

Please check the examination details below before entering your candidate information

Candidate surname	Other names
Centre Number	Candidate Number
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Pearson Edexcel International GCSE (9–1)

Friday 13 June 2025

Morning (Time: 1 hour 15 minutes)	Paper reference	4CH1/2C
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Chemistry

UNIT: 4CH1

PAPER: 2C

<p style="font-weight: bold;">You must have:</p> <p>Calculator</p>	Total Marks
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Instructions

- Use **black** ink or ball-point pen.
- If pencil is used for diagrams/sketches/graphs it must be dark (HB or B).
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided
– *there may be more space than you need.*
- Show all the steps in any calculations and state the units.

Information

- The total mark for this paper is 70.
- The marks for **each** question are shown in brackets
– *use this as a guide as to how much time to spend on each question.*

Advice

- Read each question carefully before you start to answer it.
- Write your answers neatly and in good English.
- Try to answer every question.
- Check your answers if you have time at the end.

Turn over ►

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Answer ALL questions.

Some questions must be answered with a cross ☒. If you change your mind about an answer, put a line through the box ☒ and then mark your new answer with a cross ☒.

1 This question is about elements, compounds and mixtures.

(a) The box gives the names of some substances.

calcium	fluorine	helium	neon
petrol	silicon	water	

Choose a substance from the box to answer each of these questions.

(i) Give the name of a compound. (1)

(ii) Give the name of a mixture. (1)

(iii) Give the name of a diatomic molecule. (1)

(iv) Give the name of a metal element. (1)

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(b) (i) Name the technique used to separate crude oil into fractions.

(1)

(ii) Give one use of the kerosene fraction.

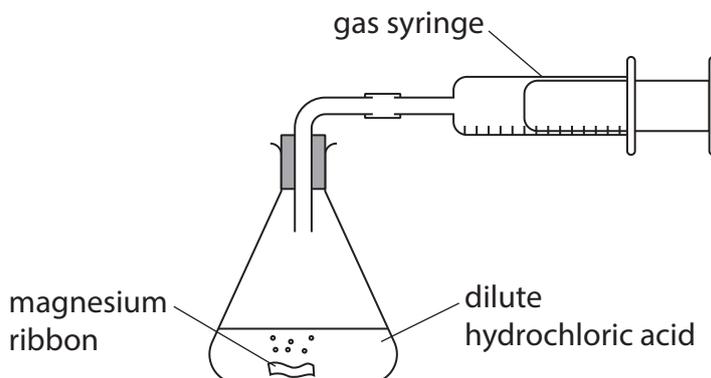
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(Total for Question 1 = 6 marks)



2 Magnesium reacts with dilute hydrochloric acid to form magnesium chloride and hydrogen.

The hydrogen is collected in a gas syringe.



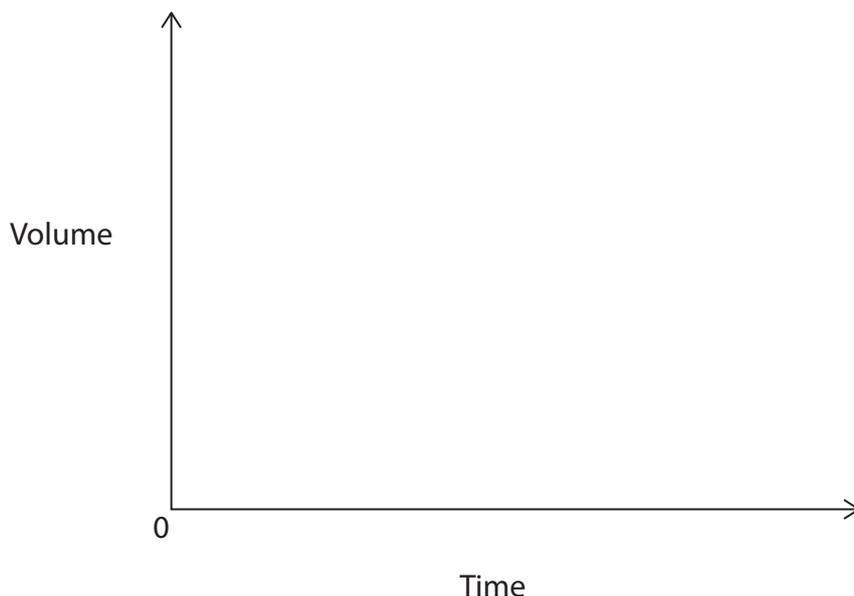
The reaction is exothermic and magnesium is in excess.

(a) State what is meant by the term **exothermic**. (1)

(b) Complete the chemical equation for the reaction. (1)



(c) Draw a sketch graph to show how the volume of hydrogen changes with time until the reaction has finished. (2)



- (d) Explain why increasing the temperature of the hydrochloric acid affects the rate of the reaction.

Refer to collision theory in your answer.

(4)

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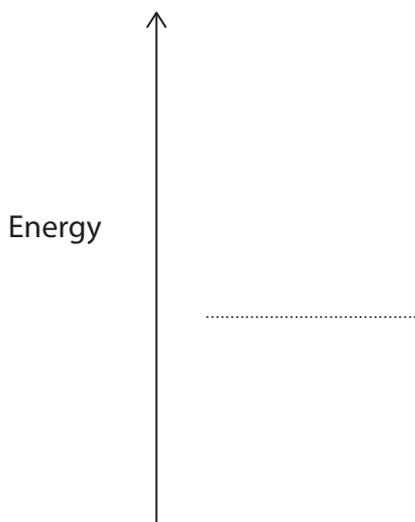
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- (e) Draw a reaction profile diagram for the reaction, showing ΔH and activation energy.

Label the reactants and products.

(5)



(Total for Question 2 = 13 marks)



3 This question is about Group 2 elements and their compounds.

(a) A sample of magnesium contains three isotopes.

Mass number of isotope	24	25	26
Percentage abundance (%)	76.5	10.5	13.0

(i) State, in terms of subatomic particles, what is meant by the term **isotopes**.

(2)

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(ii) Calculate the relative atomic mass (A_r) of magnesium.

Give your answer to one decimal place.

(3)

$A_r =$

(b) Which of these compounds is an insoluble solid?

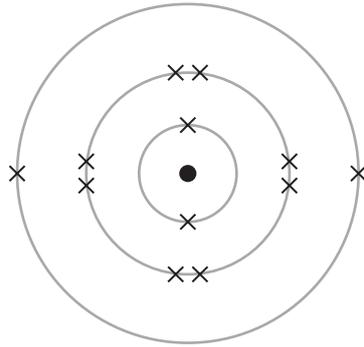
(1)

- A calcium nitrate
- B calcium sulfate
- C magnesium nitrate
- D magnesium sulfate

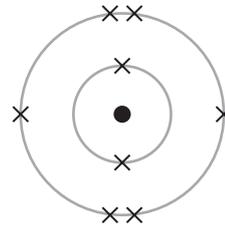


(c) Magnesium oxide is an ionic compound.

The diagrams show the atoms of magnesium and oxygen.



Mg



O

Draw diagrams to show the structure of each ion in magnesium oxide.

Include the charge on each ion.

(3)

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(e) A concentrated aqueous solution of magnesium chloride is electrolysed.

Chlorine forms at the positive electrode.

(i) Give the half-equation for the reaction at the positive electrode. (1)

(ii) Give a reason why this reaction is classified as oxidation. (1)

(Total for Question 3 = 16 marks)



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4 This question is about the reactions of ethanol and ethanoic acid.

(a) (i) Ethanol reacts with potassium dichromate(VI) to form ethanoic acid.

Name the other reagent needed in this reaction.

(1)

(ii) State the colour change that occurs when ethanol reacts with potassium dichromate(VI).

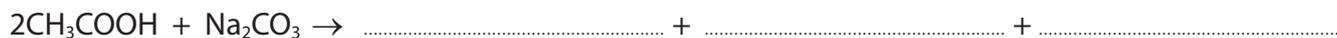
(2)

from to

(b) Ethanoic acid reacts with solid sodium carbonate.

(i) Complete the chemical equation for this reaction.

(2)



(ii) State two observations seen in this reaction.

(2)

1

2



- (c) A student prepares the ester, ethyl ethanoate, by reacting ethanol with ethanoic acid.

This is the student's method.

- mix equal amounts of ethanol and ethanoic acid in a boiling tube
- add a few drops of concentrated sulfuric acid
- place the boiling tube in a hot water bath for several minutes

- (i) State the role of sulfuric acid in this reaction.

(1)

- (ii) State why the student heats the mixture in a water bath instead of directly with a Bunsen burner flame.

(1)

- (iii) State how the student would know that ethyl ethanoate has formed.

(1)

- (iv) Draw the displayed formula of ethyl ethanoate.

(2)

(Total for Question 4 = 12 marks)



5 A student does a titration to find the concentration of a solution of sulfuric acid.

This is the student's method.

- fill a burette with sulfuric acid and record the reading
- transfer 25.0 cm^3 of potassium hydroxide solution to a conical flask
- add a few drops of methyl orange to the conical flask
- place the conical flask on a white tile
- swirl the flask continuously while adding sulfuric acid from the burette
- keep adding sulfuric acid until the end-point has been reached
- record the rough reading at the end-point

(a) (i) Name the most suitable piece of apparatus for measuring 25.0 cm^3 of potassium hydroxide solution.

(1)

(ii) State why the student places the conical flask on a white tile.

(1)

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(b) The diagram shows the final burette reading.



The initial burette reading is given in the table.

Complete the table, giving both values to the nearest 0.05 cm^3 .

(2)

final burette reading	
initial burette reading	4.15
volume of sulfuric acid added in cm^3	

(c) The student repeats the titration more accurately four times.

The table shows their results.

titration number	1	2	3	4
volume of sulfuric acid added in cm^3	27.40	27.65	27.30	27.35
concordant results				

Concordant results are those within 0.20 cm^3 of each other.

(i) Place ticks (\checkmark) in the table to show which results are concordant.

(1)

(ii) Use the concordant results to calculate the mean volume of sulfuric acid added.

(1)

mean volume of sulfuric acid = cm^3



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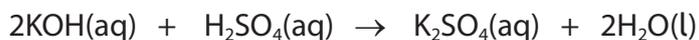
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(d) Another student also does the titration.

- (i) The student uses 25.0 cm³ of potassium hydroxide solution with a concentration of 0.500 mol/dm³.

The mean volume of sulfuric acid added is 26.50 cm³.

This is the equation for the reaction.



Calculate the concentration of the sulfuric acid.

(3)

concentration of sulfuric acid = mol/dm³

- (ii) The student repeats the titration without the methyl orange indicator.

They add exactly 26.50 cm³ of sulfuric acid to the potassium hydroxide solution.

Describe how the student could obtain pure, dry potassium sulfate crystals from the reaction mixture.

(4)

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(Total for Question 5 = 13 marks)



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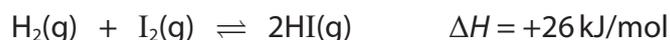
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6 Hydrogen reacts with iodine vapour to form hydrogen iodide, HI

This is the equation for the reaction.



(a) (i) Give the test for hydrogen.

(1)

(ii) State what the symbol \rightleftharpoons represents.

(1)

(b) In a sealed container at equilibrium, the reaction conditions are 500°C and 2 atm.

(i) The temperature of the reaction mixture is increased to 600°C, but the pressure is kept at 2 atm.

Explain the effect, if any, on the yield of hydrogen iodide at equilibrium.

(2)

(ii) The pressure of the reaction mixture is decreased to 1 atm but the temperature is kept at 500°C.

Explain the effect, if any, on the yield of hydrogen iodide at equilibrium.

(2)

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(c) An excess of hydrogen reacts with 50.8 g of iodine.

At equilibrium, the yield of hydrogen iodide is 80%.

Calculate the volume, in cm^3 , of hydrogen iodide produced at rtp.

Give your answer in standard form.

[at rtp, molar volume = $24\,000\text{ cm}^3$]

(4)

volume of hydrogen iodide = cm^3

(Total for Question 6 = 10 marks)

TOTAL FOR PAPER = 70 MARKS

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