

Please check the examination details below before entering your candidate information

Candidates surname					Other names				
Centre Number					Candidate Number				
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**Pearson Edexcel International GCSE (9–1)**

Time 1 hour 15 minutes

Paper reference **4CH1/2CR**

**Chemistry**

**Unit: 4CH1**

**PAPER: 2CR**

**You must have:**  
Calculator, ruler

Total Marks

### Instructions

- Use **black** ink or ball-point pen.
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided  
– *there may be more space than you need.*
- Show all the steps in any calculations and state the units.

### Information

- The total mark for this paper is 70.
- The marks for **each** question are shown in brackets  
– *use this as a guide as to how much time to spend on each question.*

### Advice

- Read each question carefully before you start to answer it.
- Write your answers neatly and in good English.
- Try to answer every question.
- Check your answers if you have time at the end.

Turn over ►

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**Answer ALL questions.**

**Some questions must be answered with a cross ☒. If you change your mind about an answer, put a line through the box ☒ and then mark your new answer with a cross ☒.**

- 1 (a) Two substances are needed to cause iron to rust.

Name these two substances.

(2)

1 .....

2 .....

- (b) The box gives the names of some substances.

calcium	copper	gold
iodine	methane	zinc

Use words from the box to answer these questions.

- (i) Give the name of a non-metallic element.

(1)

.....

- (ii) Give the name of a compound.

(1)

.....

- (iii) Give the name of the metal that is lowest in the reactivity series.

(1)

.....

**(Total for Question 1 = 5 marks)**

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P 7 0 7 0 4 A 0 3 2 0

2 Crude oil is a mixture of hydrocarbons.

(a) This passage is about the industrial separation of crude oil.

Complete the passage by adding the missing words.

(3)

Crude oil is ..... to form vapour.

The vapour is passed through a ..... column.

The refinery gases are collected at the top of the column because they have low

.....

(b) Bitumen is collected at the bottom of the column.

Give one use of bitumen.

(1)

(c) One of the hydrocarbons in crude oil is an alkane with this structural formula.



(i) Give the name of this alkane.

(1)

(ii) Calculate the relative molecular mass ( $M_r$ ) of this alkane.

(1)

$M_r =$  .....

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(d) Catalytic cracking is used to convert long-chain alkanes into shorter-chain alkanes.

Give the name of the catalyst and the temperature used in catalytic cracking.

(2)

catalyst .....

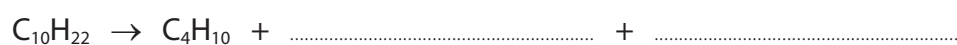
temperature .....

(e) Catalytic cracking also produces alkenes.

Decane ( $C_{10}H_{22}$ ) can undergo cracking to give  $C_4H_{10}$  and two different alkenes.

Complete the equation for this cracking process.

(2)



**(Total for Question 2 = 10 marks)**

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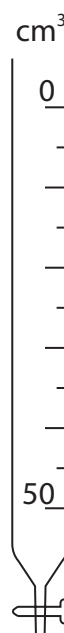
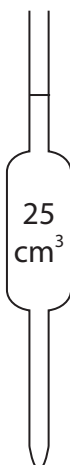
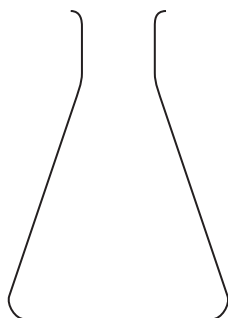
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3 A student does a titration to find the concentration of a solution of dilute sulfuric acid.

The student uses these solutions and this apparatus.

- dilute sulfuric acid
- potassium hydroxide solution of concentration  $0.240 \text{ mol/dm}^3$
- methyl orange indicator



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- (a) The student wants to find the volume of sulfuric acid needed to neutralise  $25.0\text{ cm}^3$  of the potassium hydroxide solution.

Describe how the student should do this titration.

Assume that all pieces of apparatus are clean and dry.

(6)

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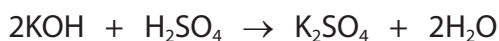
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Area for writing the answer, consisting of multiple horizontal dotted lines.



- (b) The student needs  $15.00 \text{ cm}^3$  of sulfuric acid to neutralise  $25.0 \text{ cm}^3$  of the potassium hydroxide solution.

This is the equation for the reaction.



- (i) Calculate the amount, in moles, of KOH in  $25.0 \text{ cm}^3$  of potassium hydroxide solution of concentration  $0.240 \text{ mol/dm}^3$ .

(2)

amount of KOH = ..... mol

- (ii) Calculate the amount, in moles, of  $\text{H}_2\text{SO}_4$  in  $15.00 \text{ cm}^3$  of the sulfuric acid.

(1)

amount of  $\text{H}_2\text{SO}_4$  = ..... mol

- (iii) Calculate the concentration, in  $\text{mol/dm}^3$ , of the sulfuric acid.

(2)

concentration of sulfuric acid = .....  $\text{mol/dm}^3$

**(Total for Question 3 = 11 marks)**



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4 This question is about alcohols, carboxylic acids and their reactions.

(a) The boxes give some information about a carboxylic acid.

Complete the boxes by giving the missing information.

(3)

structural formula	CH <sub>3</sub> COOH
name	
	CH <sub>2</sub> O
displayed formula	

(b) Ethanol can be oxidised to produce a carboxylic acid.

(i) Give the names of the two reagents used in this oxidation reaction.

(2)

1 .....

2 .....

(ii) Which of these colour changes occurs during the reaction?

(1)

- A** green to orange
- B** orange to green
- C** red to yellow
- D** yellow to red

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(c) Alcohols and carboxylic acids can be heated together to form esters.

- (i) State why it is better to heat the mixture using a water bath rather than directly with a Bunsen burner flame.

(1)

- (ii) An ester has the structural formula  $\text{CH}_3\text{CH}_2\text{COOCH}_3$

Which of these is the name of this ester?

(1)

- A ethyl methanoate
- B methyl ethanoate
- C methyl propanoate
- D propyl methanoate

**(Total for Question 4 = 8 marks)**

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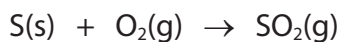
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5 This question is about three stages in the manufacture of sulfuric acid.

(a) In stage 1, sulfur is burned in oxygen to form sulfur dioxide gas.



(i) State one environmental problem caused by the release of sulfur dioxide into the atmosphere.

(1)

(ii) A mass of 6.4 tonnes of sulfur is burned to produce sulfur dioxide gas.

Calculate the maximum volume, in  $\text{dm}^3$ , of sulfur dioxide gas that can be produced at rtp.

[molar volume of sulfur dioxide gas at rtp =  $24 \text{ dm}^3$ ]

[1 tonne =  $10^6 \text{ g}$ ]

Give your answer in standard form.

(3)

maximum volume = .....  $\text{dm}^3$

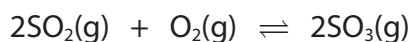
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(b) In stage 2, sulfur dioxide is reacted with oxygen to form sulfur trioxide gas.



The yield of sulfur trioxide is approximately 98%.

(i) A catalyst is used in this reaction.

Explain how a catalyst increases the rate of a reaction.

(2)

(ii) The temperature is kept constant.

Give a reason why increasing the pressure would increase the yield of sulfur trioxide.

(1)

(iii) Suggest why it is not necessary to increase the pressure in stage 2.

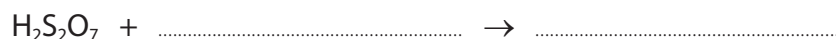
(1)

(c) In stage 3, the sulfur trioxide is reacted with concentrated sulfuric acid to form a liquid called oleum,  $\text{H}_2\text{S}_2\text{O}_7$

The oleum is then added to water to form concentrated sulfuric acid.

Complete the chemical equations for these two reactions.

(2)



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(d) Sulfuric acid reacts with ammonia to form ammonium sulfate,  $(\text{NH}_4)_2\text{SO}_4$

Calculate the percentage by mass of nitrogen in ammonium sulfate.

$[M_r \text{ of } (\text{NH}_4)_2\text{SO}_4 = 132]$

(2)

percentage = ..... %

**(Total for Question 5 = 12 marks)**

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6 A teacher prepares the insoluble salt lead(II) bromide ( $\text{PbBr}_2$ ) by mixing solutions of lead(II) nitrate and sodium bromide.

(a) Describe what the teacher should do next to obtain a pure, dry sample of lead(II) bromide.

(3)

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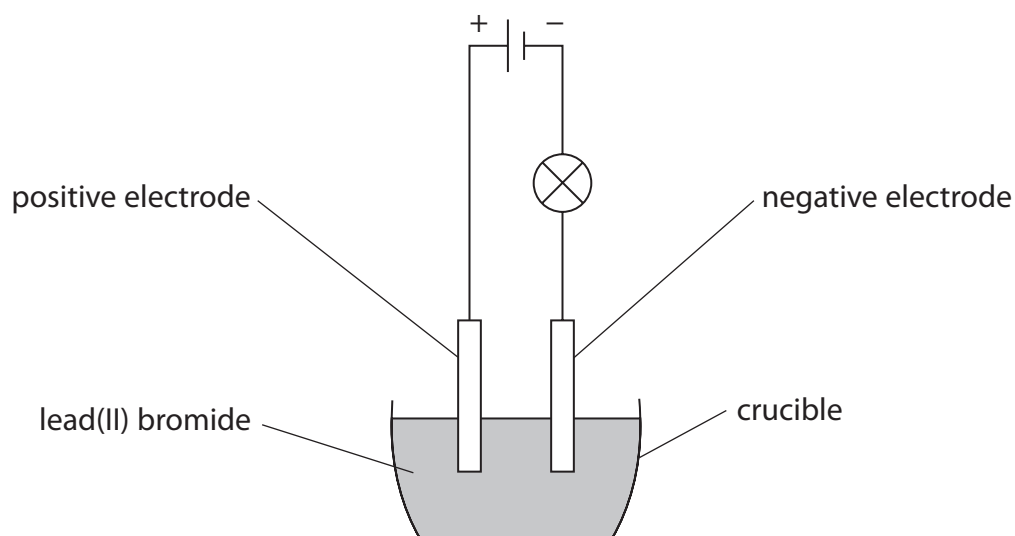
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- (b) The teacher then sets up a circuit in a fume cupboard using the pure, dry sample of lead(II) bromide.



Explain why the lamp does not light when the lead(II) bromide is solid.

(2)

.....

.....

.....

.....

- (c) The teacher heats the lead(II) bromide.

When the lead(II) bromide is molten, the lamp lights and bromine forms at the positive electrode.

- (i) State what observation would be made at the positive electrode.

(1)

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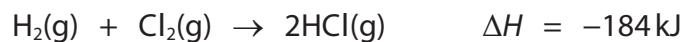
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7 The reaction between hydrogen and chlorine is exothermic.

This is the equation for the reaction.



(a) State the meaning of the term **exothermic**.

(1)

(b) The table gives the bond energies for the H—H and H—Cl bonds.

<b>Bond</b>	H—H	H—Cl
<b>Bond energy in kJ/mol</b>	436	431

Use the equation and information from the table to calculate the bond energy of the Cl—Cl bond.

(4)

bond energy = ..... kJ/mol

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(c) Explain why this reaction is exothermic.

Refer to bond-breaking and bond-making in your answer.

(3)

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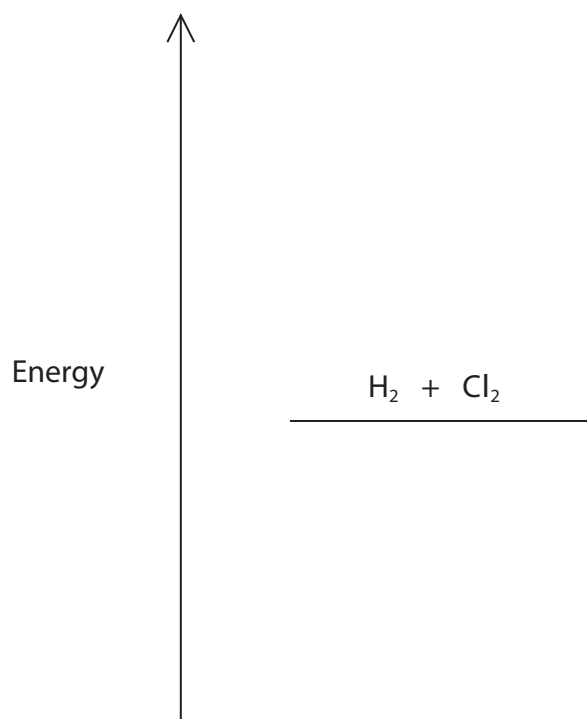
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(d) Complete the reaction profile diagram to show the position of the products, the enthalpy change ( $\Delta H$ ) and the activation energy ( $E_a$ ) for the reaction.

(4)



(Total for Question 7 = 12 marks)

TOTAL FOR PAPER = 70 MARKS



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