



Cambridge
International
AS & A Level

Cambridge Assessment International Education
Cambridge International Advanced Subsidiary and Advanced Level

CHEMISTRY

9701/34

Paper 3 Advanced Practical Skills 2

October/November 2018

MARK SCHEME

Maximum Mark: 40

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge International will not enter into discussions about these mark schemes.

Cambridge International is publishing the mark schemes for the October/November 2018 series for most Cambridge IGCSE™, Cambridge International A and AS Level components and some Cambridge O Level components.

This document consists of **8** printed pages.

PUBLISHED**Generic Marking Principles**

These general marking principles must be applied by all examiners when marking candidate answers. They should be applied alongside the specific content of the mark scheme or generic level descriptors for a question. Each question paper and mark scheme will also comply with these marking principles.

GENERIC MARKING PRINCIPLE 1:

Marks must be awarded in line with:

- the specific content of the mark scheme or the generic level descriptors for the question
- the specific skills defined in the mark scheme or in the generic level descriptors for the question
- the standard of response required by a candidate as exemplified by the standardisation scripts.

GENERIC MARKING PRINCIPLE 2:

Marks awarded are always **whole marks** (not half marks, or other fractions).

GENERIC MARKING PRINCIPLE 3:

Marks must be awarded **positively**:

- marks are awarded for correct/valid answers, as defined in the mark scheme. However, credit is given for valid answers which go beyond the scope of the syllabus and mark scheme, referring to your Team Leader as appropriate
- marks are awarded when candidates clearly demonstrate what they know and can do
- marks are not deducted for errors
- marks are not deducted for omissions
- answers should only be judged on the quality of spelling, punctuation and grammar when these features are specifically assessed by the question as indicated by the mark scheme. The meaning, however, should be unambiguous.

GENERIC MARKING PRINCIPLE 4:

Rules must be applied consistently e.g. in situations where candidates have not followed instructions or in the application of generic level descriptors.

GENERIC MARKING PRINCIPLE 5:

Marks should be awarded using the full range of marks defined in the mark scheme for the question (however; the use of the full mark range may be limited according to the quality of the candidate responses seen).

GENERIC MARKING PRINCIPLE 6:

Marks awarded are based solely on the requirements as defined in the mark scheme. Marks should not be awarded with grade thresholds or grade descriptors in mind.

Question	Answer	Marks
1(a)	I Unambiguous data recorded (mass or g) - two weighings with clear indication to what they refer. (mass of) FB 1, correctly subtracted and between 2.80 and 3.00	1
	II All the following data is recorded - rough titration: both burette readings and the titre - initial and final burette readings for two (or more) accurate titrations	1
	III Titre values recorded for accurate titrations, and correct headings and units in the accurate titration table - initial / start and (burette) reading / volume (<i>allow vol, but not V</i>) - final / end and (burette) reading / volume - titre or volume / FB 2 and used / added (<i>not 'difference' or 'total volume'</i>) - unit: / cm³ or (cm³) or in cm³ (for each heading) or cm ³ unit given for each volume recorded	1
	IV All accurate burette readings are recorded to the nearest 0.05 cm ³ . <i>The requirement to record to 0.05 applies to burette readings, including 0.00 cm³ (if this was the initial reading), but it does not apply to the titre.</i> <i>Do not award this mark if:</i> 50(.00) is used as an initial burette reading - more than one final burette reading is 50.(00) - any burette reading is greater than 50.(00)	1
	V The final accurate titre recorded is within 0.10 cm ³ of any other accurate titre.	1
	Examiner calculates supervisor mean titre and hence the supervisor's value of mean titre / mass of FB 1 . Examiner calculates candidate's mean titre and hence the candidate's value of mean titre / mass of FB 1 . Examiner compares the two values.	
	Q Award VI, VII and VIII if $\delta \leq 0.10$ (cm ³ g ⁻¹)	1
	Q Award VI and VII if $0.10 < \delta \leq 0.30$	1
	Q Award VI only if $0.30 < \delta \leq 0.50$	1

Question	Answer	Marks
1(b)	Candidate calculates the mean correctly. • Candidate must take the average of two (or more) titres that are within a total spread of not more than 0.20 cm³. • Working / explanation must be shown or ticks must be put next to the two (or more) accurate readings selected. • The mean should be quoted to 2 dp, and be rounded to nearest 0.01 cm³. (e.g. 26.667 cm³ must be rounded to 26.67 cm³) Two special cases, where the mean need not be to 2 dp: • Allow mean expressed to 3 dp only for 0.025 or 0.075 (e.g. 26.325 cm³) • Allow mean if expressed to 1 dp, if all accurate burette readings were given to 1 dp and the mean is exactly correct. (e.g. 26.0 and 26.2 = 26.1 is allowed) (e.g. 26.0 and 26.1 = 26.1 is wrong – mean should be 26.05) Do not award this mark if: • The rough titre was used to calculate the mean. • The candidate did only one accurate titration. Note: the candidate's mean will sometimes be marked correct even if it was different from the mean calculated by the Examiner for the purpose of assessing accuracy.	1
1(c)(i)	All answers quoted in parts (ii)–(v) are given to 3 or 4 sig fig. A minimum of three answers is required to earn this mark	1
1(c)(ii)	Correctly calculates moles of HCl No of moles of HCl used = $0.105 \times \frac{\text{mean titre}}{1000}$	1
1(c)(iii)	Equation with correct state symbols and moles NaHCO₃ correct $\text{NaHCO}_3(\text{aq/s}) + \text{HCl}(\text{aq}) \rightarrow \text{NaCl}(\text{aq}) + \text{CO}_2(\text{aq/g}) + \text{H}_2\text{O}(\text{l})$ and no of moles of NaHCO₃ = ans in (ii)	1
1(c)(iv)	Correctly uses (c)(iii) for moles of NaHCO₃ No of moles of NaHCO₃ in FB 1 = 10 × answer (iii)	1
1(c)(v)	Correctly uses answer (c)(iv) in calculating % by mass of NaHCO₃ $\% \text{ by mass} = \left(\frac{\text{ans (iv)} \times 84}{\text{mass of FB1 used in (a)}} \right) \times 100$ 84 may be shown as the sum of the A_r values.	1

Question	Answer	Marks
2(a)	I Six headings and units in list/table of data: • Mass of crucible and lid (<i>not 'weight' in any reading</i>) • Mass of crucible, lid and FB 1 (or 'contents before heating') • Mass of crucible, lid and residue / contents after first heating / heating 1 / owtte • Mass of crucible, lid and residue / contents after 2nd heating • Mass of FB 1 (used) • Mass of residue <i>Do not allow 'mass of FB 1 after heating' in third or fourth weighing.</i> <i>Unit must be given correctly in each case</i> <i>Inclusion or not of lid must be consistent.</i>	1
	II Four weighings recorded, all in the space provided • Four weighings (or more) recorded in the space provided. • All readings recorded to same number of decimal places. • Final weighing is within 0.04 g of the previous one. <i>If balance used read to 1 dp, the 3rd and 4th readings must be the same.</i>	1
	III Mass of FB 1 and residue , both written in the space provided • Mass of FB 1 used, correctly subtracted • Mass of FB 1 used was between 2.8 and 3.0 g • Mass of residue correctly subtracted.	1
	Accuracy marks • For assessment of accuracy, examiner must check and correct (if necessary) the masses of FB 1 used and of residue obtained by the supervisor and by the candidate. • Work out ratio $\frac{\text{mass of FB 1}}{\text{mass of residue (from lower mass weighing)}}$ for the supervisor (2 dp) • Work out ratio (mass FB 1: mass residue) for candidate (2 dp) • Calculate δ, the difference between these two ratios. Award IV and V if $\delta \leq 0.05$ Award IV if $0.05 < \delta \leq 0.10$	2

Question	Answer	Marks
2(b)(i)	Correctly calculates moles of CO₂ - No of moles CO₂ = $\frac{\text{mass lost}}{62}$ - Answer must be correct and expressed to 2–4 sig fig. <i>Mass loss must be correctly subtracted, either using original weighings or using the masses of FB1 and residue.</i>	1
2(b)(ii)	Correctly uses (i) for mass of NaHCO₃ Mass of NaHCO ₃ = ans (i) × 2 × 84 and answer to 2–4 sig fig Do not penalise sf more than once in 2(b)	1
2(b)(iii)	Correctly calculates percentage mass of NaHCO₃ $\% = \frac{\text{ans (ii)}}{\text{mass of FB1 used}} \times 100$ and answer to 2–4 sig fig.	1
2(c)(i)	To obtain constant mass (<i>owtte</i>) This shows that the reaction / decomposition is complete / only Na ₂ CO ₃ is left	1
	<i>Allow: to remove all CO₂ and H₂O</i>	1
2(c)(ii)	(Assumption is that) the impurity does not decompose when heated or The impurity does not react with air when heated or The impurity does not react with NaHCO ₃ / Na ₂ CO ₃ on heating	1
2(c)(iii)	Disagree because there is no frothing / spitting out (<i>owtte</i>)	1
2(c)(iv)	Possible answers, stating and explaining which experiment is more accurate: - Titration, because repeated titres are consistent. - Gravimetric, because fewer measurements / fewer cumulative / combined errors - Gravimetric, because heating is to constant mass. - Gravimetric, because colour change (of methyl orange) is indistinct / gradual.	1

Question	Answer	Marks
FB 4 is a mixture of $\text{Na}_2\text{CO}_3 + \text{NaCl}$		
3(a)(i)	Fizzing / effervescence / bubbling	1
	<u>Gas</u> / CO_2 turns lime water milky / cloudy white or gas gives white precipitate with lime water	1
3(a)(ii)	White ppt with AgNO_3 and ppt is soluble in (excess) ammonia / colourless solution formed.	1
3(a)(iii)	Either equation correct: $\text{Na}_2\text{CO}_3 + 2\text{HNO}_3 \rightarrow 2\text{NaNO}_3 + \text{CO}_2 + \text{H}_2\text{O}$ or $\text{Ca}(\text{OH})_2 + \text{CO}_2 \rightarrow \text{CaCO}_3 + \text{H}_2\text{O}$	1
3(b)(i)	Blue precipitate (with CuSO_4) and (pale) (not dark) blue solution (formed with H_2SO_4)	1
	Fizzing / effervescence / bubbling with acid in both this test and addition of acid to BaCO_3 ppt in next test.	1
	White precipitate (with Ba^{2+}) and precipitate dissolves / colourless solution formed (with HNO_3)	1
	Yellow / orange-yellow / yellowy-orange (solution) (formed)	1
	White / off-white precipitate (with Ag^+)	1
	Fizzes / bubbles / effervescence (with HNO_3) or ppt is <u>partially</u> soluble (in HNO_3) (<i>owtte</i>)	1
3(b)(ii)	Fizzing / effervescence / bubbling with aluminium	1
3(b)(iii)	Sodium chloride / NaCl	1
3(b)(iv)	Mark ecf on colour Yellow: alkaline / neutral / weakly acidic Orange: neutral / weakly acidic	1