

Please check the examination details below before entering your candidate information

| | | | | | | | | | |
|----------------------|----------------------|----------------------|----------------------|----------------------|----------------------|----------------------|----------------------|----------------------|----------------------|
| Candidates surname | | | | | Other names | | | | |
| Centre Number | | | | | Candidate Number | | | | |
| <input type="text"/> | <input type="text"/> | <input type="text"/> | <input type="text"/> | <input type="text"/> | <input type="text"/> | <input type="text"/> | <input type="text"/> | <input type="text"/> | <input type="text"/> |

Pearson Edexcel International GCSE (9–1)

Monday 22 May 2023

Morning (Time: 2 hours)

| | |
|-----------------|------------------------|
| Paper reference | 4CH1/1C 4SD0/1C |
|-----------------|------------------------|

Chemistry

UNIT: 4CH1

Science (Double Award) 4SD0

PAPER: 1C

You must have:
Calculator, ruler

Total Marks

Instructions

- Use **black** ink or ball-point pen.
- If pencil is used for diagrams/sketches/graphs it must be dark (HB or B).
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided
– *there may be more space than you need.*
- Show all the steps in any calculations and state the units.

Information

- The total mark for this paper is 110.
- The marks for **each** question are shown in brackets
– *use this as a guide as to how much time to spend on each question.*

Advice

- Read each question carefully before you start to answer it.
- Write your answers neatly and in good English.
- Try to answer every question.
- Check your answers if you have time at the end.

Turn over ►

P71950A

©2023 Pearson Education Ltd.

N:1/1/1/1/



Pearson

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

BLANK PAGE



P 7 1 9 5 0 A 0 3 3 2

Answer ALL questions.

Some questions must be answered with a cross ☒. If you change your mind about an answer, put a line through the box ☒ and then mark your new answer with a cross ☒.

- 1 The box gives some methods that are involved in the separation of mixtures.

| | | | |
|----------------|-------------------------|---------------------|-------------|
| chromatography | crystallisation | dissolving | evaporating |
| filtering | fractional distillation | simple distillation | |

- (a) Use words from the box to identify the method involved in each of these separations.

- (i) Give the best method for obtaining gasoline from crude oil. (1)

- (ii) Give the best method for separating the dyes in black ink. (1)

- (iii) Give the best method for obtaining pure water from seawater. (1)

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA



- (b) A sample of solid hydrated copper(II) sulfate can be obtained from a mixture of copper(II) oxide and copper(II) sulfate.

Complete the passage by using words from the box.

(4)

The mixture of copper(II) oxide and copper(II) sulfate can be separated by first

..... the copper(II) sulfate in distilled water.

The copper(II) oxide is then removed by

.....

Some of the water from the copper(II) sulfate solution is then removed by

.....

A pure sample of hydrated copper(II) sulfate is then obtained by

.....

(Total for Question 1 = 7 marks)

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA



2 This question is about the reactions of iron.

(a) Iron rusts when exposed to water and oxygen.

(i) Give the chemical name of the compound that forms when iron rusts. (1)

(ii) What type of reaction occurs when iron rusts? (1)

- A combustion
- B decomposition
- C neutralisation
- D oxidation

(iii) Galvanising is a method used to prevent iron from rusting.

Give the name of the metal used to galvanise iron. (1)

(b) When iron reacts with dilute sulfuric acid, the products are iron(II) sulfate and hydrogen.

(i) Give a chemical equation for the reaction between iron and sulfuric acid. (1)

(ii) Give a test for hydrogen. (1)

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA



DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

(c) An excess of iron is added to copper(II) sulfate solution.

(i) Name the type of reaction that occurs.

(1)

(ii) State the appearance of the solid that forms in the reaction.

(1)

(d) Give the reason why no reaction occurs when iron is added to magnesium sulfate solution.

(1)

(Total for Question 2 = 8 marks)



3 The table gives some information about three substances, X, Y and Z.

| Substance | Melting point | Conducts electricity when solid | Conducts electricity when molten | Type of bonding | Type of structure |
|-----------|---------------|---------------------------------|----------------------------------|-----------------|-------------------|
| X | low | no | no | covalent | simple molecular |
| Y | high | no | no | | |
| Z | high | no | yes | | |

(a) Complete the table by giving the missing information.

(4)

(b) Explain why substance X has a low melting point.

(2)

.....

.....

.....

.....

.....

.....

.....

(Total for Question 3 = 6 marks)



4 This question is about unsaturated hydrocarbons.

(a) Ethene (C_2H_4) is a member of the homologous series of alkenes.

(i) Give two characteristics of a homologous series.

(2)

1

.....

2

.....

(ii) Draw a dot-and-cross diagram to show the bonding in a molecule of ethene.

Show outer electrons only.

(2)

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA



DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

(b) Propene (C_3H_6) is another member of the homologous series of alkenes.

- (i) State why the empirical formula of all alkenes in this homologous series is CH_2 (1)
-
-

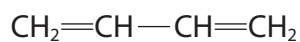
- (ii) Propene can be polymerised to form poly(propene).

Draw the displayed formula of propene and the repeat unit of poly(propene). (2)

| propene | repeat unit of poly(propene) |
|---------|------------------------------|
| | |



(c) This is the structural formula of another hydrocarbon compound.



(i) Give the molecular formula and the empirical formula of this compound.

(2)

molecular formula

empirical formula

(ii) Explain why this compound is an unsaturated hydrocarbon.

(3)

(iii) Describe a test to show that this hydrocarbon is unsaturated.

(2)

(Total for Question 4 = 14 marks)

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA



P 7 1 9 5 0 A 0 1 1 3 2

5 This question is about lithium and some of its compounds.

(a) A small piece of lithium is added to a trough containing water.

The lithium floats on the surface of the water and a vigorous reaction occurs.

(i) Give two other observations when lithium reacts with water.

(2)

1

.....

2

.....

(ii) A few drops of methyl orange are added to the solution in the trough.

Explain the final colour of the solution.

(2)

.....

.....

.....

.....

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA



(b) An unlabelled bottle containing a white powder is found in a laboratory.

Describe tests to show that the white powder in the bottle is lithium carbonate.

(5)

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

(c) Lithium carbonate has ionic bonding.

State what is meant by the term **ionic bonding**.

(2)

DO NOT WRITE IN THIS AREA

(Total for Question 5 = 11 marks)



P 7 1 9 5 0 A 0 1 3 3 2

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

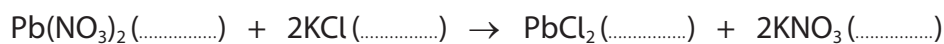
DO NOT WRITE IN THIS AREA

BLANK PAGE



6 When solutions of lead(II) nitrate and potassium chloride are mixed, a precipitate of lead(II) chloride forms.

(a) (i) Complete the equation for the reaction by adding the state symbols. (1)



(ii) Give the formula of each ion in lead(II) nitrate. (1)

lead ion

nitrate ion

(iii) Calculate the relative formula mass (M_r) of lead(II) nitrate, $\text{Pb}(\text{NO}_3)_2$ (2)

$M_r = \text{.....}$

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA



P 7 1 9 5 0 A 0 1 5 3 2

- (b) A student investigates the height of the precipitate formed when lead(II) nitrate solution is added to potassium chloride solution.

This is the student's method.

Step 1 pour 15.0 cm^3 of potassium chloride solution into a boiling tube

Step 2 add 2.0 cm^3 of lead(II) nitrate solution and allow the precipitate to settle

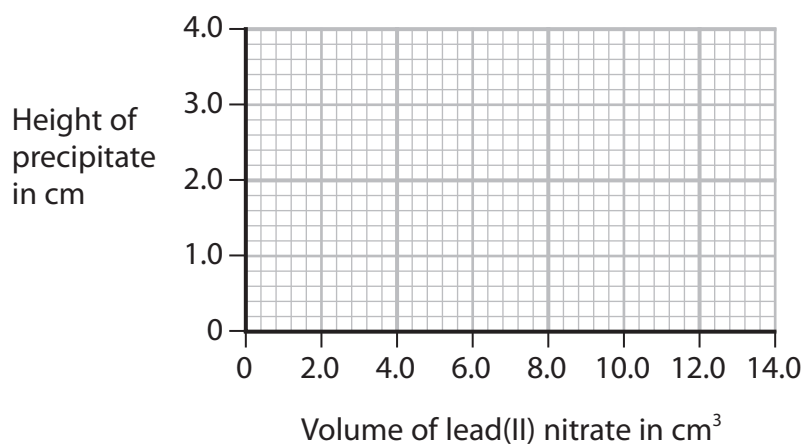
Step 3 measure the height of the precipitate

Repeat steps 2 and 3 until a total of 14.0 cm^3 of lead(II) nitrate solution has been added.

The table shows the student's results.

| | | | | | | | |
|---|-----|-----|-----|-----|------|------|------|
| Volume of lead(II) nitrate in cm^3 | 2.0 | 4.0 | 6.0 | 8.0 | 10.0 | 12.0 | 14.0 |
| Height of precipitate in cm | 0.8 | 1.6 | 2.9 | 3.2 | 3.6 | 3.6 | 3.6 |

- (i) Plot the results on the grid. (1)
- (ii) Draw a circle around the anomalous result. (1)
- (iii) Draw a line of best fit through the first four points and another line of best fit through the last three points. Make sure that the lines cross. (2)



(iv) Give two possible mistakes the student could have made to cause the anomalous result.

(2)

1

.....

2

.....

(v) State why the first line of best fit should pass through the origin.

(1)

.....

.....

(vi) Use your graph to determine the volume of lead(II) nitrate solution needed to react completely with 15.0 cm^3 of potassium chloride solution.

(1)

volume = cm^3

(Total for Question 6 = 12 marks)

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA



P 7 1 9 5 0 A 0 1 7 3 2

7 This question is about the three halogens, bromine, chlorine and iodine.

(a) Give the number of protons and the number of neutrons in an atom of iodine-127

(2)

number of protons

number of neutrons

(b) A sample of bromine contains two isotopes.

- Br-79 with relative abundance 52.8%
- Br-81 with relative abundance 47.2%

Calculate the relative atomic mass (A_r) of this sample of bromine.

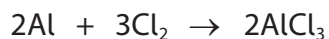
Give your answer to three significant figures.

(3)

$A_r = \dots\dots\dots$

(c) Aluminium reacts with chlorine to form aluminium chloride.

This is the equation for the reaction.



Calculate the minimum mass of chlorine needed to form 26.7 g of aluminium chloride.

[for Cl_2 , $M_r = 71$ for AlCl_3 , $M_r = 133.5$]

(3)

minimum mass of chlorine = g

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA



DO NOT WRITE IN THIS AREA

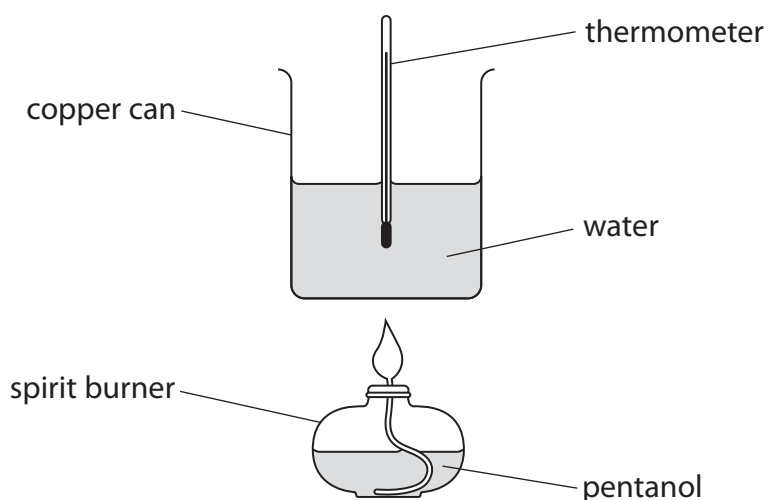
DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

BLANK PAGE



- 8 A student uses this apparatus to find the molar enthalpy change (ΔH) of combustion for the liquid fuel, pentanol.



This is the student's method.

- find the initial mass of the spirit burner and pentanol
- add 100 cm^3 of water to the copper can
- record the initial temperature of the water
- light the wick of the spirit burner to heat the water
- stir the water until the temperature rises by 35.0°C
- extinguish the flame and immediately find the final mass of the spirit burner and pentanol

- (a) (i) State why the student stirs the water.

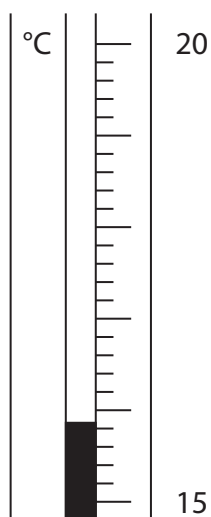
(1)

- (ii) Suggest why it is important that the student immediately finds the final mass of the spirit burner and pentanol.

(1)



(b) The diagram shows the initial temperature of the water.



Complete the table to show the temperature readings.

Give both values to the nearest 0.1 °C.

(2)

| | |
|------------------------------------|------|
| Initial temperature of water in °C | |
| Final temperature of water in °C | |
| Temperature change in °C | 35.0 |

(c) (i) Show by calculation that the heat energy (Q) supplied by the pentanol is approximately 15 000 J.

[for water, $c = 4.2 \text{ J/g/}^\circ\text{C}$]

[for 1.0 cm^3 of water, mass = 1.0 g]

(2)



(ii) The table gives the initial and final mass readings.

| | |
|---|-------|
| Initial mass of spirit burner and pentanol in g | 90.11 |
| Final mass of spirit burner and pentanol in g | 89.75 |

Use your answer to part (c)(i) and the information in the table to calculate the molar enthalpy change (ΔH) of combustion, in kJ/mol, for pentanol.

[for pentanol, $M_r = 88$]

Include a sign in your answer.

(5)

ΔH for pentanol = kJ/mol

(d) The formula of pentanol is $C_5H_{11}OH$

Write a chemical equation for the complete combustion of pentanol.

(2)

(Total for Question 8 = 13 marks)



9 This question is about different hydrated forms of sodium sulfate.

(a) A compound has the formula $\text{Na}_2\text{SO}_4 \cdot 7\text{H}_2\text{O}$

(i) How many different elements are there in the formula $\text{Na}_2\text{SO}_4 \cdot 7\text{H}_2\text{O}$? (1)

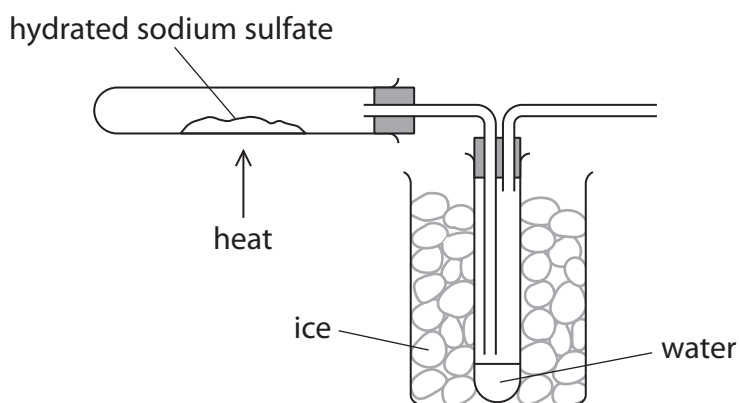
- A 3
- B 4
- C 5
- D 10

(ii) What is the total number of atoms in the formula $\text{Na}_2\text{SO}_4 \cdot 7\text{H}_2\text{O}$? (1)

- A 10
- B 22
- C 27
- D 28

(b) Another hydrated form of sodium sulfate has the formula $\text{Na}_2\text{SO}_4 \cdot x\text{H}_2\text{O}$

A student uses this apparatus to find the value of x .



DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA



This is the student's method.

- find the mass of an empty tube
 - add solid hydrated sodium sulfate to the tube
 - find the mass of the tube and hydrated sodium sulfate
 - heat the tube for several minutes
 - allow the tube to cool and find the mass of the tube and contents
- (i) Describe what the student should do next to make sure that all the water is removed from the hydrated sodium sulfate.

(2)

- (ii) Explain the role of the ice in the beaker.

(2)

- (iii) Describe how the student could prove that the liquid collected is pure water.

(2)

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA



P 7 1 9 5 0 A 0 2 5 3 2

(c) The table gives the student's results.

| | |
|--|-------|
| Mass of empty tube in g | 15.83 |
| Mass of tube and $\text{Na}_2\text{SO}_4 \cdot x\text{H}_2\text{O}$ in g | 23.88 |
| Mass of tube and Na_2SO_4 in g | 19.38 |

Use the student's results to calculate the value of x .

[for Na_2SO_4 , $M_r = 142$ for H_2O , $M_r = 18$]

(5)

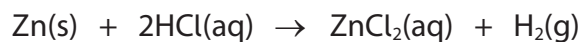
$x = \dots\dots\dots$

(Total for Question 9 = 13 marks)



10 Zinc reacts with dilute hydrochloric acid to form zinc chloride and hydrogen gas.

This is the equation for the reaction.



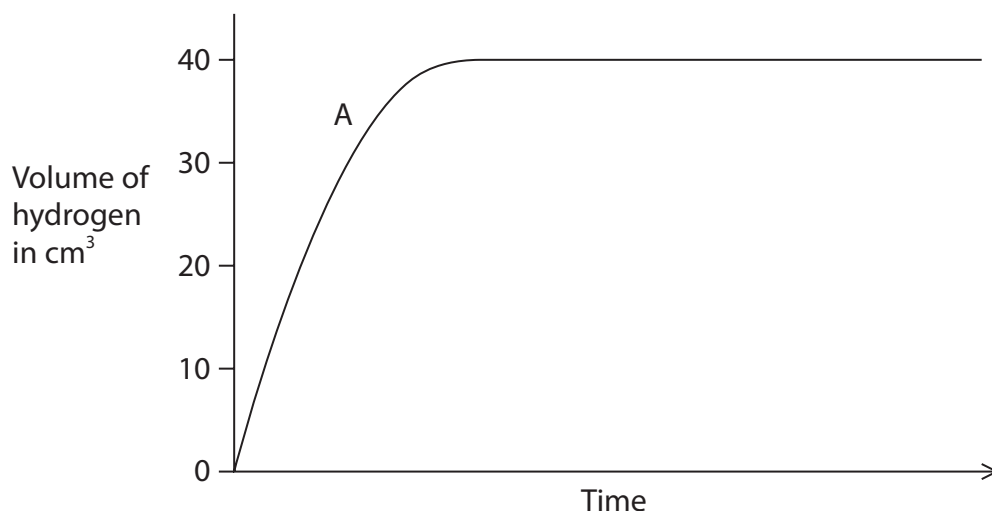
(a) In an experiment, 20 cm³ of hydrochloric acid containing 0.0036 mol are reacted with 1.3 g of zinc granules at a temperature of 30 °C.

(i) Show by calculation that the zinc is in excess.

(2)

(ii) The volume of hydrogen collected is measured at regular time intervals.

Curve A shows the results of this experiment.



The experiment is repeated using 1.3 g of zinc powder instead of zinc granules.

All other conditions are kept the same.

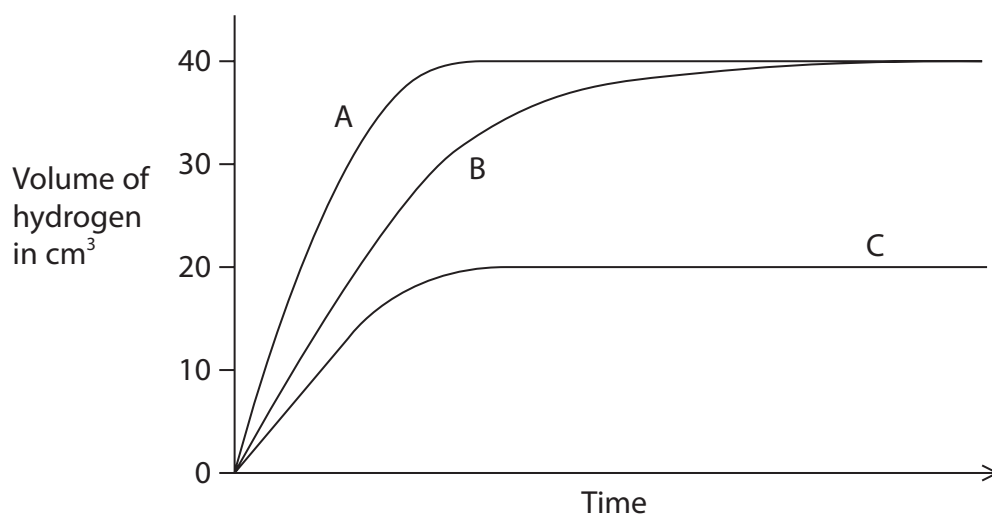
On the grid, draw the curve you would expect to obtain.

(2)



- (b) In the original experiment, 20 cm^3 of hydrochloric acid containing 0.0036 mol were reacted with 1.3 g of zinc granules at a temperature of 30°C and curve A was obtained.

The student does two more experiments and obtains curves B and C.



- (i) In one of these experiments the student repeats the original method but at a temperature of 20°C .

Explain in terms of particle collision theory why the curve obtained could be curve B.

(4)

.....

.....

.....

.....

.....

.....

.....

.....

.....

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA



- (ii) In the other experiment the student repeats the original method but uses 20 cm^3 of hydrochloric acid containing 0.0018 mol .

Explain why curve C shows the results the student obtained.

(2)

- (c) Catalysts can be used to speed up reactions.

Describe how a catalyst works.

(2)

(Total for Question 10 = 12 marks)

TOTAL FOR PAPER = 110 MARKS



DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

BLANK PAGE



DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

BLANK PAGE



P 7 1 9 5 0 A 0 3 1 3 2

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

BLANK PAGE

