

# Mark Scheme (Results)

June 2011

International GCSE

Chemistry (4CH0)  
Paper 2C

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## General Marking Guidance

- All candidates must receive the same treatment. Examiners must mark the first candidate in exactly the same way as they mark the last.
- Mark schemes should be applied positively. Candidates must be rewarded for what they have shown they can do rather than penalised for omissions.
- Examiners should mark according to the mark scheme not according to their perception of where the grade boundaries may lie.
- There is no ceiling on achievement. All marks on the mark scheme should be used appropriately.
- All the marks on the mark scheme are designed to be awarded. Examiners should always award full marks if deserved, i.e. if the answer matches the mark scheme. Examiners should also be prepared to award zero marks if the candidate's response is not worthy of credit according to the mark scheme.
- Where some judgement is required, mark schemes will provide the principles by which marks will be awarded and exemplification may be limited.
- When examiners are in doubt regarding the application of the mark scheme to a candidate's response, the team leader must be consulted.
- Crossed out work should be marked UNLESS the candidate has replaced it with an alternative response.

**INTERNATIONAL GCSE CHEMISTRY 4CH0/2C – SUMMER 2011**

Question number	Expected Answer	Accept	Reject	Marks
1	<ul style="list-style-type: none"> <li>• Fizzing occurs (box 2)</li> <li>• potassium moves around (box 4)</li> <li>• potassium melts (box 5)</li> <li>• a lilac flame is seen (box 7)</li> </ul> <p>[If more than four boxes are ticked, deduct a mark for each incorrect answer above four]</p>			1  1  1  1

**Total 4 Marks**

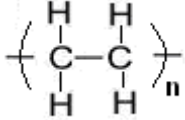
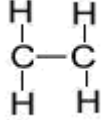
Question number	Expected Answer	Accept	Reject	Marks
2 (a)	iron			1
(b)	haematite			1
(c)	ammonia			1
(d)	sodium hydroxide			1
(e)	ammonia			1

**Total 5 Marks**

Question number	Expected answer	Accept	Reject	Marks
3 (a)(i)	Magnesium	Mg		1
(a)(ii)	It would react with the sulfuric acid / the clouds / the atmosphere / it will fizz	It is <u>too</u> reactive / <u>very</u> reactive / the <u>most</u> reactive  Dissolve <u>in the</u> (sulfuric) <u>acid/ eq</u>	“reactive” by itself	1
(b)	it has low melting point / would melt / temperature on Venus is higher than the melting point of lead (ora) IGNORE heavy / dense	Lead would be a liquid Answer using data from table e.g. 328°C is lower than temp on Venus		1
(c)	Titanium  Any two from: <ul style="list-style-type: none"> <li>it has a low density / is lightweight</li> <li>it has a high melting point / wouldn't melt / temperature on Venus is lower than the melting point of titanium (ora)</li> <li>does not react with sulfuric acid / the clouds / the atmosphere / it will not fizz</li> </ul>	Ti  The <u>probe</u> would be light  Remains solid  <b>Reason marks can be scored for copper (density mark would need to be compared to lead)</b>	Light on its own / light <u>in</u> weight on its own	1  1  1

Total 6 Marks

Question number	Expected answer	Accept	Reject	Marks
4 (a)	(i) Contains a (carbon to carbon) double bond / contains C=C / multiple bond IGNORE references to 'free' bond / spare bond	Can undergo addition reactions / does not contain the maximum number of hydrogen (atoms)		1
	(ii) (add) bromine (water) / Br <sub>2</sub> IGNORE references to any other solvent  decolourised / turns (from orange/brown to) colourless IGNORE starting colour IGNORE clear IGNORE discolour  2 <sup>nd</sup> mark dependent on 1 <sup>st</sup> mark having been awarded, but for near miss of reagent, e.g. bromine in presence of uv, observation mark can be awarded  Ignore references to any products, correctly named or otherwise	KMnO <sub>4</sub> / potassium (per)manganate (VII) with either an acid or an alkali  (purple to) colourless (if acid used) (purple to) green (if alkali used)		1  1

4 (b) (i)	H <sub>2</sub> O			1
(ii)	Dehydration	Elimination		1
(c)	 <p>1 mark for  i.e. double to single</p> <p>1 mark for rest of formula, including extension lines, brackets and the 'n'</p>	<p>CH<sub>2</sub> - CH<sub>2</sub></p> <p>n as superscript</p> <p>Max 1 for skeletal formula</p>	<p>Any double-bonded product scores 0/2</p> <p>n before the brackets</p>	2

**Total 7 Marks**

Question number	Expected answer	Accept	Reject	Marks
5 (a)	Any two from magnesium chloride, calcium chloride, iron(II) chloride  If more than 2 given, deduct 1 mark for each incorrect answer If name and formula given, both must be correct	MgCl <sub>2</sub> / CaCl <sub>2</sub> / FeCl <sub>2</sub> / ferrous chloride / iron chloride	Iron(III) chloride / ferric chloride	2
(b)	To check the reliability / repeatability (of the results/method)  IGNORE references to obtaining a mean / average IGNORE references to identifying anomalous results	To get concordant / consistent / precise results To increase the accuracy (of the results)	To increase the validity / to make it a fair test To increase the accuracy <u>of the method</u>	1
(c) (i)	4.30 circled			1
(ii)	Repeat the experiment IGNORE refs to validity / fair test	Discard / ignore this result Work out average only using other 2 results		1

5 (d)	Burette	Minor spelling mistakes, e.g. burrete / burete	Biuret	1
(e)	$\frac{(1.60 + 1.70 + 1.65)}{3}$ $= 1.65$ <p>(2<sup>nd</sup> mark consequential on 1st and can be awarded if small slip in data)</p> <p>Correct answer on its own scores 2</p>	<p>Award 1 mark for a correct average of the data for any other row except iron (II) chloride or water</p> <p>Award 1 mark for use of 1.90 and 1.95 to give 1.93</p>	<p>Any answer not to 2 d.p. (loses 2<sup>nd</sup> mark only)</p>	<p>1</p> <p>1</p>

**Total 8 Marks**

Question number	Expected answer	Accept	Reject	Marks
6 (a)	<p>Giant (structure / lattice / atomic / molecular)</p> <p>Covalent</p> <p>Idea that (covalent) bonds are broken</p> <p>(Covalent bonds) are strong / many bonds (are broken) / lots of {energy/heat} required</p> <p><b>NB No penalty for referring to graphite</b></p>	Macromolecular	<p>Max 2 for mentioning of ionic or metallic bonding or Intermolecular forces</p> <p>Bonds loosened</p>	<p>1</p> <p>1</p> <p>1</p> <p>1</p>
(b)	<p>Layers slide / slip / move over each other</p> <p>IGNORE particles in layers such as atoms, but REJECT if ions / molecules / electrons for first mark only</p> <p>Weak (intermolecular forces of) <u>attraction</u> between layers / weak van der Waals (forces of attraction) between layers</p> <p>IGNORE references to bonds <u>within</u> the layers</p>	<p>Sheets / planes slide</p> <p>Any indication that the forces are those of attraction, e.g. forces overcome / forces are broken / forces hold the layers together</p>	<p>Rows slide</p> <p>Any reference to <u>bonds</u> between layers / molecules</p>	<p>1</p> <p>1</p>
(c)	<p>Delocalised electrons</p> <p>(which) move / migrate / flow</p> <p>IGNORE references to "carrying" charge / current</p>		<p>Refs to atoms / ions / molecules scores 0/2</p>	<p>1</p> <p>1</p>

6 (d)	<p>Any two from:</p> <p>Not a giant structure IGNORE simple molecular</p> <p>Weak intermolecular (forces of ) <u>attraction</u> / weak (forces of) <u>attraction</u> between molecules / weak van der Waals (forces of attraction) between molecules</p> <p>No covalent bonds break (when melting)</p>	<p>Smaller molecules / simpler structure than diamond</p> <p>Any indication that the forces are those of attraction, e.g. forces <u>overcome</u> / forces <u>broken</u> / forces <u>hold</u> the molecules together</p> <p>First and third marking points can be awarded for correct comparisons between the two structures, e.g. buckminsterfullerene is simple molecular whereas diamond is giant covalent scores the first mark; weak intermolecular forces of attraction in buckminsterfullerene are broken as opposed to the covalent bonds in diamond (scores the 3<sup>rd</sup> mark, as well as the 2nd)</p>	<p>MAX 1 for any mention of covalent bonds are broken in Buckminster fullerene</p> <p>Any reference to <u>bonds</u> between molecules</p>	2
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Question number	Expected answer	Accept	Reject	Marks
7 (a) (i)	108/24 = 4.5	1 mark for answer of 4.8(2) (molar volume = 22.4dm <sup>3</sup> )		1
(ii)	M <sub>r</sub> of NaN <sub>3</sub> = 65 Moles of NaN <sub>3</sub> = 3 OR two thirds of (a)(i) Mass of NaN <sub>3</sub> = 195 (g) OR moles of NaN <sub>3</sub> x M <sub>r</sub> [Mark consequentially at each stage]	23 + (14 x 3)  Correct answer with no working scores 3		1  1  1
(b) (i)	Removes (harmful) sodium	Produces <u>more</u> nitrogen / gas OR bag inflates more quickly		1
(ii)	K <sub>2</sub> O(s) + SiO <sub>2</sub> (s) → K <sub>2</sub> SiO <sub>3</sub> (s) OR K <sub>2</sub> O(s) + SiO <sub>2</sub> (s) → K <sub>2</sub> SiO <sub>3</sub> (l) IGNORE same numbers of Na <sub>2</sub> O on both sides of equation			1
(c) (i)	Precipitation	<u>Double</u> decomposition	Double displacement	1
(ii)	Filtration / filter IGNORE refs to adding water	Decanting / pour off liquid	Sieving / evaporation / distillation / crystallisation / heat	1

**Total 9 Marks**

Question number	Expected answer	Accept	Reject	Marks
8 (a)	It (like water) is a colourless (liquid)  IGNORE it is clear / transparent IGNORE references to smell	it looks the same		1
(b)	(Sulfuric acid / it) contains water	Aqueous (Copper sulfate) becomes hydrated	Contains hydrogen and oxygen / the elements of water	1
(c)	Pressure: > 1 but ≤ 5 (atm)  Temperature: 350 to 550 (°C)  Catalyst: vanadium(V) oxide	Any range within this range, including 1 – 2 atm  Any range within this range  Vanadium pentoxide / V <sub>2</sub> O <sub>5</sub> / vanadium oxide / vanadium(5) oxide	Values in alternative units  Values in alternative units  other oxidation states	1  1  1

8	(d)	(i)	$\frac{30(.00) \times 0.2(00)}{1000}$  $= 0.006(00)$	Correct answer with no working scores 2		1
		(ii)	$0.003(00)$ OR $(d)(i) \div 2$	6 for 1 mark only (i.e. not dividing by 1000)		1
		(iii)	$0.003(00) \times 40$ OR $(d)(ii) \times 40$  $= 0.12(0)$	Correct answer with no working scores 2 Award 1 mark for dividing d(ii) by 25 and correctly evaluating		1
		(iv)	$0.12(0) \times 100 = 12(.0)$ OR $(d)(iii) \times 100$			1

**Total 11 Marks**  
**PAPER TOTAL: 60 MARKS**



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