

Mark Scheme (Results)

Summer 2012

International GCSE

Chemistry (4CH0) Paper 1C

Science Double Award (4SC0) Paper 1C

Edexcel Level 1/Level 2 Certificate

Chemistry (KCH0) Paper 1C

Science (Double Award) (KSC0) Paper 1C

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Publications Code UG031876

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INTERNATIONAL GCSE CHEMISTRY PAPER 1C – SUMMER 2012

Question number	Expected Answer			Accept	Reject	Marks
1 (a)				Ignore extra apparatus on top of tripod eg beaker Minimum is 2 vertical or diagonal lines with something on top		6
		measuring cylinder	B			
	 Or 2D diagram		C Ignore D			
		(top pan) balance /scale(s) / weighing machine	A or E			
1 mark for each correct answer						

Question number	Expected Answer	Accept	Reject	Marks
1 (b)	<p>M1 wear (safety) glasses / spectacles / goggles / eye protection</p> <p>M2 salt /solution / water may spit out (when evaporating the salty water) / may get in your eye IGNORE references to hazards eg toxic / irritant</p> <p>OR</p> <p>M1 use (beaker) tongs / hot hand / (rigger/oven) glove(s) (to remove / lift the basin)</p> <p>M2 basin will / may be hot</p> <p>OR</p> <p>M1 tie hair back / tuck in tie</p> <p>M2 might catch fire (in Bunsen burner)</p> <p>the reason must mention the precaution</p> <p>IGNORE reference to wearing lab. coats / protective clothing</p>	<p>It</p> <p>leave basin (to cool) <u>before removing</u></p> <p>to avoid burning hand</p>	<p>crucible tongs / plastic gloves</p>	<p>1</p> <p>1</p>
(c)	(2.9 x 2) = 5.8 (g)			1

Question number	Expected Answer	Accept	Reject	Marks
2 (a)	(i) M1 calcium M2 magnesium	Ca Mg	any other answers	1 1
	(ii) iron / zinc	Fe / Zn	any other answers	1
	(iii) calcium magnesium zinc iron copper M1 for calcium as most reactive M2 for copper as least reactive M3 for remainder in correct order	Ca Mg Zn Fe Cu		3
(b)	(i) hydrogen / H ₂		H	1
	(ii) <u>all</u> the (sulfuric) acid has reacted / <u>all</u> hydrogen (ions) have been replaced (by magnesium (ions)) OR acid has been used up/been neutralised / acid has run out IGNORE the acid saturated / excess magnesium has been added	sulphuric for sulfuric hydrogen ions / H ⁺ for acid	all the magnesium / reactants used up	1
	(iii) magnesium sulfate (solution) IGNORE incorrect formula	sulphate for sulfate MgSO ₄		1
	(iv) filtration / filter (it / magnesium / solution) / decantation / decant (off the water / solution) IGNORE references to distillation / centrifuging / washing / evaporation <u>after</u> filtration	description of filtration	sieve crystallisation	1

2	(c)	(i)	exothermic			1
		(ii)	magnesium oxide IGNORE incorrect formula	MgO		1

Question number	Expected Answer	Accept	Reject	Marks
3 (a)	<p>M1 precipitate of barium sulfate</p> <p>M2 no precipitate</p> <p>M3 precipitate of calcium sulfate</p> <p>IGNORE colours</p> <p>penalise incorrect extra observations (e.g. effervescence) ONCE only</p> <p>For M1 and M3 only: if only precipitate appears twice (with no names), penalise <u>missing</u> names once only</p> <p>if only names correct (with no precipitates), penalise omission of precipitate once only</p>	<p>sulphate for sulfate insoluble barium sulphate / BaSO₄</p> <p>no (visible) change solution (formed)</p> <p>sulphate for sulfate insoluble calcium sulfate / CaSO₄</p>	<p>incorrect name of ppt.</p> <p>incorrect name of ppt.</p>	<p>1</p> <p>1</p> <p>1</p>

Question number	Expected Answer	Accept	Reject	Marks
3 (b)	aq aq s aq			1
(c) (i)	obtain the lead(II) bromide/the residue/the solid OR remove the liquid/solution/potassium nitrate/water	separate the solid and liquid		1
(ii)	to wash away/remove the (remaining) potassium nitrate / lead(II) nitrate / potassium bromide / solution IGNORE clean	wash away / remove (remaining soluble) impurities to make it pure	make the mixture pure	1
(iii)	distilled water is pure / does not contain (dissolved) impurities / ions / substances / compounds / other chemicals (that would contaminate the lead(II) bromide) / residue / solid) IGNORE elements IGNORE references to distilled water being cleaner (ORA)	reverse argument for tap water	any suggestion that the water / impurities react	1
(iv)	to evaporate the water / to dry (the solid/crystals) / increase rate of evaporation (of water) IGNORE liquid	to avoid decomposition (if heated strongly)	to evaporate the potassium nitrate / solution any reference to crystallisation	1

Question number	Expected Answer	Accept	Reject	Marks
4 (a)	(increasing) atomic number(s) IGNORE references to electrons / electronic configurations	proton number / number of protons	mass number / RAM	1
(b) (i)	sodium / potassium	Na / K		1
(b) (ii)	fluorine / chlorine / bromine	F / Cl / Br / F ₂ / Cl ₂ / Br ₂	fluoride / chloride / bromide	1
(c) (i)	sodium OR potassium <u>AND</u> fluorine OR chlorine OR bromine OR hydrogen Answers can be in either order IGNORE incorrect symbols/formulae if names are correct	Na / K F / Cl / Br / H / F ₂ / Cl ₂ / Br ₂ / H ₂	fluoride / chloride / bromide / hydride	1
(c) (ii)	Marks do not have to be CQ on (c)(i), and all marks can be scored here for correct diagrams of the ions in a hydrogen halide M1 Na or K with 8 electrons M2 F, Cl or Br with 7 electrons IGNORE diagrams showing initial electron configurations M3 (1)+ <u>AND</u> (1)- charges correct IGNORE inner shells even if incorrect	0 electrons H with 2 electrons	Incorrect electron transfer for M1 and M2	1 1 1

Allow any combination of dots and crosses

If shown covalently bonded, then max. 1 for correct charges if given

If the position of 2 electrons shown between the two species makes it hard to be sure that the bonding is definitely ionic (and not covalent), do not award M1 or M2

Question number	Expected Answer	Accept	Reject	Marks
4 (d)	(fluorine reacts) vigorously / instantly / explosively / violently / <u>very</u> quickly / <u>very</u> rapidly	the quickest / more quickly <u>than chlorine</u>	fluorine reaction slower than chlorine reaction	1
	IGNORE references to electron transfer, even if incorrect (to form) iron(III) fluoride	ferric fluoride / FeF ₃		1
(e)	M1 colourless (IGNORE clear)	no colour	decolourised	1
	M2 orange / yellow / brown IGNORE qualifiers such as light / dark	any combination of colours on left	any other colour	1

Question number	Expected Answer	Accept	Reject	Marks
5 (a)	$2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ M1 all formulae correct (including catalyst if given) M2 correct balancing M2 DEP on M1 If catalyst included in equation, must be MnO_2 on both sides IGNORE MnO_2 above the arrow	$\text{H}_2\text{O}_2 \rightarrow \text{H}_2\text{O} + \frac{1}{2}\text{O}_2$ multiples		2
(b)	relights a glowing spill IGNORE reference to popping	splint for spill smouldering/embering for glowing		1
(c)	M1 (rate) increases M2 provides an alternative pathway / route / mechanism (for the reaction) OR hydrogen peroxide particles / molecules / reactant(s) adsorb (onto catalyst) M3 with a lower activation energy OR more particles / molecules have the (required) activation energy OR weakens the (covalent) bonds (in the hydrogen peroxide)	speeds up / goes faster / decreases time (for decomposition) lowers the activation energy by going a different way = M2 and M3 Absorb / sticks to / bonds to / provides a surface for particles / molecules / reactant(s) to react description of activation energy eg particles have enough energy to react	gives particles more kinetic energy for M2 and M3 atoms atoms	1 1 1

Question number	Expected Answer	Accept	Reject	Marks
5 (d) (i)	M1 curve starting at origin and below original curve			1
	M2 levelling off at 30 cm ³ (+/- 0.5) and anywhere between 30s and 120s			1
	(ii) M1 curve starting at origin and above original curve			1
	M2 levelling off at 60 cm ³ (+/- 0.5) and before 80s if curves incorrectly labelled then penalise each curve 1 mark, so max. 2 for the question			1

Question number	Expected Answer	Accept	Reject	Marks
6 (a)	<p>M1 both protons = 6</p> <p>M2 C-13 has 7 and C-14 has 8 (neutrons)</p>			1 1
(b)	<p>same electronic configuration(s) / structure(s)</p> <p>OR</p> <p>same <u>number</u> of electrons</p> <p>OR</p> <p>have <u>four/same number of</u> electrons in <u>outer / valence</u> shell</p> <p>IGNORE same number of electrons in inner shells IGNORE references to atomic number / same number of protons / different number of neutrons</p>	amount for number / six electrons	different number of protons	1
(c) (i)	<p>M1 the average / mean mass of an <u>atom</u> (of the element)</p> <p>M2 compared to / relative to (1/12th) the mass (of an atom) of carbon-12</p> <p>OR</p> <p>M1 mass of one mole of <u>atoms</u></p> <p>M2 compared to (mass of) 1/12th one mole / 1g of carbon-12</p>	<p>average/mean of: atomic masses / mass numbers / mass of isotopes</p> <p>on a scale where carbon-12 has a mass of 12 / compared with the mass of carbon-12 which is 12</p>	<p>mean mass of an element</p> <p>mass of one mole of the element</p>	1 1

Question number	Expected Answer	Accept	Reject	Marks
6 c (ii)	M1 $(12 \times 98.9) + (13 \times 1.1)$	(12 x 0.989) + (13 x 0.011) for first 2 marks 12.011 on its own for 2 marks 12.01 on its own for 3 marks		1
	M2 $\neq 100$			1
	M3 12.01			1
	IGNORE units			

Question number	Expected Answer	Accept	Reject	Marks
7 (a) (i)	M1 contains carbon and hydrogen (atoms / elements / particles)	C and H for carbon and hydrogen	ions / carbon molecules / hydrogen molecules / H ₂ / mixture of C and H	1
	M2 only	other equivalent words, eg solely / entirely / completely		1
	M2 DEP on M1, but allow M2 if molecules / ions / mixture used in M1			
(ii)	C ₁₀ H ₂₂ IGNORE structural formula	H ₂₂ C ₁₀	Reject superscripts / lower case c or h / full size numbers	1
(b) (i)	addition	additional		1
(ii)	M1 one of the bonds in the double bond breaks	double bond breaks / double bond becomes single bond		1
	M2 (many) <u>ethene(s)/molecules/monomers</u> join (together) OR (many) <u>ethene(s)/molecules/monomers</u> form a chain	changes (from unsaturated) to saturated		1

Question number	Expected Answer	Accept	Reject	Marks
7 (c)	Any 4 from: <ul style="list-style-type: none"> • produces smaller / shorter (chain) molecules • smaller / shorter (chain) molecules more useful (as fuels) / have greater demand • smaller / shorter (chain) molecules burn more cleanly / are used to make petrol/diesel/fuel for vehicles • crude oil richer in / has a surplus of long (chain) molecules • produces alkenes / any named alkene • alkenes used to make alcohol / polymers / plastics / chemical feedstock / any named addition polymer 	ORA low(er) demand products converted to high(er) demand products ORA		4

Question number	Expected Answer	Accept	Reject	Marks
8 (a) (i)	diffusion			1
	(ii) ammonia because it moves further (in the same time) / ammonia moved 60cm and hydrogen chloride moved 40cm OR ammonia because (white) ring right of centre / ring is further from ammonia end / closer to HCl end Do not penalise atoms in place of molecules/ particles	reverse arguments ammonia has lower density / has lighter molecules / smaller M_r references to solutions IGNORE smaller molecules		1
(b)	M1 less than 5 mins / less time (for white ring to form) M2 particles / molecules have more (kinetic) energy M3 and particles/ molecules move(s) / diffuse faster IGNORE reference to rate of reaction / more (successful/frequent) collisions Do not penalise atoms in place of molecules/particles	(forms more) quickly / sooner	gas has more energy	1 1 1

Question number	Expected Answer	Accept	Reject	Marks
8 (c)	<p>particles/molecules collide with air particles/molecules in air</p> <p>OR</p> <p>particles / molecules collide with one another / the wall (of the tube) Do not penalise collisions between ammonia and hydrogen chloride</p> <p>OR</p> <p>particles move in random direction / need many collisions (for white ring) to become visible / many particles of ammonium chloride must form (before white ring seen)</p> <p>Do not penalise atoms in place of molecules/ particles</p> <p>IGNORE reference to time taken for evaporation to take place</p> <p>IGNORE reference to time taken for reaction to take place</p>			1

Question number	Expected Answer	Accept	Reject	Marks
9 (b) (i)	M1 oxygen (atom)		oxygen molecule / O ₂ / oxide ion / oxygen ion	1
	M2 forms two bonds / smaller atom / has valency of 2 IGNORE more (oxygen) in the formula	more of them (n the diagram / structure)		1
	M2 DEP on M1, although allow M2 if oxygen mentioned but M1 not awarded because of reference to molecule/ion/O ₂			
	(ii) M1 giant (structure / lattice / atomic) IGNORE large / 3D M2 covalent	giant molecular / macromolecular		1
	M3 idea that covalent bonds are broken IGNORE bonds are loosened	overcome for broken		1
	M4 covalent bonds are strong / lots of energy required to break covalent bonds/ lots of heat required to break covalent bonds IGNORE high temperature needed Do not penalise silicon Max2 for mention of ionic or metallic bonding or intermolecular forces Max 3 if discussing diamond / carbon	many bonds are broken = M3 + M4		1

Question number	Expected Answer	Accept	Reject	Marks
10 (a)	(i) M1 Na $(1.15 \div 23) = 0.05$ (mol) O $(0.80 \div 16) = 0.05$ (mol) Accept correct alternative working		division by atomic numbers division upside down for M1 and M2	1
	M2 ratio 1:1	(moles are) the same/equal		1
	M2 DEP on M1			
	(ii) M1 $78 \div 39 = 2$	$39 \times 2 = 78$ / 78 is twice 39		1
	M2 Na ₂ O ₂	$23 \times 2 = 46$ and $16 \times 2 = 32$ (= 78)		1
	Final answer scores 2			
(b)	(i) Na ₂ O ₂ + 2H ₂ O → 2NaOH + H ₂ O ₂	multiples and fractions		2
	M1 all formulae correct M2 correct balancing M2 DEP on M1	equation csq on formula in (a)(ii), but Na and O must be in 1:1 ratio		
	(ii) Hydroxide / OH ⁻			1
	HO ⁻ / ⁻ OH			

Question number	Expected Answer	Accept	Reject	Marks
10 (b) (iii)	M1 two electrons between the oxygen atoms M2 all other electrons correct M2 DEP on M1 Allow any combination of dots and crosses			1 1

Question number	Expected Answer	Accept	Reject	Marks
11 (a) (i)	potassium / K^+	K		1
(ii)	iron(II) / Fe^{2+}			1
(iii)	iodide / I^-	I	iodine / I_2	1
(b)	<p>M1 use a (clean platinum / nichrome) wire / glass rod / silica rod</p> <p>IGNORE references to hydrochloric acid</p> <p>M2 (to put) solid / solution / M <u>in/over</u> a flame/burner</p> <p>M3 flame as either blue/roaring/non-luminous/Bunsen/blow torch</p> <p>OR</p> <p>burner described Bunsen/blow torch</p> <p>no marks if solid in a container, e.g. test tube/tray/beaker/pan</p>	<p>any method of introducing the solid into the flame, e.g. (wet) wooden spill / spatula / metal rod / tip or sprinkle in</p> <p>powder</p>	<p>any metal that will burn or melt in a flame (e.g. magnesium) or any metal that will colour the flame (e.g. copper)</p> <p>tongs / tweezers / (deflagrating) spoon</p> <p>luminous / yellow flame</p>	<p>1</p> <p>1</p> <p>1</p>

Question number	Expected Answer	Accept	Reject	Marks
11 (c) (i)	<p>reacts with / removes carbonate (ions)</p> <p>OR</p> <p>remove ions/substances/impurities that (form a) precipitate (with silver ions / silver nitrate)</p>	<p>formula</p> <p>removes ions that give a positive result (with silver ions / silver nitrate)</p>		1
(ii)	<p>M1 (hydrochloric acid) contains chloride ions</p> <p>M2 which interfere with test / make silver chloride</p> <p>OR</p> <p>M1 forms a (white) precipitate</p> <p>M2 of silver chloride</p> <p>Do not award either mark if wrong chemistry described, eg redox reactions, formation of iodine</p>	<p>gives a (white) precipitate / (false) positive result</p>	chlorine ions	1 1
(d)	<p>nitrate / NO_3^-</p> <p>If both name and formula given, both must be correct</p>			1

Question number	Expected Answer	Accept	Reject	Marks
12 (a)	$2\text{PbS} + 3\text{O}_2 \rightarrow 2\text{PbO} + 2\text{SO}_2$ <p>M1 all formulae correct M2 correct balancing</p> <p>M2 DEP on M1 IGNORE state symbols</p>	Multiples and fractions		2
(b) (i)	<p>Reduced <u>AND</u> oxygen has been removed</p> <p>IGNORE It / PbO gains electrons Do not penalise molecules</p>	arguments based on decrease in oxidation number of <u>Pb</u> / gain of electrons by <u>Pb²⁺</u> / lead ions		1
(ii)	<p>M1 Mr (PbO) = 223</p> <p>(moles method)</p> <p>M2 n (PbO) = $44.6 / 223 (=0.2)$</p> <p>M3 mass of C = $(2/2 \times 12 = 1.2)$</p> <p>(mass ratio method)</p> <p>M2 44.6 require $1 / 44.6 \times \frac{12}{446}$</p> <p>M3 44.6 require $2 / 1.2$</p> <p>Calculations with and without use of 10^6 are acceptable</p> <p>mark csq at each stage</p> <p>Correct final answer with or without working</p>	446		1 1 1

scores 3			
Final answers that may score 2 are: 0.6 / 2.4 / 4.8 / 0.12 / 12			

Question number	Expected Answer	Accept	Reject	Marks
12 (c) (i)	(silver is / it is) <u>more</u> soluble in zinc / <u>less</u> soluble in lead	soluble in zinc but insoluble in lead		1
(ii)	(it is) less than / equal to 530 (°C)		implication that Zn and Ag melting points are both less than or equal to 530 °C	1
(iii)	M1 zinc/it is lower / silver is higher			1
	M2 zinc turns into a vapour / gas (when heated) while silver remains	boils off (as a gas) first when heated		1
	M2 DEP on M1 IGNORE references to melting point			
(iv)	silver is expensive / valuable			1

Question number	Expected Answer	Accept	Reject	Marks
13 (a) (i)	4.83 (g)			1
(ii)	3.78 (g)			1
(iii)	<p>M1 $n(\text{ZnSO}_4) = 4.83 \div 161 \neq 0.03$</p> <p>M2 $n(\text{H}_2\text{O}) = 3.78 \div 18 \neq 0.21$</p> <p>M3 $x = n(\text{H}_2\text{O}) \neq n(\text{ZnSO}_4) = 7$</p> <p>CSQ on (i) and (ii)</p> <p>Do not penalise non-integer values of x</p> <p>Correct final answer with no working = 1</p> <p>Correct final answer with some <u>correct</u> working = 3</p>	<p>$(18x \div 161) = (3.78 \div 4.83)$</p> <p>$x = ((3.78 \div 4.83) \times 161) \div 18$</p> <p>$= 7$</p> <p>equivalent alternative calculations</p>		<p>1</p> <p>1</p> <p>1</p>
(b)	<p>to remove all the water</p> <p>NOT just to remove the water</p>	to make sure the solid is anhydrous / fully dehydrated		1

Question number	Expected Answer	Accept	Reject	Marks
13 (c)	M1 <u>anhydrous</u> / <u>white</u> copper sulfate IGNORE crystals	<u>anhydrous</u> cobalt chloride / <u>blue</u> cobalt chloride (solid or paper)		1
	M2 turns blue if oxidation number of copper given, must be +2 M2 DEP on M1 correct or near miss IGNORE references to determining melting and/or boiling point, even if incorrect IGNORE references to acid/base indicators or UI, even if incorrect	turns pink if oxidation number of cobalt given, must be +2 copper sulfate turns from white to blue = 2 cobalt chloride turns from blue to pink = 2 dehydrated in place of anhydrous		1

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