GradeMax

Cambridge International AS & A Level

Cambridge International Examinations
Cambridge International Advanced Subsidiary and Advanced Level

CHEMISTRY **9701/13**

Paper 1 Multiple Choice **May/June 2018**

1 hour

Additional Materials: Multiple Choice Answer Sheet
Soft clean eraser
Soft pencil (type B or HB is recommended)
Data Booklet

* 2 8 2 3 6 2 8 5 1 7 *

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.
Do not use staples, paper clips, glue or correction fluid.
Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.
DO **NOT** WRITE IN ANY BARCODES.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A**, **B**, **C** and **D**.
Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.
Any rough working should be done in this booklet.
Electronic calculators may be used.

IB18 06_9701_13/3RP
© UCLES 2018

Section A

For each question there are four possible answers, **A**, **B**, **C** and **D**. Choose the **one** you consider to be correct.

Use of the Data Booklet may be appropriate for some questions.

1 Why is the boiling point of ammonia, NH_3, higher than the boiling point of phosphine, PH_3?

A Ammonia molecules are polar; phosphine molecules are not.

B Ammonia molecules have significant hydrogen bonding; phosphine molecules do not.

C N–H covalent bonds are stronger than P–H covalent bonds.

D There is one lone pair in each ammonia molecule but no lone pair in each phosphine molecule.

2 Neutrons are passed through an electric field. The mass of one neutron relative to $\frac{1}{12}$ the mass of a ^{12}C atom and any deflection in the electric field is recorded.

Which row is correct?

	mass of neutron	behaviour of beam of neutrons in an electric field
A	0	deflected
B	1	deflected
C	0	not deflected
D	1	not deflected

3 The table refers to the electron distribution in the second shell of an atom with eight protons.

Which row is correct for this atom?

	orbital shape 		orbital shape 	
	orbital type	number of electrons	orbital type	number of electrons
A	p	2	s	4
B	p	4	s	2
C	s	2	p	4
D	s	4	p	2

© UCLES 2018

GradeMax

4 Which statement describes the bond between carbon and hydrogen in an ethene molecule?

A a π bond between an s orbital and an sp^2 orbital

B a π bond between an s orbital and an sp^3 orbital

C a σ bond between an s orbital and an sp^2 orbital

D a σ bond between an s orbital and an sp^3 orbital

5 Aspirin, $C_9H_8O_4$, M_r = 180.0, can be made by a reaction between 2-hydroxybenzoic acid, $C_7H_6O_3$, M_r = 138.0, and ethanoic anhydride, $C_4H_6O_3$, M_r = 102.0. The balanced equation for the reaction is shown.

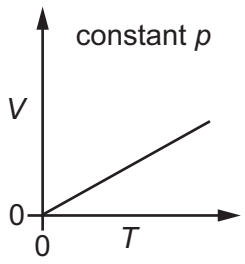
$$C_7H_6O_3 + C_4H_6O_3 \rightarrow C_9H_8O_4 + C_2H_4O_2$$

If a reaction mixture consists of 10.0 g of each of the two reactants, what is the maximum mass of aspirin that can be produced?

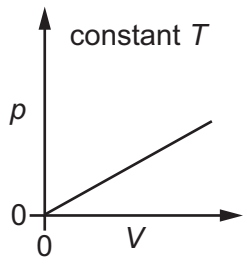
A 5.7 g **B** 10.0 g **C** 13.0 g **D** 17.6 g

6 Which diagram correctly describes the behaviour of a fixed mass of an ideal gas? (*T* is measured in K.)

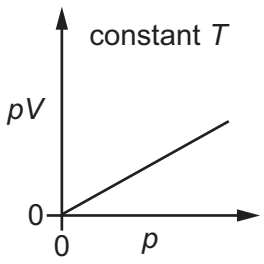
A



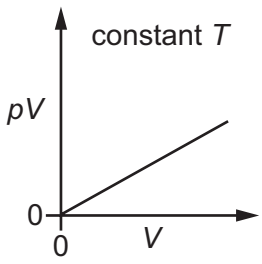
B



C



D



GradeMax

7 Anhydrous copper(II) chloride, $CuCl_2$, combines with water to form $CuCl_2.2H_2O$. The standard enthalpy changes of formation for this reaction are shown in the table.

	$\Delta H_f^\ominus$ / kJ mol^{-1}
H_2O	–286
$CuCl_2$	–206
$CuCl_2.2H_2O$	–808

What is the standard enthalpy change of the reaction shown?

$$CuCl_2 + 2H_2O \rightarrow CuCl_2.2H_2O$$

A –1586 kJ mol^{-1}

B –316 kJ mol^{-1}

C –110 kJ mol^{-1}

D –30 kJ mol^{-1}

8 Xenon hexafluoride, XeF_6, reacts with water.

$$XeF_6 + 3H_2O \rightarrow XeO_3 + 6HF$$

Which statement is correct?

A Hydrogen is reduced in this reaction.

B Hydrogen is the only element oxidised in this reaction.

C The only element oxidised in this reaction is xenon.

D This is not a redox reaction.

9 Hydrogen is produced industrially from methane as shown in the equation.

$$CH_4(g) + H_2O(g) \rightleftharpoons CO(g) + 3H_2(g) \qquad \Delta H^\ominus = +205\,kJ\,mol^{-1}$$

Which conditions would most favour the formation of hydrogen?

	pressure	temperature
A	high	high
B	high	low
C	low	high
D	low	low

© UCLES 2018

GradeMax

10 The chemical equilibrium shown is formed when ammonia is produced in the Haber process.

$$N_2 + 3H_2 \rightleftharpoons 2NH_3$$

The following concentrations are found to be present at equilibrium under certain conditions.

N_2	H_2	NH_3
$0.200\,mol\,dm^{-3}$	$0.300\,mol\,dm^{-3}$	$0.600\,mol\,dm^{-3}$

What is the numerical value of K_c under these conditions?

A 0.0150 **B** 6.0 **C** 10.0 **D** 66.7

11 The enzyme maltase speeds up the reaction between maltose and water.

$$\text{maltose + water} \xrightarrow{\text{maltase}} \text{glucose}$$

Maltase shows **specificity**.

Which statement describes the **specificity** of maltase?

A Maltase is a biological catalyst and it is a type of protein.

B Maltase is most effective between pH 6.1 and pH 6.8.

C Maltase lowers the activation energies of the reactions it catalyses.

D Maltase only speeds up a small number of chemical reactions.

12 Which description of the bonding and acid / base nature of aluminium oxide is correct?

	bonding	acid / base nature
A	simple covalent	amphoteric
B	giant covalent	basic only
C	ionic	amphoteric
D	ionic	basic only

13 X and Y are elements of the third period.

X and Y are individually heated in excess chlorine. Each product is purified and then separately added to water, producing two solutions. Both solutions have a pH of less than 5.

What could be X and Y?

A Na and P **B** Mg and A*l* **C** Mg and Si **D** Si and P

© UCLES 2018

[Turn over

GradeMax

14 An ore contains magnesium carbonate and barium carbonate. A sample of the ore is dissolved in nitric acid.

How could this solution be processed into a magnesium compound and a separate barium compound?

A Add HCl(aq), filter off the solid barium chloride.

B Add HCl(aq), filter off the solid magnesium chloride.

C Add H_2SO_4(aq), filter off the solid barium sulfate.

D Add H_2SO_4(aq), filter off the solid magnesium sulfate.

15 When calcium and calcium hydride, CaH_2, react separately with water, they each produce a white solid and a colourless gas. The white solid is the same compound in each reaction.

Which statement is correct?

A Both Ca and CaH_2 produce H_2.

B Both Ca and CaH_2 produce O_2.

C Ca produces H_2 and CaH_2 produces O_2.

D Ca produces O_2 and CaH_2 produces H_2.

16 When concentrated sulfuric acid is added to solid sodium chloride, HCl is formed but not Cl_2.

When concentrated sulfuric acid is added to solid sodium iodide, I_2 is formed.

Which statement explains these observations?

A Sulfuric acid is an oxidising agent and chloride ions are more easily oxidised.

B Sulfuric acid is an oxidising agent and iodide ions are more easily oxidised.

C Sulfuric acid is a reducing agent and chloride ions are more easily reduced.

D Sulfuric acid is a reducing agent and iodide ions are more easily reduced.

© UCLES 2018

17 Silver chloride and silver iodide form equilibria when added to water.

$$AgCl(s) \rightleftharpoons Ag^+(aq) + Cl^-(aq) \qquad K_c = K_1$$

$$AgI(s) \rightleftharpoons Ag^+(aq) + I^-(aq) \qquad K_c = K_2$$

Each equilibrium position lies well to the **left**.

Silver iodide will not dissolve in aqueous ammonia. Silver chloride will dissolve in aqueous ammonia. Another equilibrium is formed.

$$Ag^+(aq) + 2NH_3(aq) \rightleftharpoons Ag(NH_3)_2^+(aq) \qquad K_c = K_3$$

The position of this equilibrium lies to the **right**.

What is the order of magnitude for these three equilibrium constants?

A $K_1 > K_2 > K_3$

B $K_2 > K_1 > K_3$

C $K_3 > K_1 > K_2$

D $K_3 > K_2 > K_1$

18 Elements and their compounds are important as catalysts.

In which process is a compound used, rather than an element?

A catalytic converters

B Contact process

C Haber process

D hydrogenation of alkenes

19 The gaseous products of heating a mixture of $Ca(OH)_2$ and NH_4Cl are passed through solid CaO. This absorbs water vapour and a gas, **W**, is collected.

A sample of **W** is oxidised by $Cl_2(g)$ to produce two gases, **X** and **Y**.

X is an element. **Y** is acidic.

Y reacts with **W** to produce **Z**.

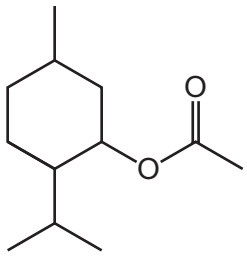
What are **X** and **Z**?

	X	**Z**
A	N_2	$CaCl_2$
B	N_2	NH_4Cl
C	O_2	$CaCl_2$
D	O_2	NH_4Cl

© UCLES 2018

GradeMax

20 Molecule G is shown.

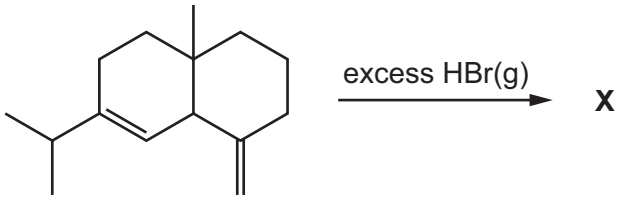


G

How many chiral centres are present in each molecule of G?

A 1 **B** 2 **C** 3 **D** 4

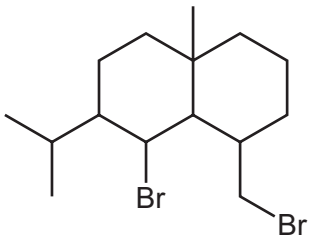
21 Sibirene, $C_{15}H_{24}$, is reacted with an excess of HBr(g). The major product is **X**.



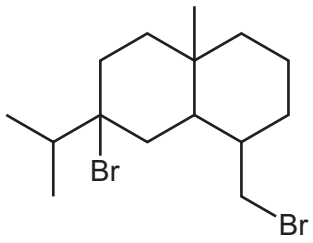
sibirene

What is the skeletal formula of **X**?

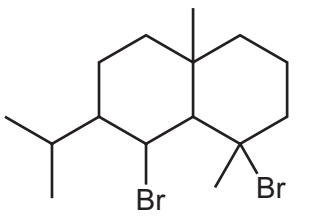
A



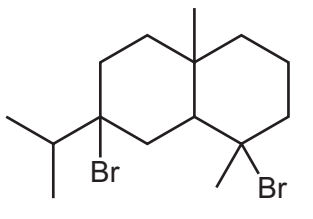
B



C



D



© UCLES 2018

22 Which statement is **not** correct?

A Combustion of PVC produces a highly acidic gas.

B PVC molecules are saturated.

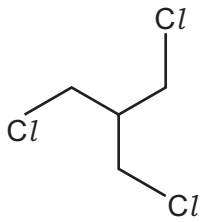
C The empirical formula of PVC is the same as the empirical formula of its monomer.

D The repeat unit of PVC is ~~(~~CHC*l*CHC*l*~~)~~.

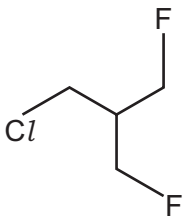
23 The presence of a halogen in an organic compound may be detected by warming the organic compound with aqueous silver nitrate.

Which compound would be the quickest to produce a precipitate?

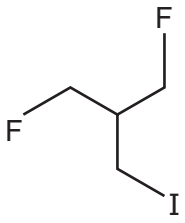
A



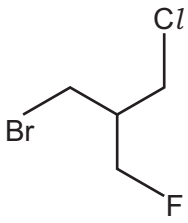
B



C



D



24 Halogenoalkanes react with NaOH(aq) either by an S_N1 mechanism or by an S_N2 mechanism. The mechanism followed by the reaction depends on the structure of the halogenoalkane.

This question is about the reaction of 3-bromo-3-ethylpentane, $(C_2H_5)_3CBr$.

Which statement is correct?

A The mechanism is S_N1, due to the stabilisation of an intermediate anion by three alkyl groups.

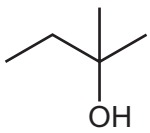
B The mechanism is S_N1, due to the stabilisation of an intermediate cation by three alkyl groups.

C The mechanism is S_N2, due to the stabilisation of an intermediate anion by three alkyl groups.

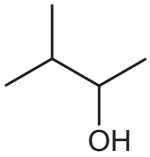
D The mechanism is S_N2, due to the stabilisation of an intermediate cation by three alkyl groups.

25 Which compound is a secondary alcohol that can be dehydrated to form an alkene with M_r = 70?

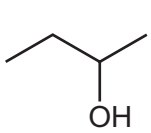
A



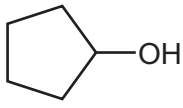
B



C



D



GradeMax

26 When 0.0075 mol of alcohol X are completely burnt in excess oxygen and the gases produced are passed through an excess of limewater (calcium hydroxide solution), 3.0 g of calcium carbonate are produced.

When X is warmed with acidified potassium dichromate(VI) there is a colour change from orange to green.

What could be the identity of X?

A $CH_3CH(OH)CH_2CH_3$

B $(CH_3)_3COH$

C $CH_3CH_2CH_2OH$

D $CH_3CH(OH)CH_3$

27 Considering **only** structural isomers, what is the number of alcohols of each type with the formula $C_5H_{12}O$?

	primary	secondary	tertiary
A	3	3	2
B	4	2	2
C	4	3	1
D	5	2	1

28 A student carried out a **two-stage** synthesis in which $CH_3CH_2CH_2Br$ was converted into $CH_3CH_2CH_2CO_2H$.

Which compound could have been formed by the first stage of this synthesis?

A $CH_3CH_2CH_2OH$

B $CH_3CH_2CH_2CHO$

C CH_3CH_2CN

D $CH_3CH_2CH_2CN$

29 An ester X has the structural formula $CH_3CO_2CH(CH_3)CH_2CH_3$. X can be prepared by heating an alcohol **Y**, under reflux, with ethanoic acid and an acid catalyst.

What is the correct name for **Y**?

A butan-1-ol

B butan-2-ol

C butan-3-ol

D methylpropan-2-ol

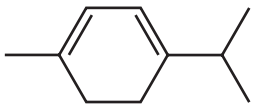
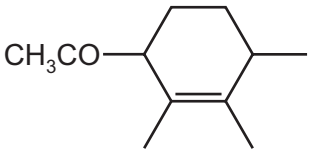
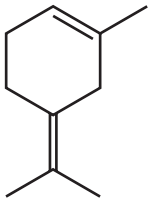
© UCLES 2018

30 Compound **S** can be extracted from natural compounds. Reacting **S** with hot, concentrated $KMnO_4$ produces the organic product, **T**. Some of the absorptions found in the infra-red spectra of **S** and **T** are described.

S has no strong absorption between 1670 and 1740 cm^{-1}.

T has a strong absorption at 1720 cm^{-1} but has **no** strong, broad absorption between 2500 and 3000 cm^{-1}.

From this information, what could be the formulae of **S** and **T**?

	S	**T**
A	$CH_3(CH_2)_5CH{=}CH_2$	$CH_3(CH_2)_5CO_2H$
B		$CH_3COCH_2CH_2COCH(CH_3)_2$
C		$CH_3COCH(COCH_3)CH_2CH_2CH(COCH_3)CH_3$
D		$HO_2CCH_2CH_2COCH_2COCH_3$

GradeMax

Section B

For each of the questions in this section, one or more of the three numbered statements **1** to **3** may be correct.

Decide whether each of the statements is or is not correct (you may find it helpful to put a tick against the statements that you consider to be correct).

The responses **A** to **D** should be selected on the basis of

A	B	C	D
1, **2** and **3** are correct	**1** and **2** only are correct	**2** and **3** only are correct	**1** only is correct

No other combination of statements is used as a correct response.

Use of the Data Booklet may be appropriate for some questions.

31 Complete combustion of a sample of a hydrocarbon gave 0.132 g of carbon dioxide and 0.054 g of water.

Which formulae could correctly represent this hydrocarbon?

1 $CH_2{=}CH_2$

2 $CH_3CH{=}CH_2$

3 $CH_3CH{=}CHCH_3$

32 Which statements are correct?

1 Magnesium carbonate decomposes at a lower temperature than calcium carbonate.

2 Calcium hydroxide is more soluble in water than magnesium hydroxide.

3 Calcium is a stronger reducing agent than magnesium.

33 Ammonia and chlorine react as shown.

$$8NH_3 + 3Cl_2 \rightarrow N_2 + 6NH_4Cl$$

Which statements are correct?

1 Each nitrogen atom is oxidised.

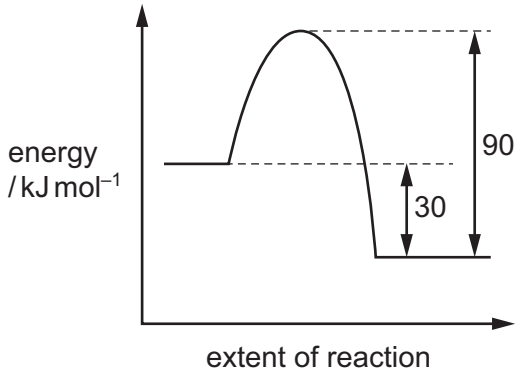
2 Each chlorine atom is reduced.

3 Ammonia behaves as a base.

© UCLES 2018

GradeMax

34 The diagram shows the reaction pathway for a reversible reaction.



Which statements are correct?

1 The forward reaction is exothermic.

2 The enthalpy change for the forward reaction is –30 kJ mol^{-1}.

3 The enthalpy change for the backward reaction is +90 kJ mol^{-1}.

35 The structure of metals is considered to be positive ions surrounded by delocalised electrons.

The melting points of the metals in Period 3 increase with increasing atomic number.

Which statements help to explain this trend from sodium to aluminium?

1 The charge on the metal ion increases.

2 There are more delocalised electrons per metal ion.

3 The radius of the metal ion decreases.

36 Under room conditions, 600 cm^3 of a gas, X, has a mass of 0.700 g.

What could X be?

1 carbon monoxide

2 ethene

3 nitrogen

37 Which compounds, on reaction with $NaBH_4$, form a compound with a chiral carbon atom?

1 $CH_3CH_2CH_2COCH_3$

2 $CH_2CHCOCH_2CH_3$

3 $CH_3CH_2COCH_2CH_3$

© UCLES 2018

[Turn over

GradeMax

The responses **A** to **D** should be selected on the basis of

A	B	C	D
1, **2** and **3** are correct	**1** and **2** only are correct	**2** and **3** only are correct	**1** only is correct

No other combination of statements is used as a correct response.

38 Compound Y is a straight chain molecule with formula $C_nH_{2n+1}X$. X is a halogen. The M_r of Y is 137. The halogen atom is on the second carbon atom in the chain.

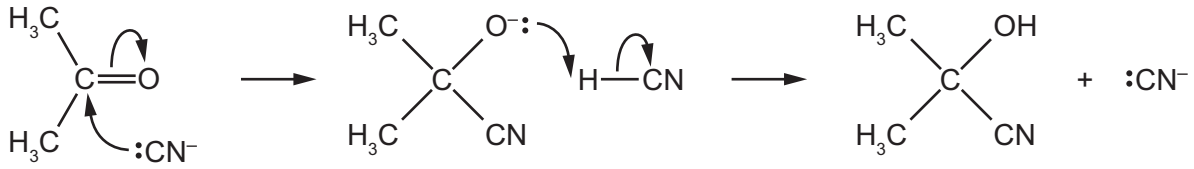
Which statements are correct?

1 Y contains a chiral centre.

2 Y can eliminate HX to form two structurally isomeric alkenes.

3 Y can eliminate HX to form two geometrically isomeric alkenes.

39 Propanone and hydrogen cyanide react together by the mechanism shown.



Which statements about this mechanism are correct?

1 CN^- is an electrophile.

2 It is an addition reaction.

3 Heterolytic bond breaking is involved.

40 Acrolein is an organic compound with the molecular formula C_3H_4O. It is used in water treatment and as a herbicide. When acrolein reacts with 2,4-dinitrophenylhydrazine an orange precipitate is obtained. Reaction of acrolein with Tollens' reagent produces a silver mirror.

Which statements are correct?

1 Acrolein reacts with alkaline aqueous iodine to produce a yellow precipitate.

2 Acrolein can be reduced to a primary alcohol.

3 Acrolein decolourises bromine water.

© UCLES 2018

GradeMax

BLANK PAGE

BLANK PAGE

Permission to reproduce items where third-party owned material protected by copyright is included has been sought and cleared where possible. Every reasonable effort has been made by the publisher (UCLES) to trace copyright holders, but if any items requiring clearance have unwittingly been included, the publisher will be pleased to make amends at the earliest possible opportunity.

To avoid the issue of disclosure of answer-related information to candidates, all copyright acknowledgements are reproduced online in the Cambridge International Examinations Copyright Acknowledgements Booklet. This is produced for each series of examinations and is freely available to download at www.cie.org.uk after the live examination series.

Cambridge International Examinations is part of the Cambridge Assessment Group. Cambridge Assessment is the brand name of University of Cambridge Local Examinations Syndicate (UCLES), which is itself a department of the University of Cambridge.