GradeMax

Cambridge International AS & A Level

Cambridge International Examinations
Cambridge International Advanced Subsidiary and Advanced Level

CHEMISTRY **9701/51**

Paper 5 Planning, Analysis and Evaluation **May/June 2016**

MARK SCHEME

Maximum Mark: 30

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge will not enter into discussions about these mark schemes.

Cambridge is publishing the mark schemes for the May/June 2016 series for most Cambridge IGCSE®, Cambridge International A and AS Level components and some Cambridge O Level components.

® IGCSE is the registered trademark of Cambridge International Examinations.

© UCLES 2016

Question	Expected Answer	Mark
1 (a)	lithium and water being labelled in an arrangement that shows them coming into contact at some time	[1]
	gas syringe OR collection over water, both using a leak-proof connection to the reaction vessel that would collect gas **after** the reagents have been mixed	[1]
	a valid separation of the two reagents	[1]
(b) (i)	cut (a small piece) AND remove oil	[1]
(ii)	no change / increase in volume OR bubbles (in collector) cease (over water) OR no more gas / hydrogen collected	[1]
(iii)	safety precautions – in a Li / LiOH / H_2 context Any two from • avoid skin contact / wear gloves / lab coat / use tongs in context of prevention of burns or corrosive contact only • keep piece of lithium / storage vessel / apparatus away from water • ensure unused lithium is all returned to storage vessel or stored under oil • keep away from naked flames / burner / sources of ignition	[2]
(iv)	$(0.1 / 7 \times 1 / 2 \times 24\,000) = 171\ cm^3$	
	correct unit **MUST** be present	[1]
(v)	use 200 cm^3 OR 250 cm^3 OR 500 cm^3 - the size must be reasonable and consistent with the volume in **b(iv)**.	
	correct unit **MUST** be present	[1]
(c) (i)	use a burette twice or 50 cm^3 pipette twice	[1]
(ii)	lithium / solid has all reacted / disappears / dissolves	[1]

© Cambridge International Examinations 2016

Question	Expected Answer	Mark
(iii)	pipette / burette / syringe / graduated pipette	[1]
(iv)	repeat the titration until results are within 0.1 cm^3	[1]
(d)	beaker : Effect: (A_r) (appears) more / larger AND Reason: (LiOH) solution more dilute (than expected)	[1]
	Flask: Effect: (A_r) None AND Reason: moles of (LiOH) (put in conical flask) remains the same / volume (i.e. moles) of LiOH not altered	[1]
		[15]

© Cambridge International Examinations 2016

Question	Expected Answer	Mark
2 (a)		

$1/T/K^{-1}$	$\log_{10}(1/t)$
3.47×10^{-3}	–1.92
3.41×10^{-3}	–1.76
3.33×10^{-3}	–1.56
3.30×10^{-3}	–1.45
3.26×10^{-3}	–1.25 or –1.26
3.22×10^{-3}	–1.28
3.19×10^{-3}	–1.18
3.16×10^{-3}	–1.08
3.12×10^{-3}	–0.95
3.05×10^{-3}	–0.90

Column values for $1/T$ correctly calculated	[1]
Column values for $\log_{10}(1/t)$ correctly calculated	[1]
3sf in $1/T$ AND 2 dp in $\log_{10}(1/t)$	[1]
(b) candidate's points plotted correctly from table in 2(a)	[1]
line of best fit drawn	[1]

© Cambridge International Examinations 2016

Question	Expected Answer	Mark
(c)	Two anomalies identified Reasons: Points to the left of the line: the time of disappearance was thought to be later OR the time was stopped too late (after reaction ended) OR the (hydrochloric acid) solution had not reached the temperature of the water bath OR the timer was started early OR magnesium folded up (reduced surface area) Points to the right of the line: the Mg may have been thought to have disappeared earlier than it did OR the timer was started late OR the timer was stopped too early (reaction still going)	[1]
(d) (i)	two co-ordinates in correct *x*, *y* format	[1]
	gradient calculated correctly from candidate's stated co-ordinates (the value **MUST** be negative unless the graph is mis-plotted) value **MUST** be to 3 significant figures Expected range –2500 to –3500	[1]
(ii)	$-E_A$ = gradient × 0.0191 OR $-E_A$ = gradient × 0.0191 then divide by 1000 OR correct transformations	[1]
	correct calculation and sign from candidate's gradient, gradient may be in calculation form, minimum 2 significant figures	[1]

© Cambridge International Examinations 2016

Question	Expected Answer	Mark
(e)	valid answer dependent on candidate's graph, e.g. reliable because most of the points on / close to the line OR unreliable as most points not on the line	[1]
(f)	Student X is correct; reaction time less OR reaction is faster AND percentage error / uncertainty will be greater OR greater error / uncertainty in time / data / recordings	[1]
(g)	reaction time is longer / rate slower AND some of the magnesium is not in contact (with the acid) OR less surface area for reaction (with HC*l*) OR only the bottom of the magnesium is reacting	[1]
(h)	initial rate lower / slower. AND the concentration of H^+ ions is lower / pH higher OR ethanoic acid less dissociated / weaker acid	[1]
		[15]

© Cambridge International Examinations 2016