



Cambridge International Examinations
Cambridge International General Certificate of Secondary Education

CHEMISTRY**0620/32**

Paper 3 Theory (Core)

May/June 2016**MARK SCHEME**Maximum Mark: 80

Published

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This document consists of **9** printed pages.

Page 2	Mark Scheme	Syllabus	Paper
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Question	Answer	Marks
1(a)(i)	B and D;	1
1(a)(ii)	C; has only one type of atom;	2 1 1
1(a)(iii)	Na ₃ P;	1
1(b)(i)	16;	1
1(b)(ii)	5;	1
1(b)(iii)	60;	1
1(c)	acidic; because phosphorus is a non-metal/it is a non-metal oxide/it would react with bases/neutralises bases/ phosphorus is on the right-hand side of the Periodic Table;	2 1 1

Page 3	Mark Scheme	Syllabus	Paper
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Question	Answer	Marks
2(a)	lead < nickel < zinc < titanium; (1 mark if one pair reversed)	2
2(b)	positive electrode: oxygen / O ₂ ; negative electrode: aluminium / Al;	2 1 1
2(c)	test: (aqueous) sodium hydroxide / (aqueous) ammonia; result: (grey-) green precipitate / solid;	2 1 1
2(d)(i)	oxygen / air; water;	2 1 1
2(d)(ii)	idea of covering surface with tin / zinc / other suitable metal / plastic / grease / oil / paint / galvanising; prevents oxygen / air or water / moisture / steam from getting to the surface;	2 1 1

Page 4	Mark Scheme	Syllabus	Paper
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Question	Answer	Marks
3(a)	reversible reaction / equilibrium;	1
3(b)	exothermic and products have less energy than reactants;	1
3(c)(i)	percentage yield decreases as temperature increases;	1
3(c)(ii)	91%;	1
3(d)	test: (acidified) potassium manganate(VII)/potassium permanganate; result: (pink solution) turns colourless;	2 1 1
3(e)	any suitable use, e.g. food preservation / manufacture of sulfuric acid;	1
3(f)	sulfur dioxide; (sulfur dioxide) loses oxygen;	2 1 1
3(g)	3 (H ₂ O);	1

Page 5	Mark Scheme	Syllabus	Paper
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Question	Answer	Marks
4(a)	any 2 from: - family/group of similar chemicals; - with same functional group; - trend in physical properties; - same general formula; - same/similar chemical reaction; - successive members differ by CH₂;	2
4(b)(i)	F and G; contain <u>only</u> carbon and hydrogen; have <u>only</u> single bonds/no double bonds;	3 1 1 1
4(b)(ii)	F / methane / CH ₄ ;	1
4(b)(iii)	H; J;	2 1 1
4(b)(iv)	contain oxygen;	1
4(c)(i)	ethanol;	1
4(c)(ii)	yes and because there is a general increase in the numbers/the numbers go up steadily; OR no and because the numbers go down then up again;	1
4(c)(iii)	65°C;	1
4(d)(i)	2 (CO); 3 (H ₂ O);	2 1 1
4(d)(ii)	poisonous/toxic;	1

Page 6	Mark Scheme	Syllabus	Paper
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Question	Answer	Marks
5(a)	liquid; 6°C is higher than the melting point and lower than the boiling point/6 °C is between the melting point and boiling point;	2 1 1
5(b)(i)	potassium chloride; iodine;	2 1 1
5(b)(ii)	<u>iodine</u> is less reactive than <u>bromine</u> / <u>bromine</u> is more reactive than <u>iodine</u> ;	1
5(c)	357 (1 mark for 1 correct row, e.g. $(4 \times 16 =) 64$ or $(2 \times 35.5) = 71$)	2
5(d)(i)	cross shown on baseline;	1
5(d)(ii)	ethanol / other organic solvent;	1
5(d)(iii)	dyes <u>move up</u> the paper and <u>separate</u> ;	1

Page 7	Mark Scheme	Syllabus	Paper
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Question	Answer	Marks
6(a)	any 5 from: • conducts electricity / conducts heat; • soft; • solid; • shiny (when cut); • malleable / ductile; • reacts with water to produce hydrogen; • bubbles / fizzes in water; • vigorous reaction with water; • floats on water / low density; • forms an alkaline solution with water; • reacts with oxygen / air to form an oxide; • reacts with chlorine to form a chloride; • suitable word equations (maximum two equations);	5
6(b)	test: put the sodium compound on <u>nichrome</u> / <u>platinum wire</u> (on the edge of a blue Bunsen burner flame); result: flame goes yellow;	2 1 1
6(c)(i)	pH 13;	1
6(c)(ii)	add (red) litmus to sodium hydroxide / dip (red) litmus into sodium hydroxide; turns blue;	2 1 1
6(d)	sulfur dioxide produced / SO_2 formed; causes breathing difficulties / harmful to eyes / coughing / damages lungs / irritates eyes / sore throat / skin burns / difficulty swallowing / headache / vomiting;	2 1 1

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Question	Answer	Marks
7(a)	flask; (gas) syringe;	2 1 1
7(b)(i)	1.0 (mol/dm ³) because the initial gradient is steeper/initial slope is steeper;	1
7(b)(ii)	steeper gradient than curve for 1.0 mol/dm ³ ; same final volume;	2 1 1
7(c)	any suitable use, e.g. fuel/reducing agent/making margarine/making ammonia/Haber process/fuel cells;	1
7(d)	dust has a (very) high surface area;	1

Page 9	Mark Scheme	Syllabus	Paper
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Question	Answer	Marks
8(a)	<u>mixture</u> of 2 or more metals / <u>mixture</u> of a metal and a non-metal;	1
8(b)	any alloy, e.g. brass, bronze etc.;	1
8(c)	any 4 from: • solder has melted; • atoms in solid (only) vibrate; • atoms in solid are regularly arranged / touching / close to each other; • atoms start to vibrate more; • atoms in liquid are irregularly arranged / close together / touching; • atoms in liquids slide over each other / atoms in liquids move slowly; • atoms move more during melting; • atoms become less regularly arranged during melting;	4
8(d)	vapour <u>spreads</u> everywhere / vapour <u>does not stay in one place</u> ;	1