



Mark Scheme (Results)

January 2018

Pearson Edexcel International GCSE
In Chemistry (4CH0) Paper 2C



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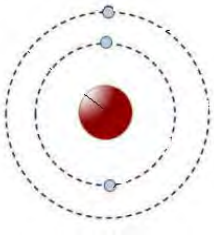
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General Marking Guidance

- All candidates must receive the same treatment. Examiners must mark the first candidate in exactly the same way as they mark the last.
- Mark schemes should be applied positively. Candidates must be rewarded for what they have shown they can do rather than penalised for omissions.
- Examiners should mark according to the mark scheme not according to their perception of where the grade boundaries may lie.
- There is no ceiling on achievement. All marks on the mark scheme should be used appropriately.
- All the marks on the mark scheme are designed to be awarded. Examiners should always award full marks if deserved, i.e. if the answer matches the mark scheme. Examiners should also be prepared to award zero marks if the candidate's response is not worthy of credit according to the mark scheme.
- Where some judgement is required, mark schemes will provide the principles by which marks will be awarded and exemplification may be limited.
- When examiners are in doubt regarding the application of the mark scheme to a candidate's response, the team leader must be consulted.
- Crossed out work should be marked UNLESS the candidate has replaced it with an alternative response.

Question number	Answer	Notes	Marks
1 (a)	P beaker Q (filter) funnel R conical flask S pipette		4
(b)	D red		1
(c)	A blue		1

Total 6 marks

Question number	Answer	Notes	Marks
2 (a)	electron(s)		1
(b)	nucleus		1
(c)	M1 proton(s) M2 neutron(s)	in either order	1 1
(d) (i)	isotope(s)		1
(ii)		ALLOW other symbols for electrons do not need to show nucleus REJECT any diagram showing a lithium ion	1
(e)	M1 $[(92.5 \times 7) + (7.5 \times 6)]/100$ M2 = 6.9	ACCEPT 6.925 and 6.93 REJECT 7 Incorrect rounding e.g. 6.92 scores 1 IGNORE units	2

Total 8 marks

Question number	Answer	Notes	Marks						
3 (a)	B neutralisation		1						
(b) (i)	<table><tr><td>after adding acid</td><td>26.30</td></tr><tr><td>before adding acid</td><td>1.75</td></tr><tr><td>volume added</td><td>24.55</td></tr></table> M1 26.30 M2 1.75 M3 24.55	after adding acid	26.30	before adding acid	1.75	volume added	24.55	If readings are correct but in wrong order, award 1 mark for M1 and M2 M3 ECF from (M1 - M2) Penalise an answer not to 2dp once only	3
after adding acid	26.30								
before adding acid	1.75								
volume added	24.55								
(ii)	Any two from M1 to obtain (two) concordant results/titres M2 to identify anomalous results/first result may be anomalous M3 to obtain a mean (of concordant results/titres)	Or equivalent e.g. titres within up to 0.2 cm³ ALLOW to obtain (very) similar results OWTTE ALLOW reference to overshooting the end point on the first titration ALLOW average	2						

Total 6 marks

Question number	Answer	Notes	Marks
4 (a) (i)	C_3H_8O	Penalise incorrect use of lowercase letters and subscripts ALLOW symbols in any order e.g. C_3OH_8 REJECT C_3H_7OH	1
(ii)	M1 (correct as X) only contains single bonds M2 (not correct as X) contains oxygen/O M3 hydrocarbon contains hydrogen/H and carbon/C only	ALLOW no double bonds REJECT if they say it is unsaturated IGNORE O_2 ALLOW (not correct as X) is an alcohol REJECT if they say it is a hydrocarbon REJECT hydrogen and carbon molecules	3
(b)	Any one from: • have a general formula • successive members differ by CH_2 • trends in physical properties • same functional group	IGNORE reference to chemical properties	1

Question number	Answer	Notes	Marks
4 (c) (i)	catalyst	ALLOW description of effect of catalyst e.g. to speed up the reaction/ to lower the activation energy IGNORE to dehydrate compound X	1
(ii)	insoluble (in water)	IGNORE reference to density ALLOW (only) slightly soluble in water	1
(iii)	contains air	ALLOW contains oxygen IGNORE contains other gases	1
(iv)	$\begin{array}{c} \text{H} & & \text{H} \\ & & \\ \text{C} = & \text{C} - & \text{C} - \text{H} \\ & & \\ \text{H} & \text{H} & \text{H} \end{array}$ M1	IGNORE bond angles ALLOW H-O-H	2
(v)	M2 H ₂ O propene		1

Total 11 marks

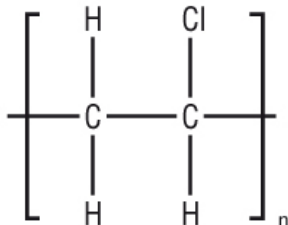
Question number	Answer	Notes	Marks						
5 (a)	<table><tr><td>Fe</td><td>iron</td></tr><tr><td>Cr</td><td>chromium</td></tr><tr><td>O</td><td>oxygen</td></tr></table>	Fe	iron	Cr	chromium	O	oxygen	All three names must be correct	1
Fe	iron								
Cr	chromium								
O	oxygen								
(b) (i)	$\text{Cr}_2\text{O}_3 + 2\text{Al} \rightarrow \text{Al}_2\text{O}_3 + 2\text{Cr}$ M1 Al_2O_3 M2 completely correct equation		2						
(ii)	M1 aluminium more reactive (than chromium) M2 as aluminium displaces chromium	ORA ALLOW aluminium higher in the reactivity series (than chromium) ALLOW aluminium removes the oxygen from chromium	2						
(iii)	Any two from M1 (redox means both) reduction and oxidation taking place (at same time) M2 chromium reduced and aluminium oxidised M3 chromium (reduced as) loses oxygen and aluminium (oxidised as) gains oxygen	ACCEPT chromium oxide ACCEPT chromium oxide ACCEPT chromium ions gain electrons (so reduced) and aluminium loses electrons (so oxidised) ALLOW answer in terms of oxidation number changes	2						

Question number	Answer	Notes	Marks
5 (c) (i)	M1 $\text{mol (Cr)} = 0.13/52 = 0.0025$ M2 $\text{vol H}_2 = \text{answer to M1} \times 24 =$ M3 $0.06(0) \text{ dm}^3$	accept 60 cm^3 If incorrect moles of H_2 e.g. 0.0025 has been multiplied or divided by 2, ECF to M3, i.e. 0.03 or 0.12 would score 2 marks	3
(ii)	Any two from M1 (sample of) chromium impure M2 gas leakage (from apparatus)/some gas escaped (before the bung was inserted) M3 not enough/less than 0.13g chromium used/did not use excess acid M4 reaction not allowed to go to completion/the volume was measured before reaction ended M5 some gas dissolved in the acid/solution		2

Total 12 marks

Question number	Answer	Notes	Marks
6 (a) (i)	$(24 + 16) = 40$		1
(ii)	Li^+ and F^-	both needed	1
(b)	Any four from M1 strong (electrostatic) forces/attraction M2 between oppositely charged ions M3 a large amount of energy needed to overcome attraction / break down lattice/break bonds M4 (MgO higher melting point as) greater charge on Mg^{2+} (than Li^+) / greater charge on O^{2-} (than F^-) M5 EITHER so stronger attraction/forces/bonds (in MgO) OR more (thermal/heat) energy required to overcome attraction / break down lattice/break the bonds (in MgO)/ORA	ACCEPT strong (ionic) bonds Must be a comparison to gain M5 MAX 2 if any reference to intermolecular forces/covalent bonding/electron sharing/molecules/metallic bonding	4
(c)	M1 (when) solid ions in fixed positions/don't move/only vibrate M2 (when) molten or in solution ions can move/mobile	IGNORE electrons ALLOW atoms REJECT electrons/atoms MAX 1 if mention of sharing of electrons/covalent bonding	2

Total 8 marks

Question number	Answer	Notes	Marks
7 (a) (i)	M1 oxidation is loss of electrons M2 chloride ions/ Cl^- lose electrons (so oxidised)	REJECT chlorine/ Cl/Cl_2 loses electrons ALLOW chloride loses electrons	2
(ii)	$2\text{H}_2\text{O} + 2\text{e}^- \rightarrow 2\text{OH}^- + \text{H}_2$ M1 all correct species M2 correctly balancing M2 dep on M1	ACCEPT multiples	2
(b)	$\text{Cl}_2 + 2\text{NaOH} \rightarrow \text{NaCl} + \text{NaOCl} + \text{H}_2\text{O}$	ACCEPT multiples	1
(c) (i)	M1 monomers join together / double bonds broken (in monomers) M2 to form a long chain (molecule) / large molecule	ALLOW link/add in place of join	2
(ii)	 M1 correct repeat structure M2 brackets and continuation bonds and n	No M1 if more than 1 repeat unit shown ACCEPT n anywhere after the brackets but not before REJECT any structure with a double bond for both marks	2

Total 9 marks

Paper Total 60 marks

