

Write your name here

Surname	Other names
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**Pearson Edexcel**  
International  
Advanced Level

Centre Number	Candidate Number												
<table border="1" style="width: 100%; height: 20px; border-collapse: collapse;"> <tr> <td style="width: 12.5%;"></td> </tr> </table>							<table border="1" style="width: 100%; height: 20px; border-collapse: collapse;"> <tr> <td style="width: 12.5%;"></td> </tr> </table>						

# Chemistry

## Advanced

### Unit 6: Chemistry Laboratory Skills II

Thursday 26 January 2017 – Afternoon <b>Time: 1 hour 15 minutes</b>	Paper Reference <b>WCH06/01</b>
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Candidates may use a calculator.	Total Marks
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### Instructions

- Use **black** ink or **black** ball-point pen.
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided  
– *there may be more space than you need.*

### Information

- The total mark for this paper is 50.
- The marks for **each** question are shown in brackets  
– *use this as a guide as to how much time to spend on each question.*
- You will be assessed on your ability to organise and present information, ideas, descriptions and arguments clearly and logically, including your use of grammar, punctuation and spelling.
- A Periodic Table is printed on the back cover of this paper.

### Advice

- Read each question carefully before you start to answer it.
- Keep an eye on the time.
- Try to answer every question.
- Check your answers if you have time at the end.

Turn over ►

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**Pearson**

**Answer ALL the questions. Write your answers in the spaces provided.**

**1** The inorganic compound **A** has two cations and one anion. Complete the table.

	Test	Observation	Inference	
(a)	Note the appearance of solid <b>A</b> .	Pale green crystals.	One of the elements present in <b>A</b> must be found in the part of the Periodic Table called ..... .....	(1)
(b)	Mix a spatula measure of solid <b>A</b> , with 20 cm <sup>3</sup> of distilled water.	Solid <b>A</b> dissolved easily to form a very pale green solution.	The <b>formula</b> of the cation responsible for the pale green colour is likely to be .....	(1)
(c)	To about 5 cm <sup>3</sup> of the solution of <b>A</b> , add 1 cm <sup>3</sup> of dilute sodium hydroxide.	A green precipitate formed which turned brown on standing.	The green precipitate is ..... The brown solid is .....	(2)
(d)	Gently warm the mixture obtained in (c) and test the gas evolved.  Test the gas evolved by ..... .....	Result of test. .....	The gas evolved is ammonia.  Therefore the <b>formula</b> of the second cation present in <b>A</b> is .....	(3)
(e)	To about 5 cm <sup>3</sup> of a fresh sample of the solution of <b>A</b> , add a few drops of ..... ..... .....  and then ..... .....	..... .....	The anion present is a sulfate.	(3)

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(f) What is the **formula** of compound **A**?

(1)

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(Total for Question 1 = 11 marks)

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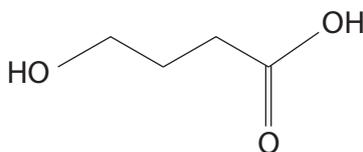
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- 2 4-hydroxybutanoic acid is a naturally occurring substance found in the human central nervous system, as well as in wine, beef and citrus fruits. 4-hydroxybutanoic acid is a sticky solid with a melting temperature of 48–50 °C. The structure of 4-hydroxybutanoic acid is



- (a) About 5 g of 4-hydroxybutanoic acid was placed in a test tube and melted by heating in a hot water bath. A small piece of sodium was added to the molten 4-hydroxybutanoic acid.
- (i) Suggest why the flammability of 4-hydroxybutanoic acid would **not** present a significant risk in this experiment, even if a Bunsen burner were used.

(1)

- (ii) Suggest **one** reason why a water bath was used in this experiment rather than a Bunsen burner.

(1)

- (iii) Describe the observation that shows a reaction is occurring.

(1)

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(iv) Give the equation for this reaction. Use a **skeletal** formula for the organic product. State symbols are not required.

(2)

(b) A spatula measure of 4-hydroxybutanoic acid was added to a solution of sodium hydrogencarbonate in a test tube. Vigorous effervescence was observed.

(i) Name the gas formed.

(1)

(ii) State the test that you would use to confirm the identity of the gas formed. Give the result of the test.

(2)

(iii) Draw a diagram of the apparatus you would use to carry out the test that you have given in 2(b)(ii).

(2)

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(c) Identifying the functional groups present is a vital part of the determination of the structure of an organic compound.

- (i) Explain fully whether or not the tests mentioned in 2(a) and 2(b) allow the two functional groups of 4-hydroxybutanoic acid to be identified.

(2)

- (ii) Use the data in the table to explain how the infrared spectrum of 4-hydroxybutanoic acid could be used to identify its two functional groups. State the relevant wavenumber ranges.

(2)

Group & vibration type	Type of compound	Wavenumber range / $\text{cm}^{-1}$
O—H stretching	Alcohols and phenols	3750–3200
	Carboxylic acids	3300–2500
C—O stretching	Esters: methanoates	1200–1180
	ethanoates	1250–1230
	propanoates	1200–1150
	benzoates	1310–1250 and 1150–1100
C=O stretching	Aldehydes	1740–1720
	Ketones	1700–1680
	Carboxylic acids: alkyl aryl	1725–1700 1700–1680
	Esters	1750–1735



(iii) Explain whether or not infrared spectra could be used to distinguish 4-hydroxybutanoic acid from its isomer, 3-hydroxybutanoic acid.

(1)

.....

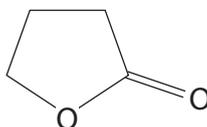
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(d) 4-hydroxybutanoic acid reacts on heating in the presence of a catalyst to form 4-butyrolactone which is used in food flavouring, and as a superglue remover. The structure of 4-butyrolactone is



(i) By considering the functional group present in 4-butyrolactone, name the type of reaction which occurs when 4-hydroxybutanoic acid forms 4-butyrolactone.

(1)

.....

(ii) Identify a suitable catalyst for this reaction.

(1)

.....

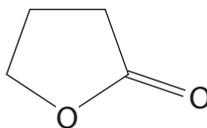


(iii) There are three proton environments in 4-butyrolactone.

Circle each of these proton environments on the structure below and label them X, Y and Z.

Complete the table to give the relative areas under the three peaks and the splitting pattern expected for each peak in the high resolution proton nmr spectrum of this compound.

(3)



Proton environment	Relative peak height	Splitting pattern
X		
Y		
Z		

(Total for Question 2 = 20 marks)

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- 3 The first stage in a preparation of an azo dye is the formation of benzenediazonium chloride by the reaction of phenylamine (aniline) with nitrous acid at about 5 °C. The equation for the reaction is



- (a) (i) Nitrous acid is unstable at room temperature, so it is prepared only when it is needed. Identify the **two** reagents required to prepare nitrous acid for this reaction. (2)

- (ii) State how a temperature of 'about 5 °C' would be maintained for this preparation. (1)

- (b) If the benzenediazonium chloride is allowed to warm up, it decomposes:



One way of following the progress of this reaction in a rates experiment is to use a continuous monitoring method, such as measuring the volume of nitrogen produced.

- (i) Suggest an advantage of using a continuous monitoring method rather than a sampling method. (1)

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- (ii) Draw a diagram of an apparatus suitable for measuring the volume of nitrogen produced at various times during the decomposition of benzenediazonium chloride at 45 °C.

(3)

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P 4 8 3 6 9 A 0 1 1 1 6

- (c) Data from the experiment described in 3(b)(ii) are given in the table. The initial concentration of the benzenediazonium chloride was  $0.071 \text{ mol dm}^{-3}$  and the concentrations of the benzenediazonium chloride as the reaction proceeded were calculated from the volumes of nitrogen collected.

Time / min	Volume nitrogen / $\text{cm}^3$	$1000 \times [\text{C}_6\text{H}_5\text{N}_2\text{Cl}] / \text{mol dm}^{-3}$
0	0	71.0
3	10.8	57.8
6	19.3	
9	26.3	39.0
12	32.4	31.5
18	41.3	20.7
24	46.5	14.4
30	50.4	9.6
$\infty$	58.3	0

$\infty$  = time when reaction is complete.

- (i) How would you know when the reaction was complete?

(1)

- (ii) After a given time, the fraction of nitrogen produced is equal to the fraction of the benzenediazonium chloride that is used up. For example, when  $\frac{1}{4}$  of the nitrogen is produced,  $\frac{3}{4}$  of the benzenediazonium chloride is left. This is used to calculate the concentrations of the benzenediazonium chloride in the table. Calculate the concentration of benzenediazonium chloride at 6 min. You must show your working.

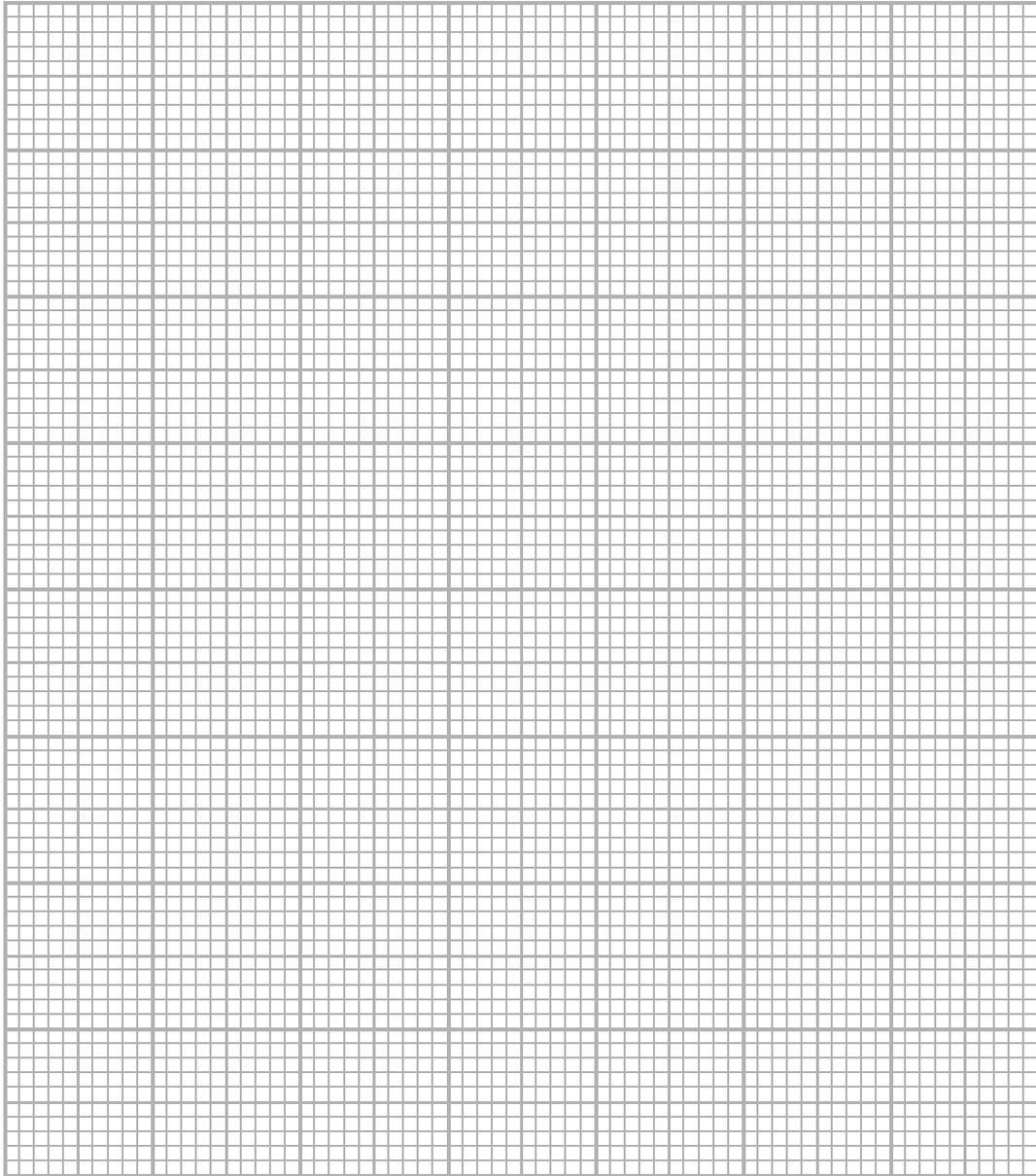
(1)

$[\text{C}_6\text{H}_5\text{N}_2\text{Cl}]$  at 6 min = .....  $\text{mol dm}^{-3}$



- (iii) Plot a graph of  $1000 \times [\text{C}_6\text{H}_5\text{N}_2\text{Cl}]$  (on the vertical axis) against time.  
Use appropriate scales, and label the axes of the graph.

(3)



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P 4 8 3 6 9 A 0 1 3 1 6

(iv) From the graph, estimate **two** half lives of this reaction. Show your working on the graph.

(1)

(v) Use the half lives that you have estimated to deduce the order of the reaction. Justify your answer.

(1)

(d) Benzenediazonium chloride can be converted into an azo dye by mixing it with phenol in an alkaline solution. Pure crystals of the dye may be obtained by recrystallisation of the impure solid using ethanoic acid as the solvent.

(i) Phenol is an extremely caustic compound. Suggest **one** safety precaution, apart from eye protection and wearing a laboratory coat, that would reduce the risk of using phenol.

(1)

(ii) In the first stage of recrystallisation, the impure solid is dissolved in the minimum volume of hot solvent. Explain why it is important to use the **minimum** volume.

(1)

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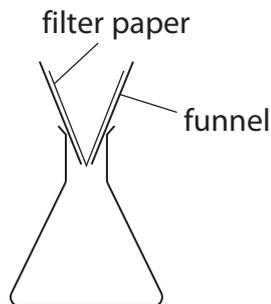
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(iii) In the second stage of recrystallisation, the hot solution is filtered using simple (or gravity) filtration. A filter funnel without a stem is often used (see diagram). Explain the advantage of using a funnel like this.

(1)



(iv) When crystallisation is complete, the mixture is filtered under reduced pressure (vacuum filtration). State the type of impurities which are removed at this stage and explain why this method of filtration is preferred to gravity filtration.

(2)

**(Total for Question 3 = 19 marks)**

**TOTAL FOR PAPER = 50 MARKS**

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# The Periodic Table of Elements

	1	2											3	4	5	6	7	0 (8)		
			(1)	(2)											(13)	(14)	(15)	(16)	(17)	(18)
	6.9	9.0	6.9	9.0											10.8	12.0	14.0	16.0	19.0	4.0
	Li	Be	Li	Be											B	C	N	O	F	He
	lithium	beryllium	lithium	beryllium											boron	carbon	nitrogen	oxygen	fluorine	helium
	3	4	3	4											5	6	7	8	9	2
	23.0	24.3	23.0	24.3											27.0	28.1	31.0	32.1	35.5	20.2
	Na	Mg	Na	Mg											Al	Si	P	S	Cl	Ne
	sodium	magnesium	sodium	magnesium											aluminium	silicon	phosphorus	sulfur	chlorine	neon
	11	12	11	12											13	14	15	16	17	10
	39.1	40.1	39.1	40.1											69.7	72.6	74.9	79.0	79.9	83.8
	K	Ca	K	Ca											Ga	Ge	As	Se	Br	Kr
	potassium	calcium	potassium	calcium											gallium	germanium	arsenic	selenium	bromine	krypton
	19	20	19	20											31	32	33	34	35	36
	85.5	87.6	85.5	87.6											114.8	118.7	121.8	127.6	126.9	131.3
	Rb	Sr	Rb	Sr											In	Sn	Sb	Te	I	Xe
	rubidium	strontium	rubidium	strontium											indium	tin	antimony	tellurium	iodine	xenon
	37	38	37	38											49	50	51	52	53	54
	132.9	137.3	132.9	137.3											204.4	207.2	209.0	[209]	[210]	[222]
	Cs	Ba	Cs	Ba											Tl	Pb	Bi	Po	At	Rn
	caesium	barium	caesium	barium											thallium	lead	bismuth	polonium	astatine	radon
	55	56	55	56											81	82	83	84	85	86
	[223]	[226]	[223]	[226]											200.6	204.4	209.0	[209]	[210]	[222]
	Fr	Ra	Fr	Ra											Hg	Tl	Pb	Bi	Po	At
	francium	radium	francium	radium											mercury	thallium	lead	bismuth	polonium	astatine
	87	88	87	88											80	81	82	83	84	86
	[227]	[227]	[227]	[227]											106.4	106.4	106.4	106.4	106.4	106.4
	Ac*	Ra	Ac*	Ra											Pt	Au	Hg	Tl	Pb	Bi
	actinium	radium	actinium	radium											platinum	gold	mercury	thallium	lead	bismuth
	89	88	89	88											78	79	80	81	82	83
	[261]	[261]	[261]	[261]											106.4	106.4	106.4	106.4	106.4	106.4
	Rf	Rf	Rf	Rf											Pd	Ag	Cd	In	Sn	Sb
	rutherfordium	rutherfordium	rutherfordium	rutherfordium											palladium	silver	cadmium	indium	tin	antimony
	104	104	104	104											46	47	48	49	50	51
	[262]	[262]	[262]	[262]											106.4	106.4	106.4	106.4	106.4	106.4
	Db	Db	Db	Db											Pd	Ag	Cd	In	Sn	Sb
	dubnium	dubnium	dubnium	dubnium											palladium	silver	cadmium	indium	tin	antimony
	105	105	105	105											46	47	48	49	50	51
	[264]	[264]	[264]	[264]											106.4	106.4	106.4	106.4	106.4	106.4
	Bh	Bh	Bh	Bh											Pd	Ag	Cd	In	Sn	Sb
	bohrium	bohrium	bohrium	bohrium											palladium	silver	cadmium	indium	tin	antimony
	107	107	107	107											46	47	48	49	50	51
	[266]	[266]	[266]	[266]											106.4	106.4	106.4	106.4	106.4	106.4
	Sg	Sg	Sg	Sg											Pd	Ag	Cd	In	Sn	Sb
	seaborgium	seaborgium	seaborgium	seaborgium											palladium	silver	cadmium	indium	tin	antimony
	106	106	106	106											46	47	48	49	50	51
	[147]	[147]	[147]	[147]											106.4	106.4	106.4	106.4	106.4	106.4
	Pm	Pm	Pm	Pm											Pd	Ag	Cd	In	Sn	Sb
	promethium	promethium	promethium	promethium											palladium	silver	cadmium	indium	tin	antimony
	61	61	61	61											46	47	48	49	50	51
	[237]	[237]	[237]	[237]											106.4	106.4	106.4	106.4	106.4	106.4
	Np	Np	Np	Np											Pd	Ag	Cd	In	Sn	Sb
	neptunium	neptunium	neptunium	neptunium											palladium	silver	cadmium	indium	tin	antimony
	93	93	93	93											46	47	48	49	50	51
	[231]	[231]	[231]	[231]											106.4	106.4	106.4	106.4	106.4	106.4
	Pa	Pa	Pa	Pa											Pd	Ag	Cd	In	Sn	Sb
	protactinium	protactinium	protactinium	protactinium											palladium	silver	cadmium	indium	tin	antimony
	91	91	91	91											46	47	48	49	50	51
	[238]	[238]	[238]	[238]											106.4	106.4	106.4	106.4	106.4	106.4
	U	U	U	U											Pd	Ag	Cd	In	Sn	Sb
	uranium	uranium	uranium	uranium											palladium	silver	cadmium	indium	tin	antimony
	92	92	92	92											46	47	48	49	50	51
	[242]	[242]	[242]	[242]											106.4	106.4	106.4	106.4	106.4	106.4
	Pu	Pu	Pu	Pu											Pd	Ag	Cd	In	Sn	Sb
	plutonium	plutonium	plutonium	plutonium											palladium	silver	cadmium	indium	tin	antimony
	94	94	94	94											46	47	48	49	50	51
	[243]	[243]	[243]	[243]											106.4	106.4	106.4	106.4	106.4	106.4
	Am	Am	Am	Am											Pd	Ag	Cd	In	Sn	Sb
	americium	americium	americium	americium											palladium	silver	cadmium	indium	tin	antimony
	95	95	95	95											46	47	48	49	50	51
	[247]	[247]	[247]	[247]											106.4	106.4	106.4	106.4	106.4	106.4
	Cm	Cm	Cm	Cm											Pd	Ag	Cd	In	Sn	Sb
	curium	curium	curium	curium											palladium	silver	cadmium	indium	tin	antimony
	96	96	96	96											46	47	48	49	50	51
	[251]	[251]	[251]	[251]											106.4	106.4	106.4	106.4	106.4	106.4
	Cf	Cf	Cf	Cf											Pd	Ag	Cd	In	Sn	Sb
	californium	californium	californium	californium											palladium	silver	cadmium	indium	tin	antimony
	98	98	98	98											46	47	48	49	50	51
	[254]	[254]	[254]	[254]											106.4	106.4	106.4	106.4	106.4	106.4
	Es	Es	Es	Es											Pd	Ag	Cd	In	Sn	Sb
	einsteinium	einsteinium	einsteinium	einsteinium											palladium	silver	cadmium	indium	tin	antimony
	99	99	99	99											46	47	48	49	50	51
	[253]	[253]	[253]	[253]											106.4	106.4	106.4	106.4	106.4	106.4
	Fm	Fm	Fm	Fm											Pd	Ag	Cd	In	Sn	Sb
	fermium	fermium	fermium	fermium											palladium	silver	cadmium	indium	tin	antimony
	100	100	100	100											46	47	48	49	50	51
	[256]	[256]	[256]	[256]											106.4	106.4	106.4	106.4	106.4	106.4
	Md	Md	Md	Md											Pd	Ag	Cd	In	Sn	Sb
	mendeleevium	mendeleevium	mendeleevium	mendeleevium											palladium	silver	cadmium	indium	tin	antimony
	101	101	101	101											46	47	48	49	50	51
	[257]	[257]	[257]	[257]											106.4	106.4	106.4	106.4	106.4	106.4
	Lr	Lr	Lr	Lr											Pd	Ag	Cd	In	Sn	Sb
	lawrencium	lawrencium	lawrencium	lawrencium											palladium	silver	cadmium	indium	tin	antimony
	103	103	103	103											46	47	48	49	50	51

Elements with atomic numbers 112-116 have been reported but not fully authenticated

\* Lanthanide series  
\* Actinide series

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