



Cambridge International Examinations
Cambridge International General Certificate of Secondary Education

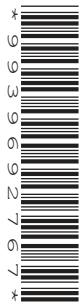
CANDIDATE
NAME

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COMBINED SCIENCE

0653/31

Paper 3 (Extended)

October/November 2016

1 hour 15 minutes

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use an HB pencil for any diagrams or graphs.

Do not use staples, paper clips, glue or correction fluid.

DO NOT WRITE IN ANY BARCODES.

Answer **all** questions.

Electronic calculators may be used.

You may lose marks if you do not show your working or if you do not use appropriate units.

A copy of the Periodic Table is printed on page 24.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part question.

This document consists of **23** printed pages and **1** blank page.



2

- 1 Fig. 1.1 shows a solar-powered pump used to fill a water trough from a well on a farm. Animals in a field on the farm drink from the trough.

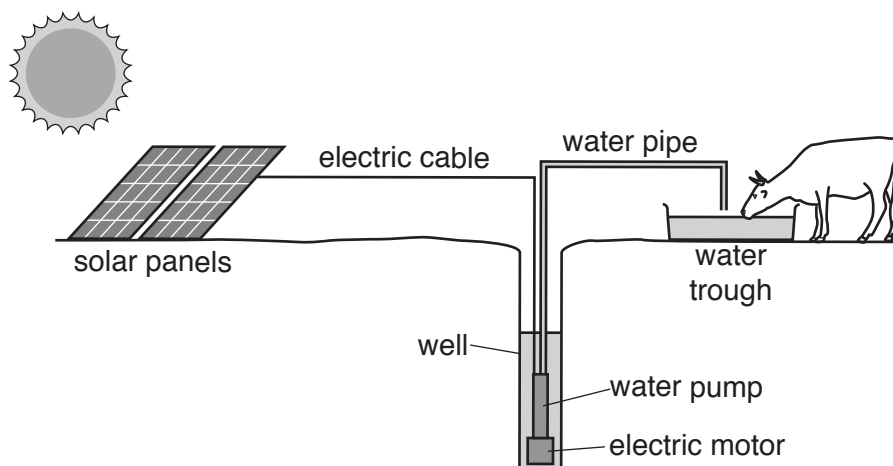


Fig. 1.1

The solar panels use sunlight to generate electricity.
 The motor uses the electricity to drive the pump.
 The pump moves water from the well to the water trough.

- (a) (i) Complete the sequence of energy transformations that takes place as the water pipe is filled when the pump is switched on.

from energy

to **electrical** energy

to energy

to **potential** energy

[2]

- (ii) The pipe from the pump to the trough is 9 m long. When the pump starts, it takes 12 s for water to move from the well to the trough.

Calculate the average speed of the water through the pipe.

speed = m/s [1]

3

- (b) Fig. 1.2 shows a speed/time graph for one water molecule as it moves from the well through the pipe to the trough.

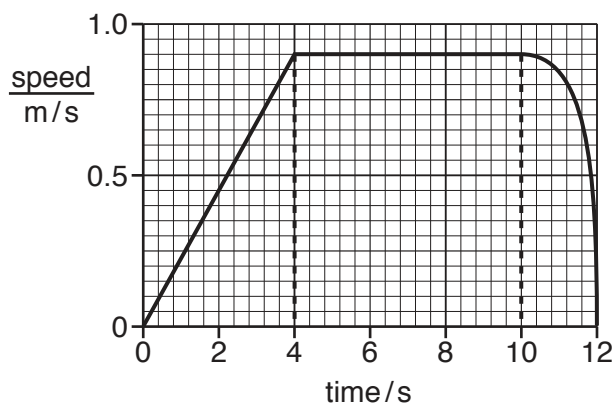


Fig. 1.2

- (i) State the maximum speed of the water molecule through the pipe.

speed = m/s [1]

- (ii) Use Fig. 1.2 to describe how the motion of the molecule between 0 s and 4 s differs from its motion between 10 s and 12 s.

.....

.....

.....

.....[2]

- (c) The water level in the well is 6 m below the water in the trough.

- (i) Calculate the work done in raising 10 kg of water from the well to the trough.
(gravitational field strength, $g = 10 \text{ N/kg}$)

State any formula you use and show your working.

formula

working

work done = J [2]

4

- (ii)** When the pump is working steadily, it takes 2 minutes to pump 10 kg of water from the well into the trough.

Use your answer to **(i)** to calculate the output power of the pump.

State any formula you use and show your working.

formula

working

power = W [2]

Please turn over for Question 2.

6

- 2 An atom of sodium has the symbol shown.



- (a) Deduce the number of neutrons and the number of electrons in an atom of sodium.

neutrons

electrons

[2]

- (b) State the electronic structure of the sodium atom.

.....

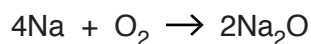
[1]

- (c) Describe the relationship between the electronic structure of the sodium atom and the number of its group in the Periodic Table.

.....

.....[1]

- (d) The equation for the reaction between sodium and oxygen is



- (i) Describe, in terms of electrons, the formation of a sodium ion from a sodium atom.

.....[1]

- (ii) Use the Periodic Table on page 24 to identify the element which has atoms with the same electronic structure as a sodium ion.

.....[1]

7

- (e) All of the Group I metals are solid at 20 °C.

Table 2.1 shows the melting points of some of the Group I metals.

Table 2.1

metal	melting point/°C
Li	
Na	98
K	64
Rb	39
Cs	

Complete Table 2.1 by suggesting the melting points of lithium, Li, and caesium, Cs. [2]

- (f) Solid sodium chloride, NaCl, contains sodium ions, Na⁺, and chloride ions, Cl⁻.

observations

1. Sodium ions cannot be separated from chloride ions in solid sodium chloride by any simple separation method.
2. An aqueous solution of sodium chloride can be separated into sodium chloride and water by the process of evaporation.

Explain these observations.

observation 1

.....

.....

observation 2

.....

.....[2]

8

- 3 Some cress seeds are placed on damp cotton wool in a dish. They are left to germinate in a light-proof box for a few days.

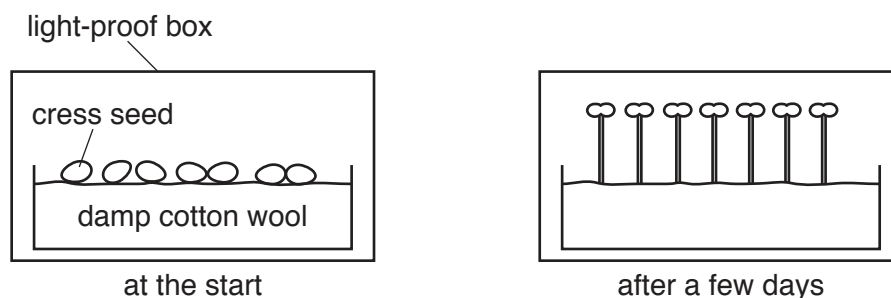


Fig. 3.1

- (a) Fig. 3.1 shows that the seedlings grow directly upwards.

Name this response.

.....[1]

- (b) The dish of seedlings is removed from the box and lit from one side.

Fig. 3.2 shows what happens.

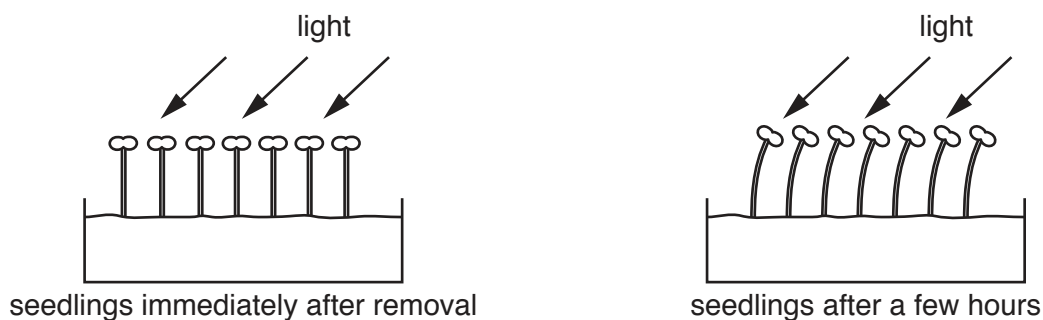


Fig. 3.2

Explain in detail how hormones in the seedlings cause the response shown in Fig. 3.2. You may use a diagram to illustrate your answer.

.....

.....

.....

.....

.....[3]

9

- (c) A student thinks that the germination of seeds might be affected by acid rain.

The student has five dishes each containing 20 seeds. He waters them with solutions that have a range of pH values.

Table 3.1 shows the results after a few days.

Table 3.1

pH of solution	number of seeds that germinate
3	1
4	3
5	3
6	14
7	16

- (i) Describe the effect of pH on the germination of the seeds.

.....

[2]

- (ii) Explain why the germination of seeds is affected by pH.

.....

[1]

- (iii) The pH of acid rain is usually between pH 4 and 5.

Calculate the percentage germination in this pH range.

answer = % [1]

10

(d) (i) Describe how the activities of humans cause acid rain.

.....

.....

.....

.....

.....[2]

(ii) Suggest how acid rain affects the productivity of farms which grow crops.

Explain your answer.

.....

.....[1]

Please turn over for Question 4.

12

- 4 There is a water trough in a field for animals to drink from. On a hot day, the level of the water in the trough drops even when there are no animals drinking from it.

- (a) (i) Explain in terms of the energy and movement of water molecules why the level of water goes down in the trough on a hot day.

.....

.....

.....

.....

.....

.....[3]

- (ii) The thermal energy causing the process in (i) is provided by the Sun.

Suggest one way in which the loss of water from the trough could be reduced without placing a lid on the trough.

.....

.....[1]

- (b) Fig. 4.1 shows the farmer looking at a stick lying in the water trough.

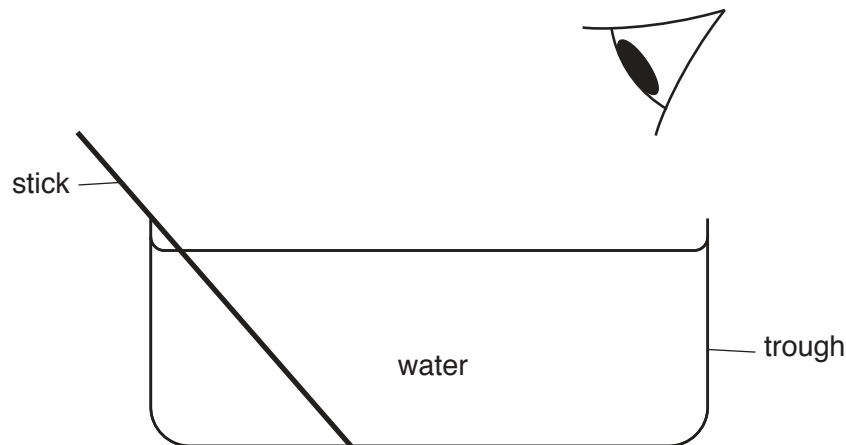


Fig. 4.1

The farmer says that the stick looks bent where it goes into the water.

- (i) Name the property of light that causes the stick to look bent.

.....[1]

- (ii) On Fig. 4.1, draw a ray diagram using two rays from the bottom of the stick to show why the stick appears bent when the farmer looks at it from the side. [3]

13

- (c)** The farmer speaks to someone on his cell phone (mobile phone).

Identify and describe the two different types of wave that enable two people to speak to each other using cell phones.

.....

.....

.....

.....

.....

.....[3]

- 5 (a) Iron is extracted in the blast furnace, as shown in Fig. 5.1.

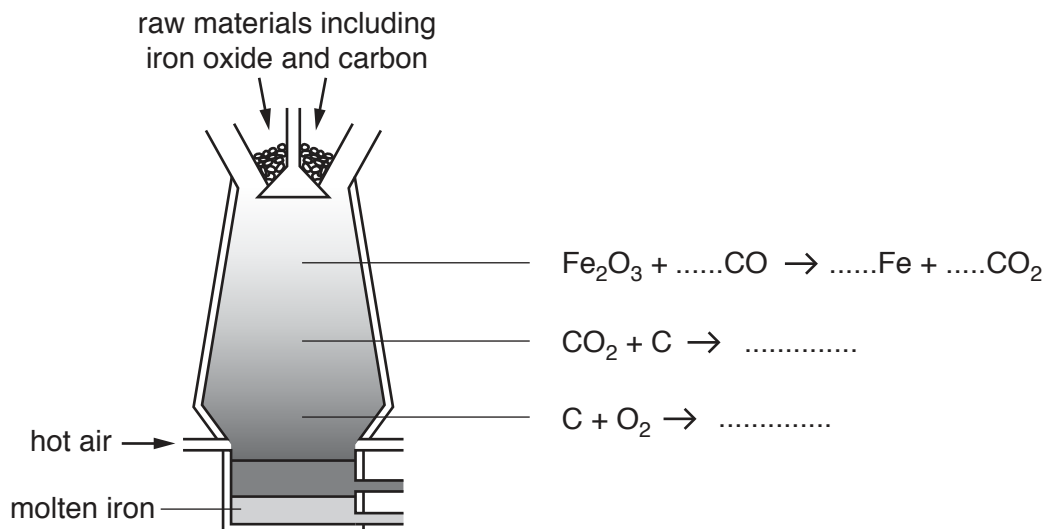


Fig. 5.1

- (i) Complete and balance the equations in Fig. 5.1. [3]

- (ii) The three equations in Fig. 5.1 represent redox reactions.

Identify **one** substance in these equations which is reduced.
Explain your answer briefly.

substance

explanation

.....[2]

- (b) Aluminium is extracted from aluminium oxide, Al_2O_3 , by the process of electrolysis, as shown in Fig. 5.2.

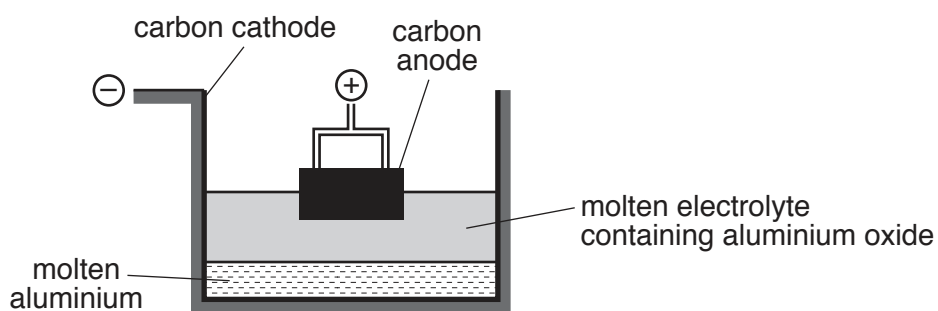


Fig. 5.2

- (i) Suggest why aluminium cannot be extracted from aluminium oxide by heating with carbon in a blast furnace.

.....

.....[1]

15

- (ii) The electrolysis of aluminium oxide does not occur if it is solid.

Explain, in terms of the ions present, why the electrolyte must be molten.

.....
.....
.....[1]

- (iii) Describe briefly, in terms of electrons, what happens to the aluminium ions, Al^{3+} , and the oxide ions, O^{2-} , when aluminium and oxygen are formed during electrolysis.

aluminium ions
.....
oxide ions
.....[1]

- (c) (i) Suggest why aluminium alloys, rather than iron alloys, are used in aircraft construction.

.....[1]

- (ii) Suggest why aluminium alloys, rather than pure aluminium, are used in aircraft construction.

.....[1]

- 6 (a) Microorganisms are added to milk to make yoghurt. They use the nutrients in the milk to make lactic acid. The lactic acid causes the milk to thicken.

Fig. 6.1 shows how the lactic acid production of a microorganism varies with temperature.

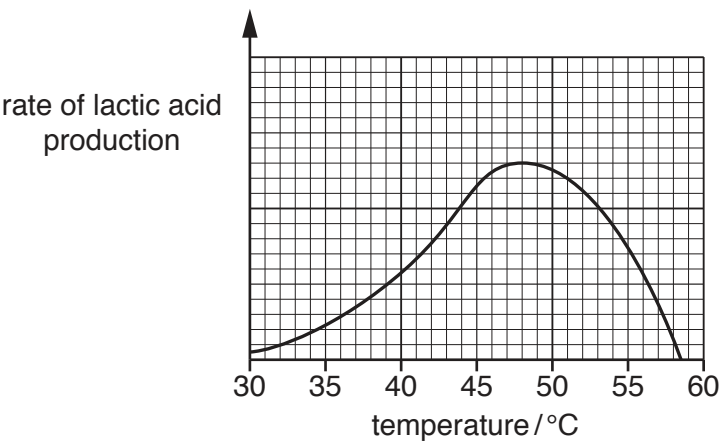


Fig. 6.1

- (i) State the optimum temperature for lactic acid production by this microorganism.
-[1]
- (ii) Describe evidence from Fig. 6.1 that shows that lactic acid production depends on enzyme action in the microorganism.
-
-
-[1]

- (b) Yoghurt can be made from different types of milk.

Table 6.1 shows some of the substances found in two different types of milk.

Table 6.1

substance	milk A/g per 100cm ³	milk B/g per 100cm ³
protein	3.4	7.2
fat	3.6	0.1
carbohydrate	4.7	4.9
calcium (mineral salt)	0.1	0.1
fibre	none	none

17

- (i) Milk **B** contains more protein than milk **A**.

State why the body needs protein.

.....
[1]

- (ii) A balanced diet contains five nutrient groups, as well as fibre and water.

Name the nutrient group, present in milk, which is **not** listed in Table 6.1.

.....
 Describe the importance of including this nutrient group in a balanced diet.

[2]

- (iii) Suggest which type of milk, **A** or **B**, would be a healthy choice for someone who does not wish to develop coronary heart disease.

Explain your answer.

type of milk
 explanation

[1]

- (c) Sometimes fresh fruit is added to yoghurt in the final stages of manufacture. This improves the yoghurt by adding fibre.

Explain the importance of fibre in the diet.

.....
[1]

18

- 7 Fig. 7.1 shows a circuit diagram for a solar-powered electric motor.

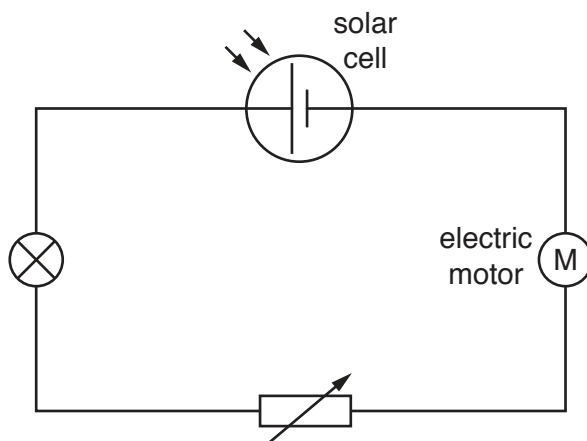


Fig. 7.1

- (a) When the motor is turning, the potential difference (p.d.) across the motor is 1.5 V, across the lamp is 0.5 V, and across the variable resistor is 0.6 V.

- (i) Calculate the p.d. across the solar cell.

p.d. across cell = V [1]

- (ii) The current in the circuit is 0.3 A.

Calculate the resistance of the lamp in the circuit.

State the formula you use and show your working.

formula

working

resistance = Ω [2]

19

- (iii)** Calculate the power output of the solar cell.

State any formula you use and show your working.

formula

working

power = W [2]

- (b)** The purpose of the lamp in the circuit in Fig. 7.1 is to indicate that the motor is working.

Explain why the lamp is connected in series and not in parallel with the motor.

.....

.....[1]

- 8 Fig. 8.1 shows fractional distillation and catalytic cracking at an oil refinery.

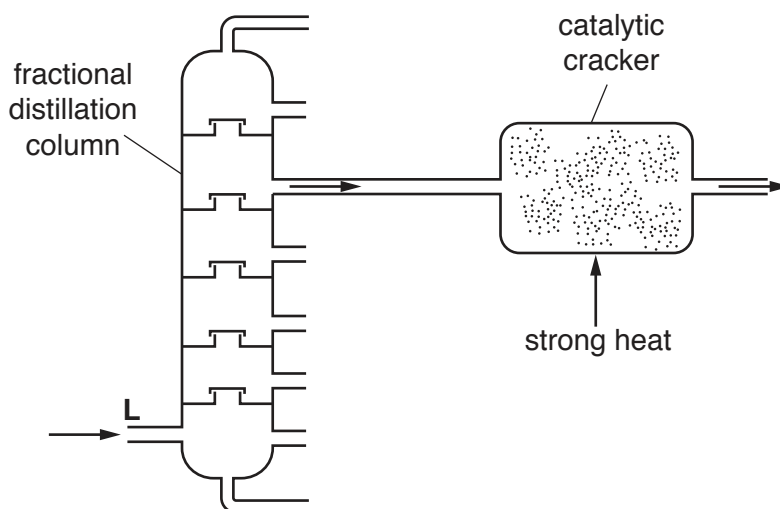


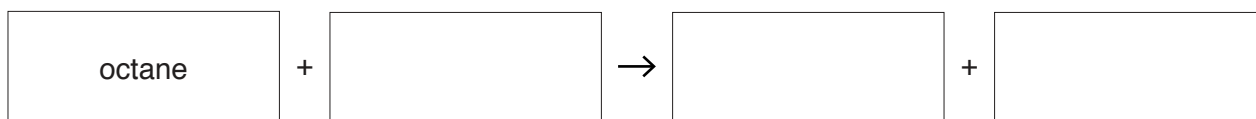
Fig. 8.1

- (a) (i) Name the raw material that enters the fractional distillation column at **L**.
[1]
- (ii) State the physical property which enables the different molecules to be separated during fractional distillation.
[1]
- (b) Cracking molecules of the hydrocarbon $C_{20}H_{42}$ produces a mixture of compounds, as shown below.
- $$C_{20}H_{42} \rightarrow 3C_2H_4 + 2X + C_8H_{18}$$
- (i) Deduce the formula of compound **X**.
[1]
- (ii) Name the compound with the formula C_2H_4 .
[1]
- (iii) Describe a simple chemical test that can distinguish octane, C_8H_{18} , from C_2H_4 .
 Give the results for both compounds.
- test

 C_8H_{18} result.....
 C_2H_4 result[2]

21

(c) Give the word equation for the complete combustion of octane.



[1]

Please turn over for Question 9.

- 9 Fig. 9.1 shows a fetus shortly before it is born.

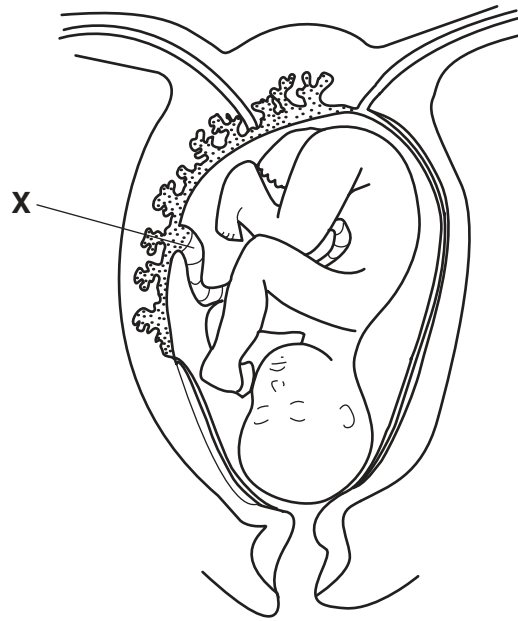


Fig. 9.1

- (a) Complete the sentences below using the words or phrases provided.

You may use each word or phrase once, more than once or not at all.

amnion an artery a vein carbon dioxide
oxygen placenta umbilical cord

Structure **X** is the It contains
..... which carries blood to the
..... to collect for the
growing fetus.

[4]

- (b) (i) Describe **two** ways in which HIV can be transmitted.

1.
2.[2]

- (ii) Describe how the immune system is affected by HIV/AIDS.

.....
.....
.....[2]

The Periodic Table of Elements

Group																												
I	II											III	IV	V	VI	VII	VIII											
		1Hhydrogen1 Key atomic number atomic symbol name relative atomic mass																										
3Li lithium 7	4Be beryllium 9																5B boron 11	6C carbon 12	7N nitrogen 14	8O oxygen 16	9F fluorine 19	2He helium 4						
11Na sodium 23	12Mg magnesium 24																13Al aluminium 27	14Si silicon 28	15P phosphorus 31	16S sulfur 32	17Cl chlorine 35.5	18Ar argon 40						
19K potassium 39	20Ca calcium 40	21Sc scandium 45	22Ti titanium 48	23V vanadium 51	24Cr chromium 52	25Mn manganese 55	26Fe iron 56	27Co cobalt 59	28Ni nickel 59	29Cu copper 64	30Zn zinc 65	31Ga gallium 70	32Ge germanium 73	33As arsenic 75	34Se selenium 79	35Br bromine 80	36Kr krypton 84											
37Rb rubidium 85	38Sr strontium 88	39Y yttrium 89	40Zr zirconium 91	41Nb niobium 93	42Mo molybdenum 96	43Tc technetium —	44Ru ruthenium 101	45Rh rhodium 103	46Pd palladium 106	47Ag silver 108	48Cd cadmium 112	49In indium 115	50Sn tin 119	51Sb antimony 122	52Te tellurium 128	53I iodine 127	54Xe xenon 131											
55Cs caesium 133	56Ba barium 137	57–71 lanthanoids		72Hf hafnium 178	73Ta tantalum 181	74W tungsten 184	75Re rhenium 186	76Os osmium 190	77Ir iridium 192	78Pt platinum 195	79Au gold 197	80Hg mercury 201	81Tl thallium 204	82Pb lead 207	83Bi bismuth 209	84Po polonium —	85At astatine —											
87Fr francium —	88Ra radium —	89–103 actinoids		104Rf rutherfordium —	105Db dubnium —	106Sg seaborgium —	107Bh bohrium —	108Hs hassium —	109Mt meitnerium —	110Ds darmstadtium —	111Rg roentgenium —	112Cn copernicium —	114Fl flerovium —		116Lv livermorium —													
lanthanoids		57La lanthanum 139	58Ce cerium 140	59Pr praseodymium 141	60Nd neodymium 144	61Pm promethium —	62Sm samarium 150	63Eu europium 152	64Gd gadolinium 157	65Tb terbium 159	66Dy dysprosium 163	67Ho holmium 165	68Er erbium 167	69Tm thulium 169	70Yb ytterbium 173	71Lu lutetium 175												
actinoids		89Ac actinium —	90Th thorium 232	91Pa protactinium 231	92U uranium 238	93Np neptunium —	94Pu plutonium —	95Am americium —	96Cm curium —	97Bk berkelium —	98Cf californium —	99Es einsteinium —	100Fm fermium —	101Md mendelevium —	102No nobelium —	103Lr lawrencium —												

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.)