Please check the examination details below before entering your candidate information

Candidate surname	Other names

Centre Number | Candidate Number

Pearson Edexcel International GCSE (9–1)

Friday 17 May 2024

Morning (Time: 2 hours)

Paper reference 4CH1/1CR 4SD0/1CR

Chemistry

UNIT: 4CH1
Science (Double Award) 4SD0
PAPER: 1CR

You must have:
Calculator, ruler

Total Marks

Instructions

- Use **black** ink or ball-point pen.
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided – *there may be more space than you need.*
- Show all the steps in any calculations and state the units.

Information

- The total mark for this paper is 110.
- The marks for **each** question are shown in brackets – *use this as a guide as to how much time to spend on each question.*

Advice

- Read each question carefully before you start to answer it.
- Write your answers neatly and in good English.
- Try to answer every question.
- Check your answers if you have time at the end.

Turn over ▶

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F:1/1/1/1/1/1/1

P 7 3 4 2 4 A 0 1 3 2

Pearson

The Periodic Table of the Elements

1	2											3	4	5	6	7	0
							1 H hydrogen 1										4 He helium 2
7 Li lithium 3	9 Be beryllium 4											11 B boron 5	12 C carbon 6	14 N nitrogen 7	16 O oxygen 8	19 F fluorine 9	20 Ne neon 10
23 Na sodium 11	24 Mg magnesium 12											27 Al aluminium 13	28 Si silicon 14	31 P phosphorus 15	32 S sulfur 16	35.5 Cl chlorine 17	40 Ar argon 18
39 K potassium 19	40 Ca calcium 20	45 Sc scandium 21	48 Ti titanium 22	51 V vanadium 23	52 Cr chromium 24	55 Mn manganese 25	56 Fe iron 26	59 Co cobalt 27	59 Ni nickel 28	63.5 Cu copper 29	65 Zn zinc 30	70 Ga gallium 31	73 Ge germanium 32	75 As arsenic 33	79 Se selenium 34	80 Br bromine 35	84 Kr krypton 36
85 Rb rubidium 37	88 Sr strontium 38	89 Y yttrium 39	91 Zr zirconium 40	93 Nb niobium 41	96 Mo molybdenum 42	[98] Tc technetium 43	101 Ru ruthenium 44	103 Rh rhodium 45	106 Pd palladium 46	108 Ag silver 47	112 Cd cadmium 48	115 In indium 49	119 Sn tin 50	122 Sb antimony 51	128 Te tellurium 52	127 I iodine 53	131 Xe xenon 54
133 Cs caesium 55	137 Ba barium 56	139 La* lanthanum 57	178 Hf hafnium 72	181 Ta tantalum 73	184 W tungsten 74	186 Re rhenium 75	190 Os osmium 76	192 Ir iridium 77	195 Pt platinum 78	197 Au gold 79	201 Hg mercury 80	204 Tl thallium 81	207 Pb lead 82	209 Bi bismuth 83	[209] Po polonium 84	[210] At astatine 85	[222] Rn radon 86
[223] Fr francium 87	[226] Ra radium 88	[227] Ac* actinium 89	[261] Rf rutherfordium 104	[262] Db dubnium 105	[266] Sg seaborgium 106	[264] Bh bohrium 107	[277] Hs hassium 108	[268] Mt meitnerium 109	[271] Ds darmstadtium 110	[272] Rg roentgenium 111	Elements with atomic numbers 112–116 have been reported but not fully authenticated						

Key

relative atomic mass
atomic symbol
name
atomic (proton) number

**** The lanthanoids (atomic numbers 58–71) and the actinoids (atomic numbers 90–103) have been omitted.***

The relative atomic masses of copper and chlorine have not been rounded to the nearest whole number.

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Answer ALL questions.

Some questions must be answered with a cross in a box ☒. If you change your mind about an answer, put a line through the box ☒ and then mark your new answer with a cross ☒.

1 This question is about atomic structure.

(a) The table shows the number of protons, neutrons and electrons in five species, V, W, X, Y and Z.

The letters represent the species but are **not** symbols from the Periodic Table.

Species	Number of protons	Number of neutrons	Number of electrons
V	29	38	27
W	12	12	12
X	9	10	10
Y	6	6	8
Z	7	7	10

Choose letters from the table to answer these questions.

Each letter may be used once, more than once or not at all.

(i) Which species is an atom?

(1)

.....

(ii) Which species is an ion with a positive charge?

(1)

.....

(iii) Which species is an ion with a 3– charge?

(1)

.....

(b) (i) State what is meant by the term **atomic number**.

(1)

.....

.....

(ii) State what is meant by the term **mass number**.

(1)

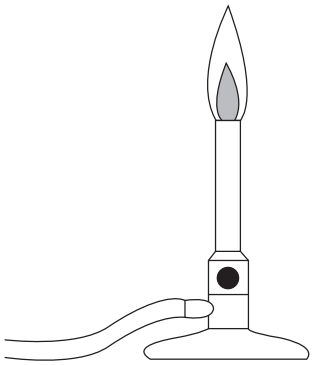
.....

.....

(Total for Question 1 = 5 marks)

2 This question is about methane, CH_4

The diagram shows a Bunsen burner that uses methane.



(a) During combustion, methane reacts with a gas in the air.

Give the name of this gas.

(1)

.....

(b) Give the two products of the complete combustion of methane.

(2)

.....

.....

(c) During the incomplete combustion of methane, carbon monoxide forms.

(i) Give a reason why carbon monoxide forms during incomplete combustion.

(1)

.....

.....

(ii) State why carbon monoxide is poisonous.

(1)

.....

.....

(d) The equation shows the reaction of methane with bromine.

$$CH_4 + Br_2 \rightarrow CH_3Br + HBr$$

Give the name of this type of chemical reaction.

(1)

.....

(Total for Question 2 = 6 marks)

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3 This question is about elements, mixtures and compounds.

(a) The box gives some methods used to separate mixtures.

crystallisation	filtration
fractional distillation	simple distillation

Choose methods from the box to answer these questions.

(i) Identify a method to remove sand from a mixture of sand and seawater.

(1)

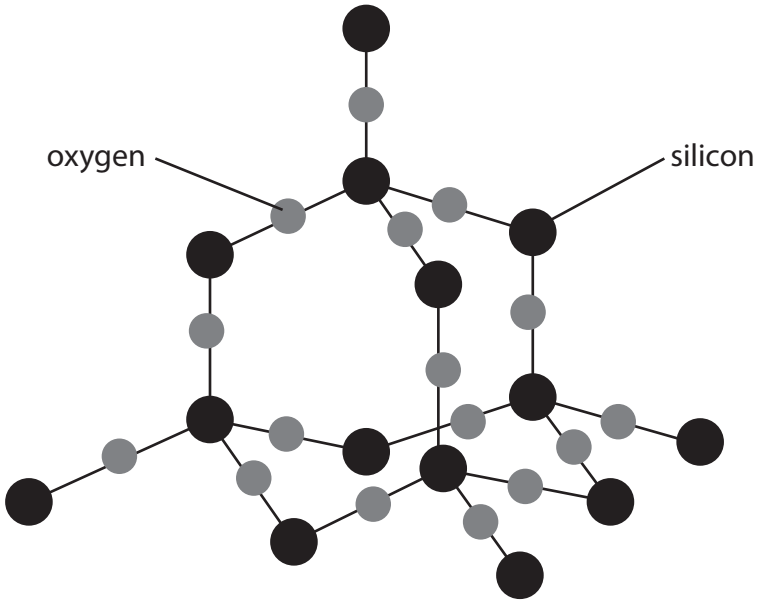
.....

(ii) Identify a method to separate a mixture of liquids with different boiling points.

(1)

.....

(b) The diagram shows part of the structure of silicon dioxide.



Explain why silicon dioxide is a compound.

(2)

.....

.....

.....

.....

(c) The molecular formula of the compound insulin is $C_{257}H_{383}N_{65}O_{77}S_6$

(i) Determine the number of different elements in $C_{257}H_{383}N_{65}O_{77}S_6$

(1)

.....

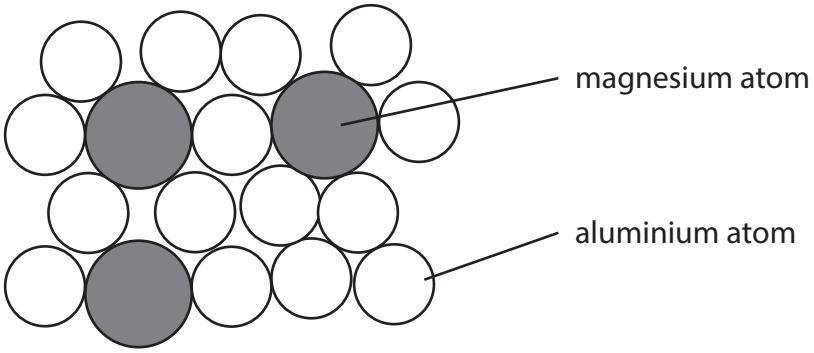
(ii) Determine the number of atoms in a molecule of $C_{257}H_{383}N_{65}O_{77}S_6$

(1)

number of atoms = ...

(d) Magnalium is a mixture of magnesium atoms and aluminium atoms.

The diagram shows a sample of magnalium.



Calculate the percentage of magnesium atoms in this sample.

(2)

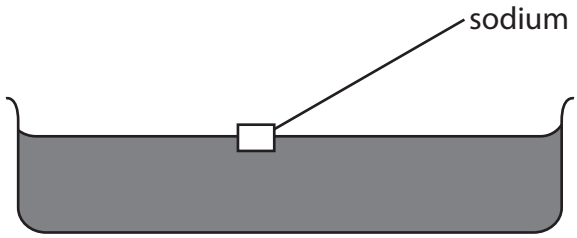
percentage = ... %

(Total for Question 3 = 8 marks)

4 This question is about the alkali metals.

A teacher demonstrates the reaction between sodium and water.

The teacher fills a trough with water and then adds a piece of sodium.



(a) The sodium reacts with the water, forming bubbles of hydrogen gas and a colourless solution.

State two other observations that would be made.

(2)

1 .....

.....

2 .....

.....

(b) Give a test to show that, at the end of the reaction, the solution contains sodium ions.

(2)

.....

.....

.....

.....

(c) Lithium, sodium and potassium react in a similar way when added to water.

(i) State, with reference to the electronic configurations of atoms, why these elements have similar reactions.

(1)

.....

.....

(ii) The table shows the atomic radius of a lithium atom, a sodium atom and a potassium atom.

Atom	Atomic radius in cm
lithium	1.82×10^{-12}
sodium	2.27×10^{-12}
potassium	2.80×10^{-12}

Deduce the relationship between the atomic radius and the reactivity of the metals.

(1)

.....

.....

(Total for Question 4 = 6 marks)

5 Chromatography is used to separate the components in a mixture.

(a) Diagram 1 shows the apparatus used to separate the different dyes in a food colouring.

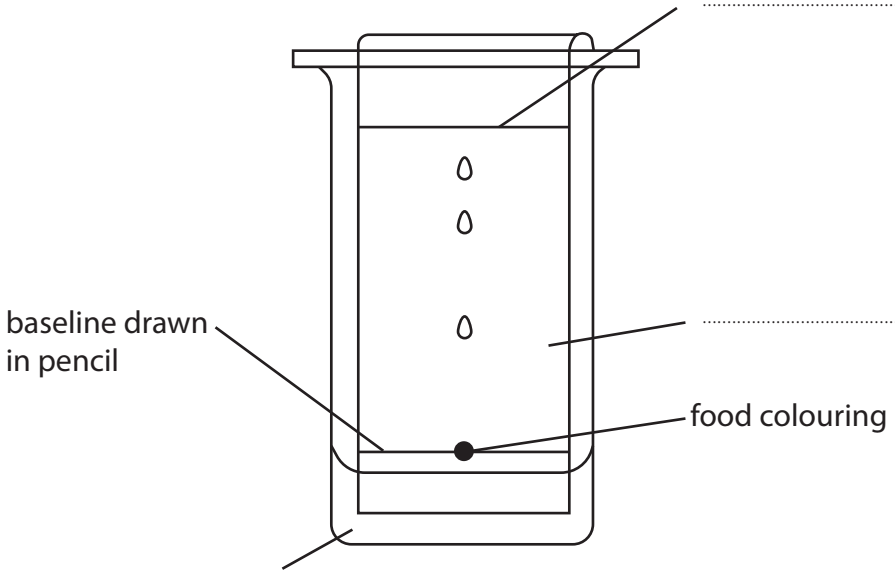


Diagram 1

(i) Complete the diagram by adding the missing labels.

(3)

(ii) Give a reason why the baseline is drawn in pencil.

(1)

(b) Diagram 2 shows a chromatogram produced from four different food colourings, W, X, Y and Z.

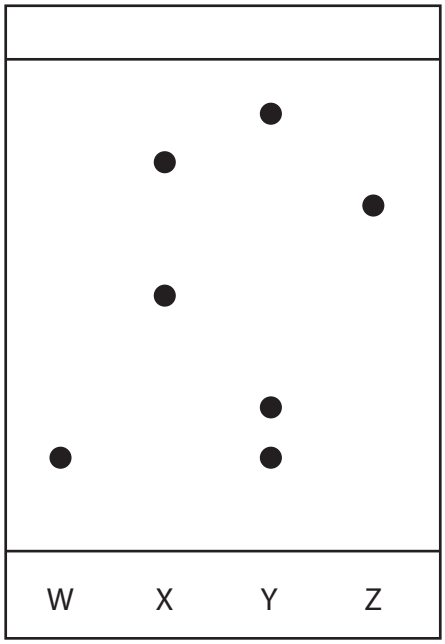


Diagram 2

(i) Which two food colourings contain the same dye?

(1)

- ☒ **A** W and X
- ☒ **B** W and Y
- ☒ **C** X and Z
- ☒ **D** Y and Z

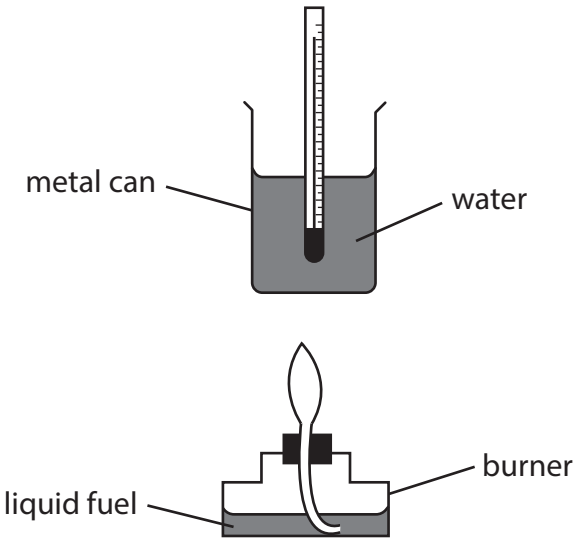
(ii) Calculate the R_f value of the dye in food colouring W.

(2)

R_f = ...

(Total for Question 5 = 7 marks)

6 A student uses this apparatus to find the heat energy released by the combustion of liquid fuels.



(a) Explain what is meant by the term **fuel**.

(2)

.....

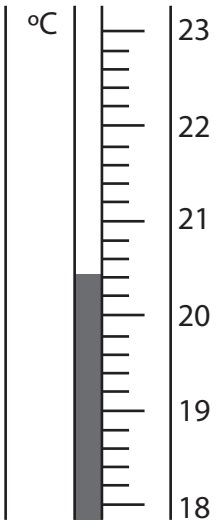
.....

.....

.....

(b) (i) In one experiment, the student uses liquid ethanol as the fuel.

The thermometer shows the temperature of the water at the start of the experiment.



Complete the table by giving the temperatures to the nearest 0.1 °C.

(2)

temperature of the water at the start in °C	
highest temperature reached in °C	
temperature rise in °C	57.2

(ii) The metal can contains water of mass 150 g.

Show, by calculation, that the heat energy change (Q) for this reaction is approximately 36 000 J.

[for water, c = 4.2 J/g/ °C]

(2)

Q = ... J

(iii) In the experiment, 2.3 g of ethanol (M_r = 46) is burned.

Calculate the molar enthalpy change (ΔH), in kJ/mol, for the combustion of ethanol, C_2H_5OH

Include a sign in your answer.

Give your answer to two significant figures.

(4)

ΔH = ... kJ/mol

(c) In this experiment, the calculated value of ΔH is less than the value given in a data book.

Give a possible reason for the difference in values.

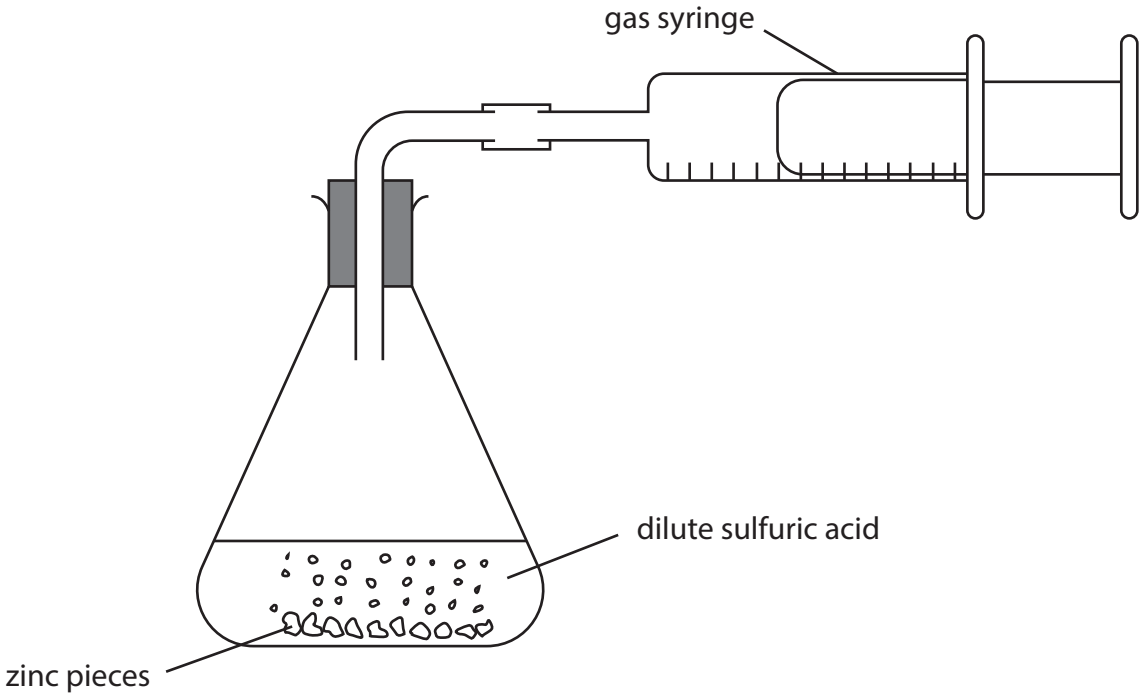
(1)

.....

.....

(Total for Question 6 = 11 marks)

7 A student uses this apparatus to investigate the rate of reaction between dilute sulfuric acid and an excess of small pieces of zinc.



This is the student's method.

Step 1 use 50 cm^3 of dilute sulfuric acid

Step 2 add approximately 5 g of small zinc pieces to the sulfuric acid

Step 3 quickly connect the gas syringe

Step 4 record the reading on the gas syringe every 30 seconds until the reaction stops

(a) (i) Name a suitable piece of apparatus to measure the volume of sulfuric acid.

(1)

.....

(ii) Give a reason why the mass of zinc pieces does not need to be measured accurately.

(1)

.....

.....

(iii) Give a reason why the student quickly connects the gas syringe in step 3.

(1)

.....

.....

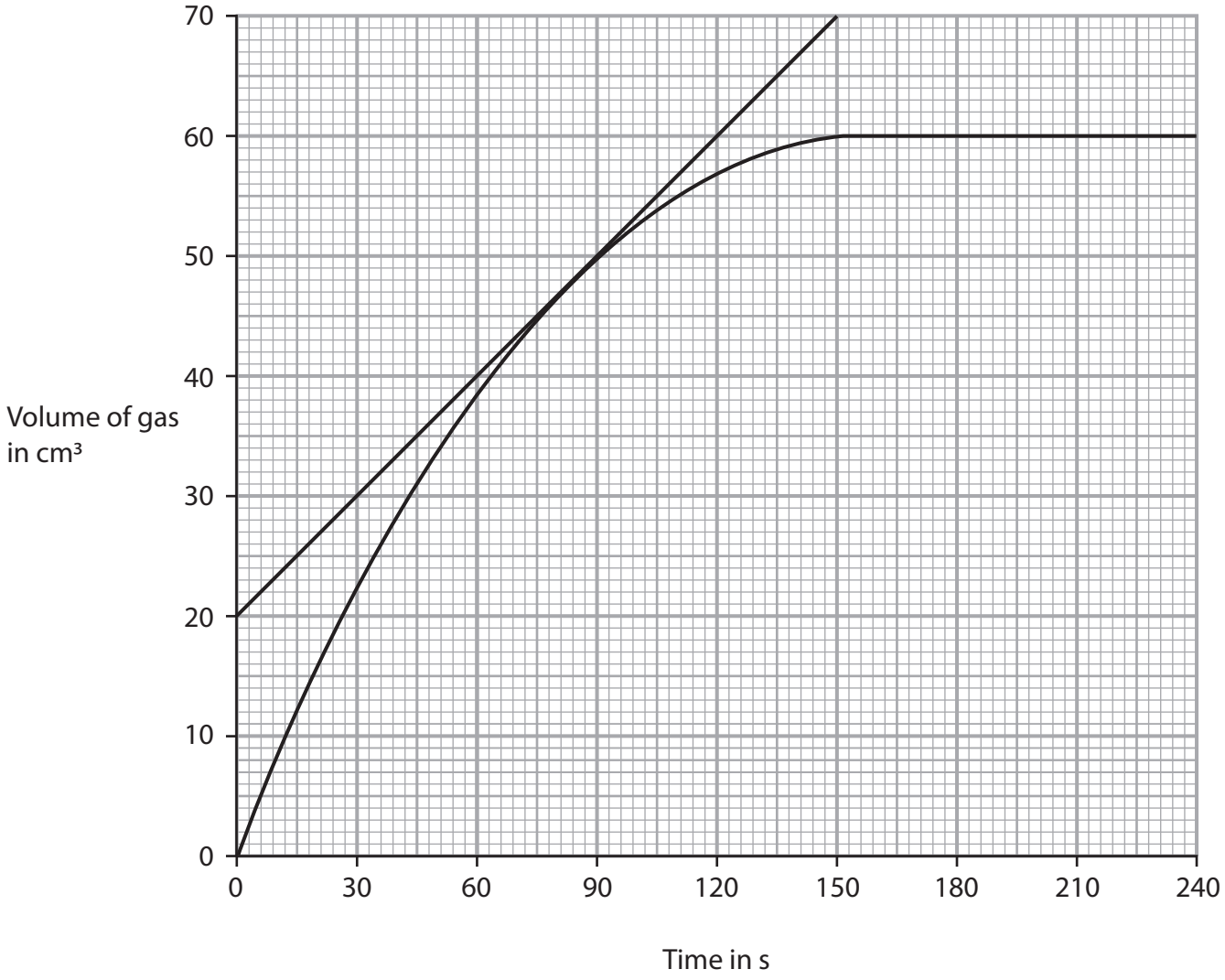
(iv) State how the student would know when the reaction stops.

(1)

.....

.....

(b) The graph shows the volume of gas collected in the syringe during the experiment.



(i) A tangent to the curve has been drawn at a time of 80 s.

Use the tangent to calculate the rate of reaction at 80 s.

Show your working on the graph.

Give the unit.

(3)

rate of reaction = ... unit .....

(ii) Explain the shape of the graph in these regions.

(6)

from 0 s to 60 s

from 60 s to 150 s

from 150 s to 240 s

(Total for Question 7 = 13 marks)

8 This question is about crude oil.

(a) Describe how crude oil is separated into fractions by fractional distillation.

(4)

(b) Some of the products of fractional distillation are then cracked.

This equation represents a reaction that occurs during cracking.

$$C_{15}H_{32} \rightarrow C_8H_{18} + C_3H_6 + 2C_2H_4$$

Explain why cracking is an important process in the oil industry.

(4)

(c) Fuels obtained from crude oil may contain impurities.

Explain how an impurity found in fuels can cause an environmental problem.

(3)

(Total for Question 8 = 11 marks)

9 (a) The table shows the formulae of some positive and negative ions, and the formulae of some compounds containing these ions.

	Cl^-	**O^{2-}**	**SO_4^{2-}**
Na^+		Na_2O	Na_2SO_4
NH_4^+	NH_4Cl		
Zn^{2+}	$ZnCl_2$		$ZnSO_4$

(i) Complete the table by giving the formulae of the missing compounds.

(3)

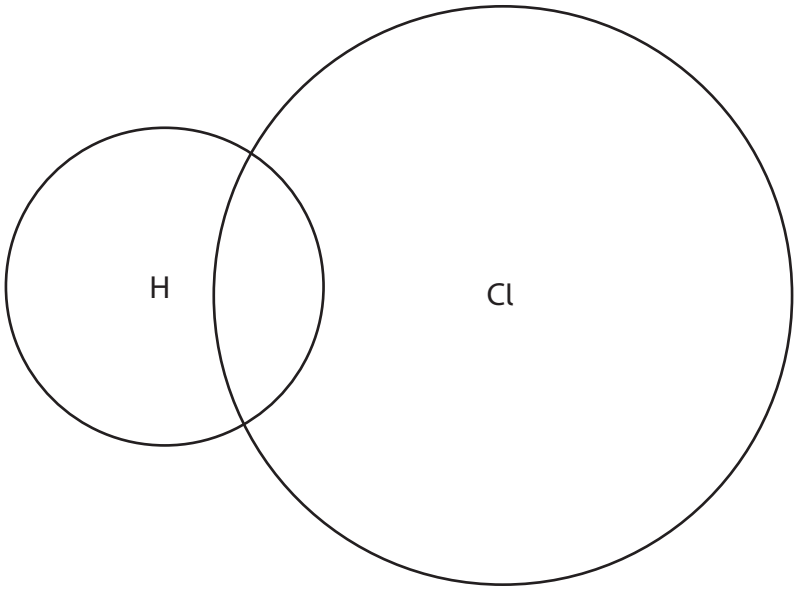
(ii) Give the name of the compound with the formula $ZnSO_4$

(1)

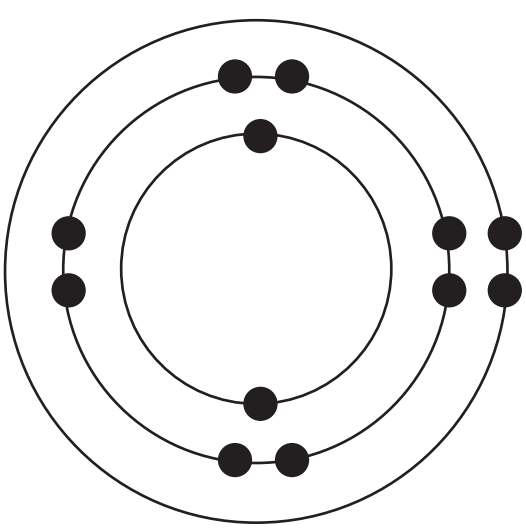
(b) Hydrogen chloride and magnesium chloride have different types of bonding and have different structures.

(i) Complete the dot-and-cross diagram to show the outer shell electrons in a molecule of hydrogen chloride.

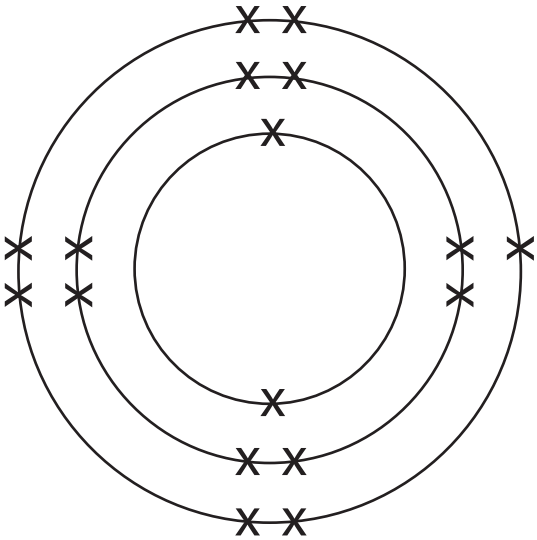
(2)



(ii) The diagram shows the electronic configuration of a magnesium atom and of a chlorine atom.



magnesium



chlorine

Draw the electronic configuration of a magnesium ion and of a chloride ion in the boxes.

Show the charge on each ion.

(3)

magnesium ion

chloride ion

(iii) Explain why magnesium chloride has a much higher melting point than hydrogen chloride.

Refer to structure and bonding in your answer.

(5)

(Total for Question 9 = 14 marks)

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10 This is the equation for the decomposition of hydrogen peroxide.

$$2H_2O_2(aq) \rightarrow 2H_2O(l) + O_2(g)$$

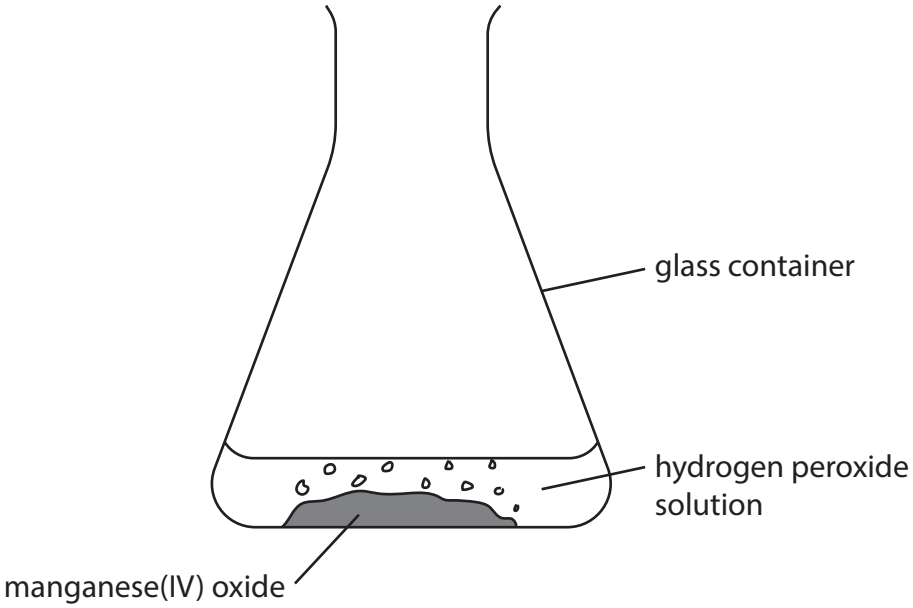
The rate of reaction increases when a catalyst of manganese(IV) oxide is added.

(a) Describe how a catalyst increases the rate of a reaction.

(2)

(b) A student adds 50 cm^3 of hydrogen peroxide solution to a glass container and then adds 1.0 g of manganese(IV) oxide.

The diagram shows the apparatus the student uses.



(i) Name the glass container the student uses.

(1)

(ii) The student waits until the hydrogen peroxide solution completely decomposes.

Describe how the student could then show that the manganese(IV) oxide was a catalyst and not a reactant.

(3)

.....

.....

.....

.....

.....

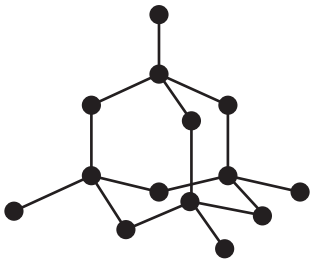
.....

(Total for Question 10 = 6 marks)

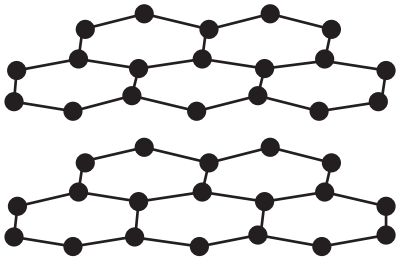
11 Diamond and graphite are both forms of the element carbon.

Diamond and graphite both have covalent bonds and giant covalent structures.

The diagram represents the structure of diamond and the structure of graphite.



diamond



graphite

(a) Give a reason why diamond is an element.

(1)

(b) Describe the forces of attraction in a covalent bond.

(2)

(c) (i) Explain why graphite conducts electricity.

(2)

(ii) Explain why diamond is hard but graphite is soft.

(4)

.....

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.....

.....

(d) Another form of carbon has molecules with the formula C_x

x represents the number of carbon atoms in each molecule.

Each molecule of C_x has a mass of 1.40×10^{-21} g.

One mole of C_x contains 6.02×10^{23} molecules.

Calculate the M_r of C_x and the value of x

[for carbon, $A_r = 12$]

(3)

M_r = ...

x = ...

(Total for Question 11 = 12 marks)

12 This question is about the metal tantalum, Ta.

Tantalum metal can be produced by heating tantalum chloride ($TaCl_5$) and hydrogen gas in a furnace.

The other product of the reaction is hydrogen chloride.

(a) Complete the equation for the reaction.

(1)

..... $TaCl_5(s)$ + ... $H_2(g)$ → ..... $Ta(s)$ + ... $HCl(g)$

(b) As tantalum chloride is heated, the mass of solid in the furnace decreases leaving tantalum as the only solid product.

The table shows the mass of solid in the furnace at one-hour intervals.

Time in hours	Mass of solid in the furnace in kg
0	2510
1	2207
2	1960
3	1506
4	1329
5	1267
6	1267
7	1267

(i) State how the data in the table shows that the reaction is complete.

(1)

.....

.....

(ii) Use the data to show that the formula of tantalum chloride is $TaCl_5$

[for tantalum, $A_r = 181$ for chlorine, $A_r = 35.5$]

(3)

QUESTION 12 CONTINUES ON NEXT PAGE.

(c) Another method of extracting tantalum is by reacting tantalum(V) oxide with carbon.

This is the equation for the reaction.

$$Ta_2O_5(s) + 5C(s) \rightarrow 2Ta(s) + 5CO(g)$$

(i) Explain why this is a redox reaction.

(2)

.....

.....

.....

.....

(ii) 2000 mol of tantalum(V) oxide is heated with 500 000 g of carbon.

Show by calculation that the carbon is in excess.

[for carbon, $A_r = 12$]

(2)

(iii) Calculate the maximum mass, in grams, of tantalum that can be obtained from 2000 mol of tantalum(V) oxide.

[for tantalum, $A_r = 181$]

(2)

mass = ... g

(Total for Question 12 = 11 marks)

TOTAL FOR PAPER = 110 MARKS