



Mark Scheme (Results)

Summer 2016

Pearson Edexcel International GCSE  
in Chemistry (4CH0) Paper 2CR

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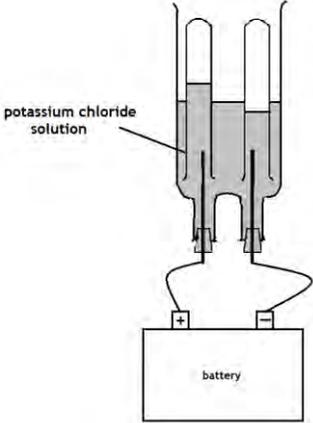
## General Marking Guidance

- All candidates must receive the same treatment. Examiners must mark the first candidate in exactly the same way as they mark the last.
- Mark schemes should be applied positively. Candidates must be rewarded for what they have shown they can do rather than penalised for omissions.
- Examiners should mark according to the mark scheme not according to their perception of where the grade boundaries may lie.
- There is no ceiling on achievement. All marks on the mark scheme should be used appropriately.
- All the marks on the mark scheme are designed to be awarded. Examiners should always award full marks if deserved, i.e. if the answer matches the mark scheme. Examiners should also be **prepared to award zero marks if the candidate's response is not worthy of credit according to the mark scheme.**
- Where some judgement is required, mark schemes will provide the principles by which marks will be awarded and exemplification may be limited.
- When examiners are in doubt regarding the application of the mark **scheme to a candidate's response, the team leader must be consulted.**
- Crossed out work should be marked UNLESS the candidate has replaced it with an alternative response.

<b>Question number</b>	<b>Answer</b>	<b>Notes</b>	<b>Marks</b>
1 a	A (the crystal dissolves)		1
b	A (it is all blue)		1
c i	4		1
ii	21		1

Question number	Answer	Notes	Marks						
2 a	<b>M1</b> oxygen / air <b>M2</b> water (vapour) / moisture	<b>ACCEPT</b> O <sub>2</sub> but not O <b>ACCEPT</b> H <sub>2</sub> O <b>IGNORE</b> steam	2						
b	(hydrated) iron(III) oxide	<b>ACCEPT</b> iron oxide / ferric oxide <b>REJECT</b> ferrous oxide and iron with other oxidation numbers <b>IGNORE</b> iron trioxide  <b>ACCEPT</b> Fe <sub>2</sub> O <sub>3</sub> (.xH <sub>2</sub> O) <b>IGNORE</b> all other formulae  If both name and formula given mark name only	1						
c	<table border="1" style="width: 100%; border-collapse: collapse;"> <tr> <td data-bbox="338 850 728 959"><b>M1</b> (galvanising)</td> <td data-bbox="728 850 1120 959">bucket / car body / railway bridge</td> </tr> <tr> <td data-bbox="338 959 728 1034"><b>M2</b> (oiling)</td> <td data-bbox="728 959 1120 1034">bicycle chain / car engine</td> </tr> <tr> <td data-bbox="338 1034 728 1109"><b>M3</b> (painting)</td> <td data-bbox="728 1034 1120 1109">car body / railway bridge</td> </tr> </table>	<b>M1</b> (galvanising)	bucket / car body / railway bridge	<b>M2</b> (oiling)	bicycle chain / car engine	<b>M3</b> (painting)	car body / railway bridge	<b>DO NOT AWARD M3</b> for car body/railway bridge if already scored for <b>M1</b>	3
<b>M1</b> (galvanising)	bucket / car body / railway bridge								
<b>M2</b> (oiling)	bicycle chain / car engine								
<b>M3</b> (painting)	car body / railway bridge								
d	<b>M1</b> zinc corrodes/oxidises/reacts in preference to iron  <b>M2</b> (because) zinc is more reactive than iron / zinc (atoms) lose electrons more readily (than do iron atoms)	<b>REJECT</b> zinc rusts <b>IGNORE</b> reference to sacrificial protection <b>ACCEPT</b> for <b>M1</b> zinc atoms react with iron(II) ions  <b>ACCEPT</b> for <b>M2</b> iron(II) ions are converted to iron atoms	2						

Question number	Answer	Notes	Marks
3 a	C (nitrogen)		1
b	A (argon)		1
c	<b>M1</b> (formula) CuO  <b>M2</b> (colour) black	<b>ACCEPT</b> correct formula as a <b>product</b> of an equation. The equation need not be balanced <b>IGNORE</b> names  <b>IGNORE</b> brown <b>REJECT</b> all other colours	2
d i	C (dilute hydrochloric acid)		1
ii	A (calcium carbonate)		1
iii	in a (gas) syringe / downward delivery in air	<b>ALLOW</b> downward delivery	1
e i	$\text{CO}_2(\mathbf{g}) + \text{Ca}(\text{OH})_2(\mathbf{aq}) \rightarrow \text{CaCO}_3(\mathbf{s}) + \text{H}_2\text{O}(\mathbf{l})$	<b>ACCEPT</b> upper case letters <b>IGNORE</b> words	1
e ii	white precipitate forms / liquid goes milky/cloudy	<b>ACCEPT</b> usual alternatives for precipitate	1

Question number	Answer	Notes	Marks						
4 a		<p><b>M1</b> both bungs inserted AND electrodes connected to battery</p> <p><b>M2</b> both tubes inverted over electrodes</p> <p><b>M3</b> solution placed in the voltameter and labelled as potassium chloride / KCl(aq)</p> <p>For <b>M3</b>, ignore all three liquid levels, except that the level in the voltameter must be above the bottoms of both tubes if present</p>	3						
b	<table border="1" data-bbox="338 852 1003 1062"> <thead> <tr> <th data-bbox="338 852 510 919">Polarity</th> <th data-bbox="510 852 1003 919">Equation</th> </tr> </thead> <tbody> <tr> <td data-bbox="338 919 510 986">-(ve)</td> <td data-bbox="510 919 1003 986"><math>(2\text{H}_2\text{O} + 2\text{e}^- \rightarrow \text{H}_2 + 2\text{OH}^-)</math></td> </tr> <tr> <td data-bbox="338 986 510 1062">+(ve)</td> <td data-bbox="510 986 1003 1062"><math>2\text{Cl}^- \rightarrow \text{Cl}_2 + 2\text{e}^-</math></td> </tr> </tbody> </table>	Polarity	Equation	-(ve)	$(2\text{H}_2\text{O} + 2\text{e}^- \rightarrow \text{H}_2 + 2\text{OH}^-)$	+(ve)	$2\text{Cl}^- \rightarrow \text{Cl}_2 + 2\text{e}^-$	<p><b>M1</b> for <math>2\text{Cl}^- \rightarrow \text{Cl}_2 + 2\text{e}^-</math></p> <p><b>ACCEPT</b> <math>2\text{Cl}^- - 2\text{e}^- \rightarrow \text{Cl}_2</math></p> <p><b>M2</b> for -(ve) in top row <b>AND</b> +(ve) in bottom row</p> <p><b>ACCEPT</b> negative and positive</p> <p><b>IGNORE</b> cathode and anode</p>	2
Polarity	Equation								
-(ve)	$(2\text{H}_2\text{O} + 2\text{e}^- \rightarrow \text{H}_2 + 2\text{OH}^-)$								
+(ve)	$2\text{Cl}^- \rightarrow \text{Cl}_2 + 2\text{e}^-$								
c	<p>burns with a pop / squeak</p> <p>OR</p> <p>use burning/lit spill / use flame to see if pop/squeak</p>	<p>Must be reference to test and result</p> <p>Reference to spill/match with no indication of flame is not enough</p> <p><b>ACCEPT</b> splint for spill</p> <p><b>REJECT</b> reference to glowing spill/splint</p> <p>Ignore flame extinguished</p> <p>'Squeaky pop test' alone is not sufficient</p>	1						

Question number	Answer	Notes	Marks												
5 a i	<table border="1" data-bbox="360 341 1115 632"> <thead> <tr> <th data-bbox="360 341 548 493">Atomic number</th> <th data-bbox="548 341 736 493">Mass number</th> <th data-bbox="736 341 925 493">Number of protons</th> <th data-bbox="925 341 1115 493">Number of neutrons</th> </tr> </thead> <tbody> <tr> <td data-bbox="360 493 548 563"></td> <td data-bbox="548 493 736 563"></td> <td data-bbox="736 493 925 563">19</td> <td data-bbox="925 493 1115 563">20</td> </tr> <tr> <td data-bbox="360 563 548 632">19</td> <td data-bbox="548 563 736 632">41</td> <td data-bbox="736 563 925 632"></td> <td data-bbox="925 563 1115 632"></td> </tr> </tbody> </table>	Atomic number	Mass number	Number of protons	Number of neutrons			19	20	19	41			<p><b>M1</b> for 19 protons in top row AND atomic number of 19</p> <p><b>M2</b> for 20 neutrons in top row</p> <p><b>M3</b> for mass number of 41</p> <p><b>ACCEPT</b> <math>\frac{(6 \times 7.4) + (7 \times 92.6)}{100}</math></p> <p>Answer must be to 1 dp Correct final answer without working scores 2 marks</p>	3
Atomic number	Mass number	Number of protons	Number of neutrons												
		19	20												
19	41														
ii	<p><b>M1</b> <math>(6 \times 0.074) + (7 \times 0.926)</math></p> <p><b>M2</b> = 6.9</p>		2												
b	<p>any two from</p> <ul data-bbox="371 962 1003 1385" style="list-style-type: none"> <li>• effervescence/fizzing/bubbles</li> <li>• potassium moves/darts/floats</li> <li>• potassium leaves white trail</li> <li>• potassium forms into a ball</li> <li>• potassium becomes smaller/disappears</li> <li>• (lilac) flame</li> </ul>	<p><b>ACCEPT</b> (hydrogen) gas given off/evolved/formed/produced</p> <p><b>IGNORE</b> name of gas</p> <p><b>ACCEPT</b> melts</p> <p><b>ACCEPT</b> dissolves</p> <p><b>IGNORE</b> colour of flame / explodes</p>	2												

Question number	Answer	Notes	Marks
5 c i	pink	<b>ALLOW</b> red <b>IGNORE</b> purple	1
	ii		1
d	<p><b>M1</b> potassium loses its outer/valence electron more easily/readily</p> <p><b>M2</b> because it is further from (the attraction of) nucleus (and therefore less strongly attracted to the nucleus)</p>	<p><b>IGNORE</b> references to more shells / larger atomic radius / more shielding / more screening</p> <p><b>ACCEPT</b> reverse arguments as long as it is clear that lithium is being considered</p>	2

Question number	Answer	Notes	Marks
6 a	<p><b>M1</b> twice as much/more carbon dioxide removed (per mole reacted)</p> <p><b>M2</b> produces oxygen (for breathing)</p>	<p><b>ACCEPT</b> reverse arguments for both <b>M1</b> and <b>M2</b> eg lithium hydroxide removes less CO<sub>2</sub> and does not produce oxygen scores 2</p> <p><b>IGNORE</b> references to the need to remove water in reaction 1</p>	2
b i	<p><b>M1</b> <math>n(\text{CO}_2) = \frac{100}{44}</math> OR 2.27(27....) (mol)</p> <p><b>M2</b> <math>n(\text{LiOH}) = \text{answer to M1} \times 2</math> OR 4.54(54....) (mol)</p> <p><b>M3</b> <math>m(\text{LiOH}) = (\text{answer to M3} \times 24) = 110</math> (g)</p> <p><b>OR</b></p> <p><b>M1</b> 48 (g) reacts with 44 (g)</p> <p><b>M2</b> x (g) reacts with 100 (g)</p> <p><b>M3</b> x = 110 (g)</p>	<p><b>ACCEPT</b> any number of sig figs except one eg 109 / 109.1 / 109.09 / 109.0909.....</p> <p>Award 3 marks for correct final answer without working</p> <p>108.96 (from 2.27) scores 3 marks 110.4 (from 2.3) scores 3 marks</p>	3

Question number	Answer	Notes	Marks
6 b ii	<p><b>M1</b> <math>n(\text{Li}_2\text{O}_2) = \frac{100}{46} = 2.17(3913\dots)</math> mol (= <math>n\text{CO}_2</math>)</p> <p><b>M2</b> volume of <math>\text{CO}_2 = \text{answer to M1} \times 24\,000</math></p> <p><b>M3</b> = 52 000 (<math>\text{cm}^3</math>)</p>	<p><b>ACCEPT</b> any number of sig figs except one eg 52 170, 52 174, 52 173.9, etc</p> <p>Award 3 marks for correct final answer without working</p> <p>52 080 (from 2.17) scores 3 marks 52 800/53 000 (from 2.2) scores 3 marks</p>	3

Question number	Answer	Notes	Marks						
7 a	<p><b>M1</b> (step 1) nitric acid</p> <p><b>M2</b> (step 2) magnesium carbonate is insoluble / magnesium carbonate does not form a solution</p> <p><b>M3</b> (step 3) boiling off all the water (will not produce a hydrated salt)</p>	<p><b>ACCEPT</b> sulfuric acid should be used</p> <p><b>REJECT</b> the use of reagents that would not work, eg magnesium chloride</p>	3						
7 b i	<table border="1" data-bbox="338 772 871 981"> <tbody> <tr> <td data-bbox="338 772 719 842"><b>M1</b> (after)</td> <td data-bbox="719 772 871 842">23.80</td> </tr> <tr> <td data-bbox="338 842 719 912"><b>M2</b> (before)</td> <td data-bbox="719 842 871 912">2.15</td> </tr> <tr> <td data-bbox="338 912 719 981"><b>M3</b> (volume added)</td> <td data-bbox="719 912 871 981">21.65</td> </tr> </tbody> </table>	<b>M1</b> (after)	23.80	<b>M2</b> (before)	2.15	<b>M3</b> (volume added)	21.65	<p>If both readings are correct but in the wrong order, award 1 mark for <b>M1</b> and <b>M2</b></p> <p><b>M3</b> CQ on the values given for <b>M1</b> and <b>M2</b></p> <p>Penalise missing trailing zeros once only</p>	3
<b>M1</b> (after)	23.80								
<b>M2</b> (before)	2.15								
<b>M3</b> (volume added)	21.65								
b ii	<p><b>M1</b> (the calculated) volume will be higher</p> <p><b>M2</b> because it includes the air (contained in the tip of the burette)</p>	<p><b>M2</b> dep on <b>M1</b></p>	2						

c	<p>i ticks in columns 2 and 4</p> <p>ii <b>M1</b> <math>\frac{26.45 + 26.25}{2}</math></p> <p><b>M2</b> 26.35 (cm<sup>3</sup>)</p>	<p>CQ on any combination of ticked results</p> <p>If no results are ticked then <b>M1</b> can only be awarded if the values from columns 2 and 4 are averaged</p> <p>If only one column ticked then no marks can be awarded in (c)(ii)</p> <p>CQ on results averaged Answers should be to 2dp, except trailing zero not needed</p> <p>Correct final answer without working scores 2</p>	<p>1</p> <p>2</p>
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Question number	Answer	Notes	Marks
7 d	<p><b>M1</b> heat/boil until crystals form in a sample of solution that has been removed (and cooled)</p> <p><b>M2</b> leave (the solution) to cool (so that crystals form)</p> <p><b>M3</b> filter (to obtain crystals)</p> <p><b>AND</b></p> <p>suitable method of drying crystals</p>	<p><b>ACCEPT</b> heat/boil to produce a (hot) saturated/concentrated solution</p> <p><b>ACCEPT</b> heat until crystals start/begin to form</p> <p><b>ALLOW</b> (heat/boil to) evaporate some of the water</p> <p><b>ALLOW</b> heat/boil to crystallisation point</p> <p><b>IGNORE</b> references to filtering before heating</p> <p><b>M2</b> dep on <b>M1</b></p> <p><b>ACCEPT</b> decant/pour off the liquid/(excess)solution for filter</p> <p>eg place in (warm) oven / leave to dry (in warm place) / use filter paper / use kitchen towel</p> <p><b>REJECT</b> any reference to heating directly with a flame, eg with a Bunsen</p> <p><b>IGNORE</b> reference to washing crystals</p> <p><b>M3</b> dep on <b>M1</b></p> <p>If <b>M1</b> not scored then award 1 mark out of 3 for leaving the solution until the water evaporates fully</p>	3



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